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, 2012

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(Pb, Cd, Zn Cu).

5 mg/L

100 mg/L

, <90 m, 90-315 m, 315-1000 m 1000-4000 m,  
315-1000 m.

1 mm.

XRD, XRF, TG,

95%.

Pb Cu,

Zn Cd

93%.

( )

Pb>Cd>Zn=Cu.

Zn Cu

Pb Cd

Overpopulation and expansion of industrial activities of recent years has led to a steadily increasing pollution of water supplies with heavy metals. With the rapid development of industries such as metal plating facilities, mining operations, fertilizer industries, tanneries, batteries, paper industries and pesticides, etc., heavy metals wastewaters are directly or indirectly discharged into the environment increasingly, especially in developing countries. Faced with more and more stringent regulations, nowadays heavy metals are the environmental priority pollutants and are becoming one of the most serious environmental problems. So these toxic heavy metals should be removed from the wastewater to protect the people and the environment. Unlike organic contaminants, heavy metals are not biodegradable and tend to accumulate in living organisms and many heavy metal ions are known to be toxic or carcinogenic. As a consequence, these toxic elements is necessary to be eliminated in the effluents so as people and environment to be protected.

Adsorption is one of the most interesting methods because of its high efficiency in removing heavy metals from different fluxes. A lot of studies have been held in order to be found effective organic and inorganic adsorbents. These materials are abundant and possess, generally, strong affinity and high capacity, two of the most basic properties of adsorbents. The removal of heavy metals that is based on their sorption on mineral surfaces is constitutes an alternative. Carbonate minerals and rocks strongly interact with metallic ions removing them from aqueous solutions. Calcite, magnesite and dolomite have been studied as potential adsorbents of heavy metals from solutions. The interaction mechanism seems to be a combination of ion exchange and precipitation on the carbonate surface.

The main objective of the present work is the study of dolomite and dolomitic marble are studied as cheap and effective inorganic adsorbents as a function of the particle size distribution and the flow conditions of watewaters with ions of heavy metals and toxic metallic ions (Pb, Cd, Zn and Cu).

Particularly, batch and column sorption experiments were conducted aiming at determining the retention capacity for the above metals in stable and continuous flow conditions, respectively. For comparison reasons, in the case of column experiments, calcitic marble is also studied. The adsorption of metals is examined either in monometallic solutions of each metal, or in their polymetallic solution, with concentrations of 100 mg/L for both zinc and copper and 5 mg/L for lead and cadmium. More precisely, in the adsorption under stable flow conditions both types of solutions are used, while in the adsorption under continuous flow conditions only polymetallic solutions is implemented. Basic parameter of the study is particle size distribution as well. In the case of stable flow conditions, dolomite is examined in different particle fractions, <90  $\mu$ m, 90-315  $\mu$ m, 315-1000  $\mu$ m, 1000-4000  $\mu$ m, while dolomitic marble only in that of 315-1000  $\mu$ m. In the case of continuous flow conditions, the adsorbentsø particle size is <1 mm. The investigation of the mechanism of the adsorption is held through the leachability of the adsorbents. The implementation of analysis methods of XRD, XRF, TG, as well as of chemical analysis and the determination of loss of ignition and z-potential gives the ability of characterizing adsorbents not only before but also after the process of adsorption.



Experimental data show that under stable flow conditions, dolomite and dolomitic marble have a very satisfactory adsorption capacity in the studied particle size distribution for all the metals, with adsorption percentages generally beyond 95%. On the contrary, in common aqueous environment, the metallic ions behave differently for both adsorbents. In particular, retention of Pb and Cu is considerably favored, while that of Zn and Cd is much reduced.

In the case of adsorption under continuous flow conditions, the three adsorbents appear high levels of retention, beyond 93%, for all the metallic ions. Besides, the maximum adsorption performance tends to be achieved directly for all the metals.

In both cases of flow it is noticed that the difference in the mineralogical tissue between dolomite and dolomitic marble, i.e. the bigger particle size of the grains (crystals) and the preferred direction of the carbonate crystals of dolomitic marble, as well as the more massive structure of the marble do not seem to differentiate the adsorption capacity of the adsorbents.

The study of leachability of adsorbents which come from stable and continuous flow conditions show that it is greater in fixed beds and in both cases the leachability order followed is  $Pb > Cd > Zn = Cu$ . Experimental data conclude that in stable flow conditions lead and cadmium, except for ionexchange, complexation and precipitation as basic adsorption mechanisms, are expected to be physically adsorbed on the surface and be trapped in small percentage, while the retention of zinc and copper is attributed to adsorption through ion exchange and formation of carbonate salts through complexation. In the case of fixed beds, adsorption through ion exchange, surface complexation and precipitation is indicated as retention mechanism of Zn and Cu cations, while for Pb and Cd, apart from adsorption through surface precipitation and ion exchange, physical sorption also takes place.

Last but not least, the characterization of the adsorbents derived from the fixed beds reveals the formation of new substance during the process of adsorption, which are the same for the dolomite and the dolomitic marble and different for the calcitic marble.



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## 1.1

### 1.1.1

(CO<sub>3</sub>)<sup>2-</sup> [1].

O-C-O C-O 120° CO<sub>3</sub> 1,25-1,35 Å (CO<sub>3</sub>)<sup>2-</sup> [2].

Mn<sup>2+</sup>, Zn<sup>2+</sup>, 150 Ca<sup>2+</sup>, Mg<sup>2+</sup>, Ba<sup>2+</sup>, Sr<sup>2+</sup>, Pb<sup>2+</sup>, Fe<sup>2+</sup>, [4]. [2]. [3].

[1]. [2,3,5]. [5]. : Ca, Mg, Fe, Mn Zn (CO<sub>3</sub>)<sup>2-</sup> Ca Mg, Fe Mn [5].

### 1.1.2

( 1.1).

1.1 [2].

[2].

			(Å)			
			a	c		
	CaCO <sub>3</sub>	2.71	4.989	17.061	1.658	1.486
	MgCO <sub>3</sub>	3.00	4.632	15.012	1.700	1.509
	FeCO <sub>3</sub>	3.97	4.691	15.379	1.875	1.635
	MnCO <sub>3</sub>	3.70	4.768	15.635	1.816	1.597
	ZnCO <sub>3</sub>	4.43	4.652	15.025	1.850	1.625
Otavite	CdCO <sub>3</sub>	4.96	4.923	16.787	1.828	1.607
Gaspéite	NiCO <sub>3</sub>	4.39	4.608	14.805	1.930	1.721

1914 (1.2),  
(CO<sub>3</sub>)  
(CO<sub>3</sub>)  
Ca  
CO<sub>3</sub>

Cl Ca  
NaCl [3].  
CO<sub>3</sub>

Ca  
[6].

X  
NaCl

R<sub>3</sub>c.  
W. L. Bragg  
Na.



1.1: [37].

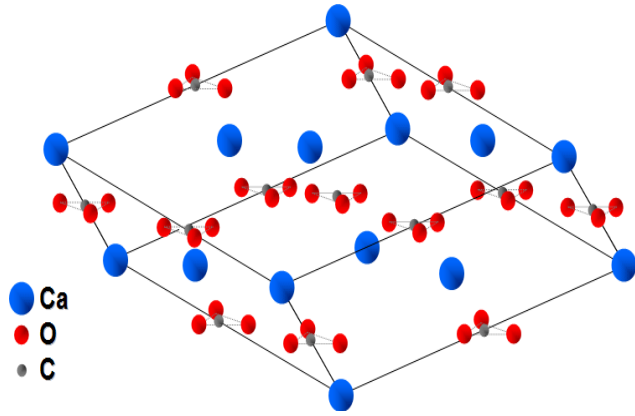
CO<sub>3</sub>  
{0001}

[2].  
CO<sub>3</sub>  
60

Ca [2,6]. T  
(CO<sub>3</sub>).

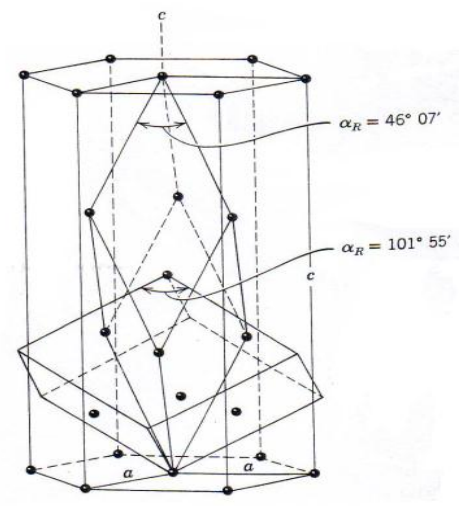
Ca  
(CO<sub>3</sub>) [3].

Ca  
CO<sub>3</sub>  
(CO<sub>3</sub>)  
[6].  
1.2)  
Ca



● Ca  
● O  
● C

(a)



(b)

**1.2:** (a)  $\text{CaCO}_3$ , (b)  $\text{H}_2\text{O}$  ( [3,38]. )

{10 $\bar{1}$ 1} ( 1.2),



(a)

(b)

**1.3:** (a) (b) [37].

$\text{CaCO}_3$  ( 2).  
 Ca:O (=0.714)  
 6- 8- (=0.732)

Ca : 6- Ca  
 [3].  
 MO<sub>6</sub> c/a.  
 otavite gaspéite [2].

**1.1.3**

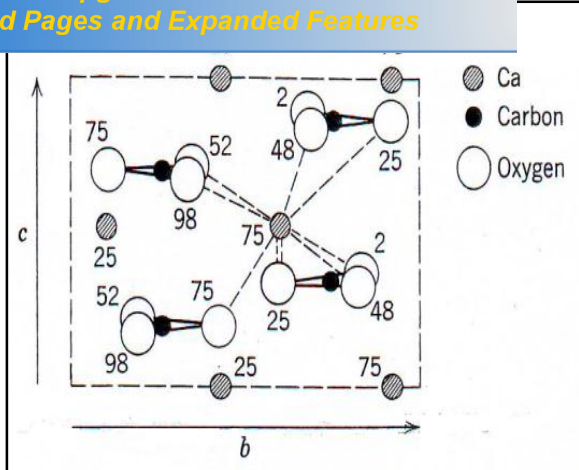
1.0 Å), (CO<sub>3</sub>) ( 6-  
 Pmcn. CaCO<sub>3</sub> ( 1.5a)  
 Ca 6- ( )  
 BaCO<sub>3</sub>, SrCO<sub>3</sub> PbCO<sub>3</sub>, 9- 1.5b). CaCO<sub>3</sub>  
 [3]. 1.2 [2].



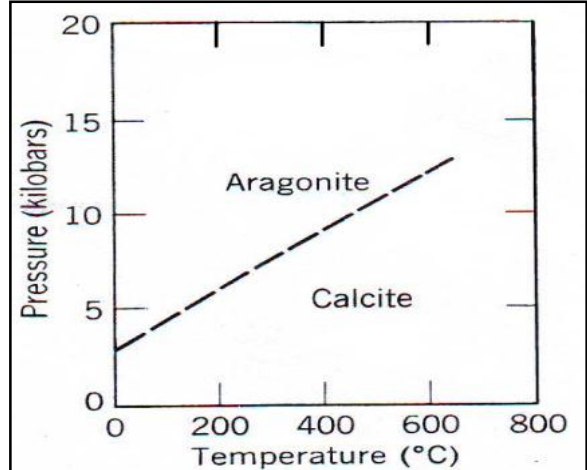
**1.4:** [37,43].

[2]. CaCO<sub>3</sub> (CO<sub>3</sub>), (CO<sub>3</sub>)  
 c, (CO<sub>3</sub>)  
 [3]. {001}  
 [2].  
 [3].

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(a)



(b)

1.5 : (a) H

CaCO<sub>3</sub>,

(100).

a. To

Ca-O (b) H

[3].

(CO<sub>3</sub>)

Ca,

Ca,

Ca [6].

[3,6].

1.2:

[2].

			(Å)					
			a	b	c			
	CaCO <sub>3</sub>	2.95	4.960	7.964	5.738	1.531	1.681	1.684
	SrCO <sub>3</sub>	3.7	5.090	8.358	5.997	1.517	1.661	1.667
	BaCO <sub>3</sub>	4.3	5.312	8.896	6.428	1.529	1.676	1.677
	PbCO <sub>3</sub>	6.55	5.180	8.492	6.134	1.803	2.074	2.076

Ca

Ba ,

BaCa(CO<sub>3</sub>)<sub>2</sub>,

CaMg(CO<sub>3</sub>)<sub>2</sub>

barytocalcite,  
CaCO<sub>3</sub>-MgCO<sub>3</sub>.

( 1.2).

( 1.3).

kutnahorite minrecordite,  
[2].

R<sub>3</sub>  
Ca

Mg, Fe, Zn Mn.

[3,19,20]:

| |CO<sub>3</sub>|B|CO<sub>3</sub>| |CO<sub>3</sub>|B|CO<sub>3</sub>| |CO<sub>3</sub>|B|í

Ca

Ca

[6].

[2,3].

Ca<sup>2+</sup>

Ca Mg  
CO<sub>3</sub>  
Mg<sup>2+</sup> (33%)

Ca Mg, 2-fold

CaCO<sub>3</sub> MgCO<sub>3</sub>, Ca:Mg=1:1 [3].

Fe Mn

[19-21].

1.3:

[2].

			(Å)			
			a	c		
Kutnahorite	CaMg(CO <sub>3</sub> ) <sub>2</sub>	2.85	4.807	16.003	1.679	1.500
	Ca(Mg,Fe)(CO <sub>3</sub> ) <sub>2</sub>	3.01	4.811-4.831	16.042-16.166	1.690-1.750	1.510-1.548
Minrecordite	CaMn(CO <sub>3</sub> ) <sub>2</sub>	3.12	4.873	16.349	1.727	1.535
	CaZn(CO <sub>3</sub> ) <sub>2</sub>	3.45	4.818	16.029	1.750	1.550

CaCO<sub>3</sub> MgCO<sub>3</sub> ( 1.6b).

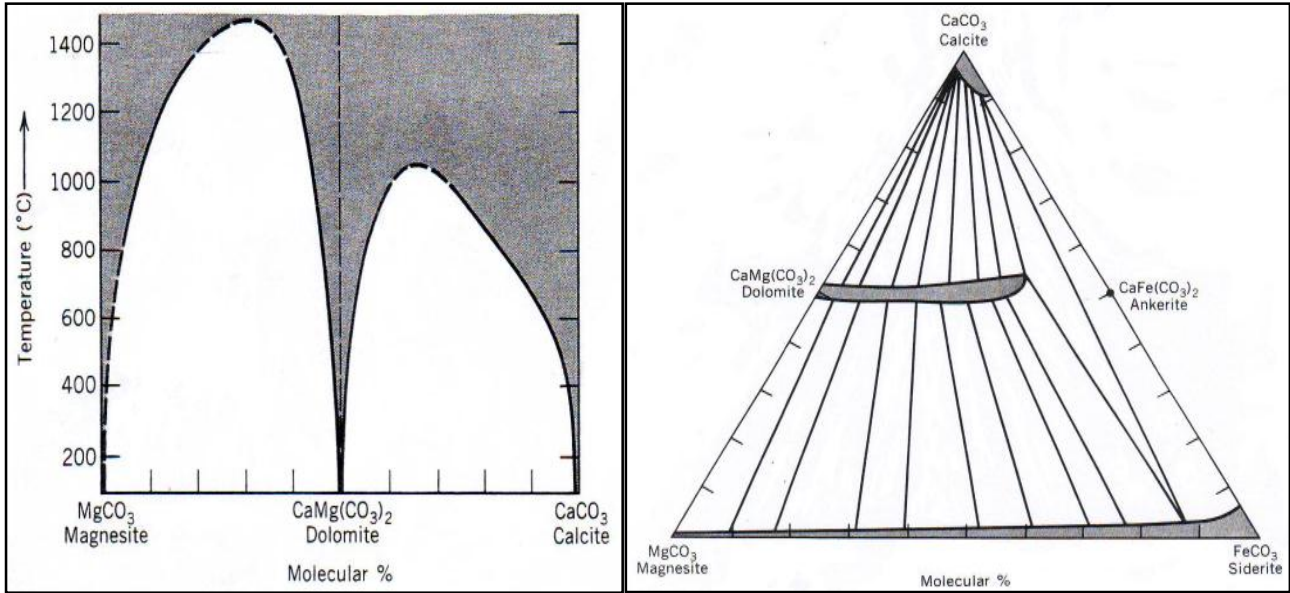
( 700 C),

Ca:Mg=1:1, 6a.

1000-1100 C



1.6a [3].



(a)

(b)

1.6 : (a) CaCO<sub>3</sub>-MgCO<sub>3</sub>

CO<sub>2</sub>

(b)

CaO-MgO-FeO-CO<sub>2</sub>.

[3].

1.1.5

huntite

Oldoinyo nyerereite [2].

[3].

1.1.6

O<sup>2-</sup>, OH<sup>-</sup>

Cl<sup>-</sup>.



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,  $\text{FeCO}_3$ ,  $\text{Mn}^{2+}$ ,  $\text{Cu}^{2+}$ , [4].

( , , - , . . ) , . . .

actonolite, . . .

( ) Fe Pb, [4].  $[\text{CO}_3]^{2-}$

( , , , . . ), ( , , , . . ), . . . [4].

$\text{CO}_2$ , HCl, [4].

, , . . .

, , . . .

( , ) ( , ) [4].

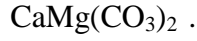
( , ) ( , ) ( , )

( , ) [4]. ( , ) ( , )

(The Iceland spar), ( , ) ( , ) parasite bastnäsité dawsonite

[4].

1.1.7.1



[13].

[2,15-17,21,22].

1.1.7.2

: CaO 30,4%, MgO 21,7%  $\text{CO}_2$  47,9%. O

Ca:Mg

58:42

47,5:52,5 [3].

Ca:Mg=1:1 ( 6a)

Mn  
Mg

[5].

Fe Mn  
 $\text{Fe}^{2+}$  , ,

,  $\text{CaFe}(\text{CO}_3)_2$  [6].

$\text{CaMg}(\text{CO}_3)_2$

$\text{Fe}^{2+}$

$\text{Fe}^{2+}/\text{Mg}=2,6$ .

1,

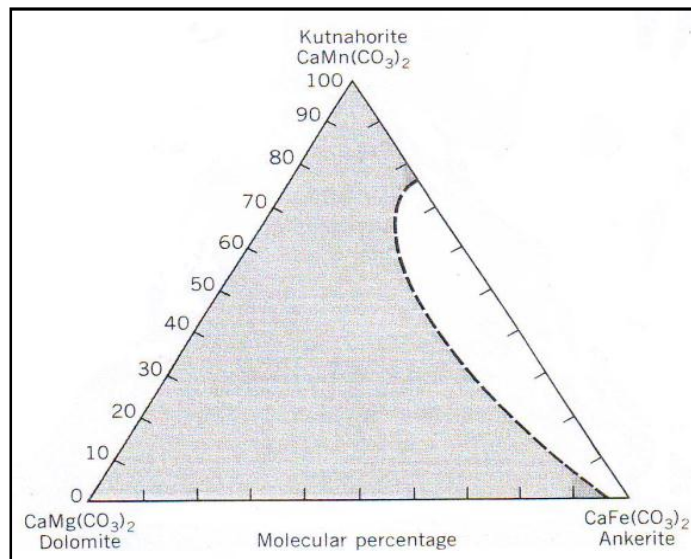
[8]. Mn

kutnahorite,  $\text{CaMn}(\text{CO}_3)_2$  ( 1.7) [3,6].

Co, Pb, Zn, Ce

Ca

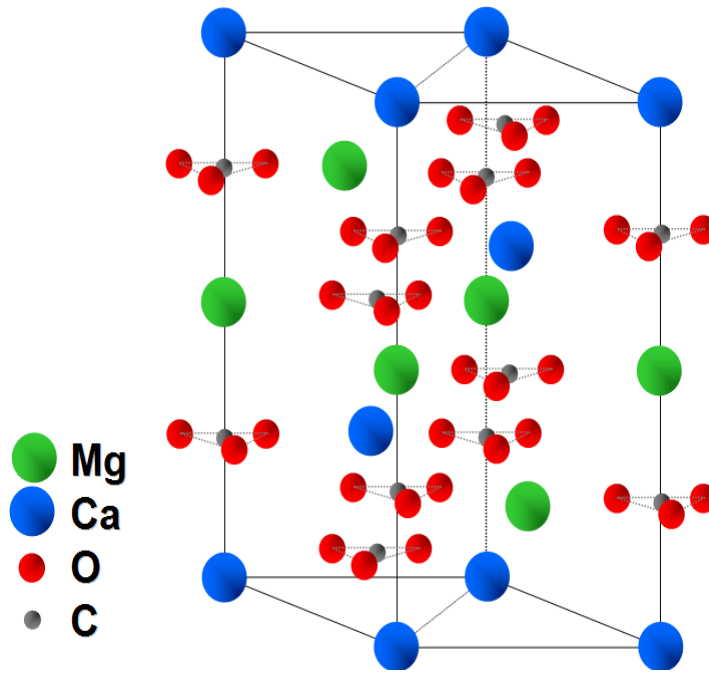
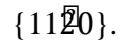
[5].



1.7:

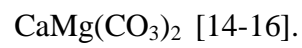
$\text{CaMn}(\text{CO}_3)_2$ -  $\text{CaMg}(\text{CO}_3)_2$ -  $\text{CaFe}(\text{CO}_3)_2$

[3].



1.8:

[38]



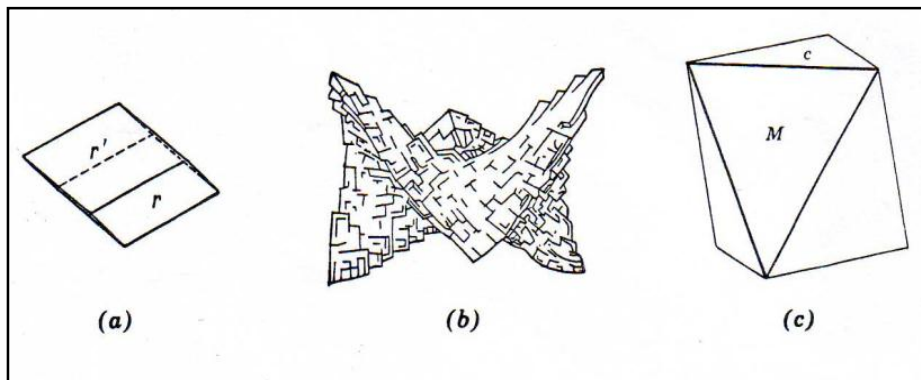
[1].

( 4. ).

1.4: [5,3].

	a = 4,84 b = 15,96 c = 15,96
Z	3
	$R\sqrt{3}$
	$\sqrt{3}$
d	2,88 (10) 2,19 (4) 2,01 (3) 1,800 (1) 1,780 (1)

( 9c) , ( 9a) ,  
 {0001} {10 $\bar{1}$ 1} 1.9b , 1.10) ,  
 {02 $\bar{2}$ 1} ,  
 [3,5,14].



1.9: [3].



1.10:



[39,40]

1.1.7.5

( 1.11) cleavage

[3-6,10,14].



1.11:



[37].

1.1.7.6

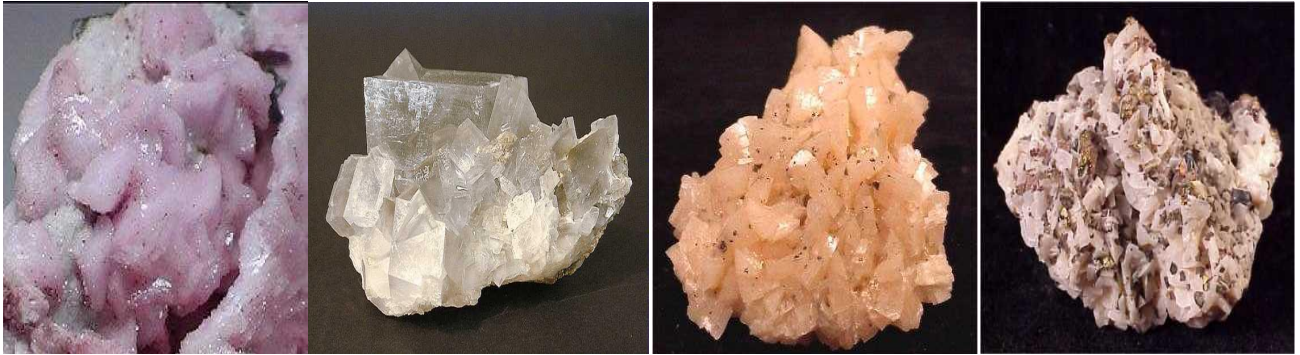
{021}.

2,85.

{101}

3,5-4

) , =1,500 =179. , R.I. ( , [3,5,6,10,14].



1.12:

[37].

1.1.7.7

- $\{10\bar{1}1\}$
- a) , , :
  - b) , , [2,7].
  - c) , , HCl, HCl.
  - d)  $\{01\bar{1}2\}$ ,  $\{02\bar{2}1\}$  ( HCl. ) [1,2,3]. [2,8].

Yellow , Alizarin Red S 20 Titan Trypan

Alizarin Red S [2].  
/15 HCl, [1].

[2].

X-ray Powder Diffraction  
3.03 Å

2.88 Å [2].

[8].

550 C

[1].

[3].

### 1.1.7.8

Mg [3,5].

Ca

[6].

(dolostones)

[3].

(dolostone)

[7].

, quartz

, talcose

[3,5,6].

Binnenthal  
Guanajuato

Traversella

Piedmont

Tyrol,

Donbass,



Joplin Missouri [3,6].  
 Missouri  
 Cornwall  
 -quartz  
 Mother Lode Ontario  
 Porcupine Larder Lake Quebec.  
 ( sericite) [3,4,6].

**1.1.7.9**

Mg, Mg,  
 Portland,  
 CaO, (MgO),  
 [3,4,6,7,10].  
 [7].

**1.2**

**1.2.1**

15% 0,25% 2,7  
 75% ( 7.000)  
 [12,15,36].  
 [12,15].  
 □  
 ■ ( )  
 ■ ( . . )



) [9].  
 , , ,  
 ( , , . )  
 : , , ,  
 , . . [11].

## 1.2.2

### 1.2.2.1

(CaCO<sub>3</sub>)  
 ( ,  
 ). 90% 10% ,  
 , , ,  
 , , ,  
 , , , [11,13,36].  
 [9]. 200-2.000 2,7 3.  
 , [13].



1.13: [37,41].

50%

[9].

[11].  
MgCO<sub>3</sub>

: Ca- ( <5% MgCO<sub>3</sub>), Mg- ( 5-35% MgCO<sub>3</sub>) [9].

1.2.2.2

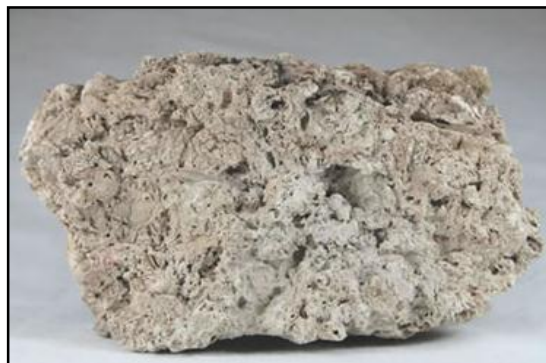
ó ó ( )

( ), [13].

[9,13].

[9].

[13].



1.14:

CaCO<sub>3</sub>,  
CaCO<sub>3</sub> [41].

[3,13].



(a)

(b)

1.15: (a)

(b)

[41].

### 1.2.3

#### 1.2.3.1

[13].  
 $\text{CaMg}(\text{CO}_3)_2$ ,  
 : 21,9% MgO, 30,4% CaO 47,7%  $\text{CO}_2$  [9].  
 90-100 % 19,6-21,7%  
 MgO [13]. 57% mole Ca 43% mole Mg  
 50% mole Ca Mg.  
 Frenchman D. Dolomieu,  
 $\text{Fe}^{2+}$ [9]. 18 [13].

[13,51,50].

[13].

[11].



1.16:

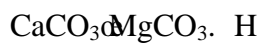
[37,45,46].

80:1 / 1:3 3:1

[11].

[15].

1.2.3.2 ó



Mg pH

[13].



.



11-13%

[11,12,15-17,28,31-34]:

1.5:

[11].

	%
	0-10
	10-50
	50-90
	90-100

Mg

( )

[13].

[31].

1.

2.

[3,29].

[15,16].

[13].

[13,33].



1.17:



[37].

[13].

Funalui



1.2.4.1

3, / [3,9,30]. Mohs 2,7. Mohs (3,5-4) (2,9). [13].



1.18:



[37,42].

[30]. [9,13] [32].

[9].

ó  
[9,13].

õ ö,

[9].

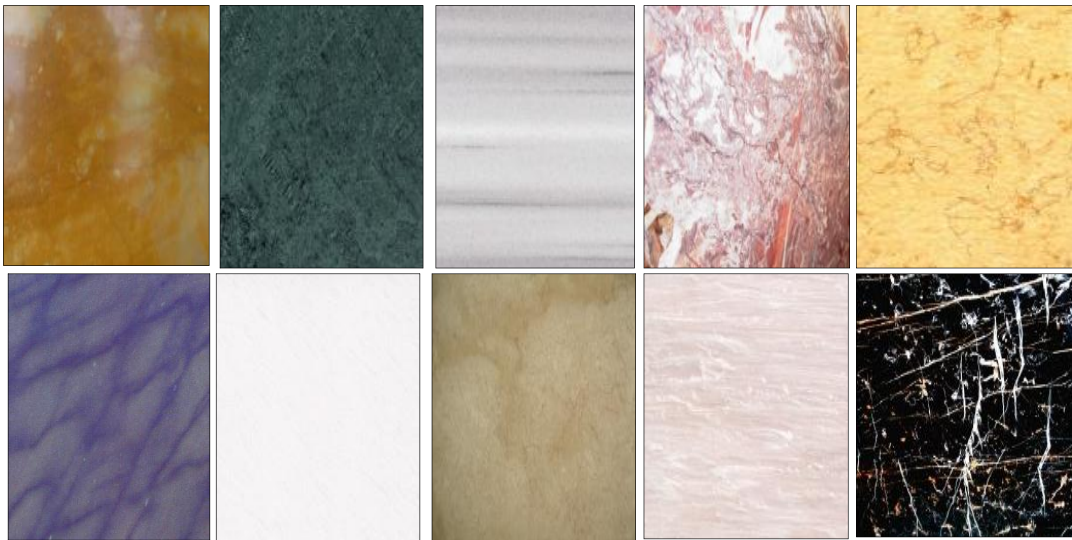
[13].

[3,9].

( . . )

( . . . . . ),

[9].



**1.19:**

[44].

:  $d=0,01-0,5$  mm,  $d=0,06-2$  mm  $d=2-6$  mm.  
 $d>6$  mm.

[13].

« »,

[11].  
 $(2.700 \text{ kg/m}^3)$   
 $1.500 \text{ kg/cm}^3$  (  
 Mohs [11].

3 4

[11].  
 [29].

[11].  
 [32].

15 mm,

25 mm  
 35 mm.

« »,

[11].  
 ( ) [9].



1.20:

[44].

[13].

[9].

16.000.000 tn/year,  
2.500.000 tn.

[24-26]

### 1.2.4.2

∅

∅

[3,31].

( )

[13,31].

[31].

[5].

### 1.2.5.1

1993		2004	
Commercial	Industrial Marble Co. Ionian Kalk	Ionian Kalk, Zafranas-Petrochem, Dionysos-Pentelikon	
>99,6% CaCO <sub>3</sub> , <0,07% Al <sub>2</sub> O <sub>3</sub> , <0,02% SiO <sub>2</sub>	<0,01% Fe <sub>2</sub> O <sub>3</sub> .		30
65%		150.000	
		(>96%),	
	Ionian Kalk	2004	
	150.000 /		
	Zafranas-Petrochem		
		100.000	50%
	Dionysos-Pentelikon		9
	98%	0,5%	0,5%
1,5 Mohs)			1%
		40.000	300.000 /
			5

1.2.5.2

[9,13].

15  
2003

3.500 clinker, 25% 30%

petcoke, 500.000 / 755  
15%

46 33

7-10

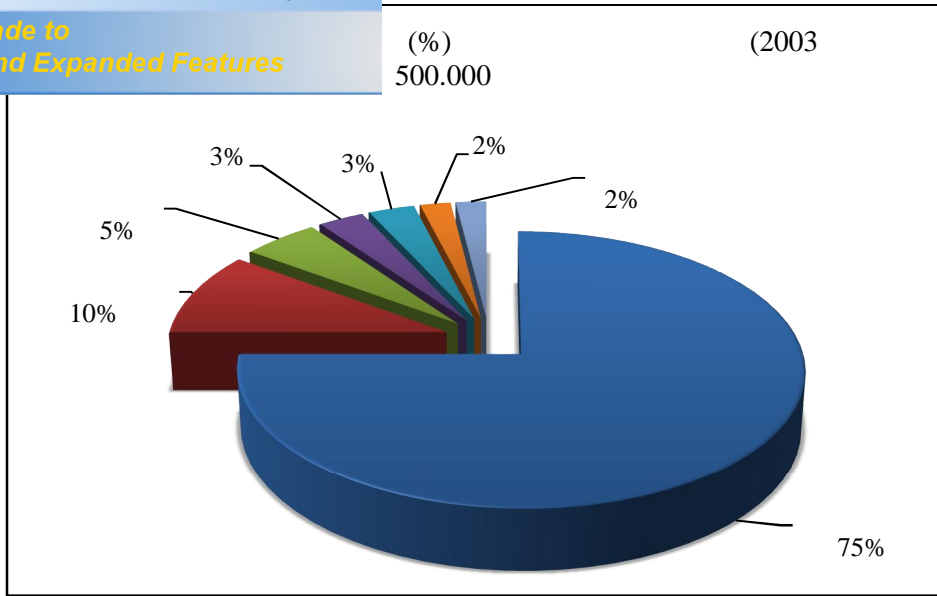
50-70 p/ [9].

ITALCEMENTI, LAFARGE BETON,

2010

50 20% 2009.

2011 [42].



1.21: (%) (2003) [9].

1.2.5.3

[13].

[9].

MgO>20%,

(

[13].

37%

( ... )

(

[9].

[15].

**1.2.5.4**

( ) .

( 1.12 )

1.28

210.000 m<sup>3</sup> .

86

( 1800 p/m<sup>3</sup>).

1.800 km<sup>2</sup>

40

1,8-2,2

300.000 /

700 . p.

3.500

( , , ),

20

40.000 . 60%

( ) .

:

- 
- 
-





**PDF Complete**

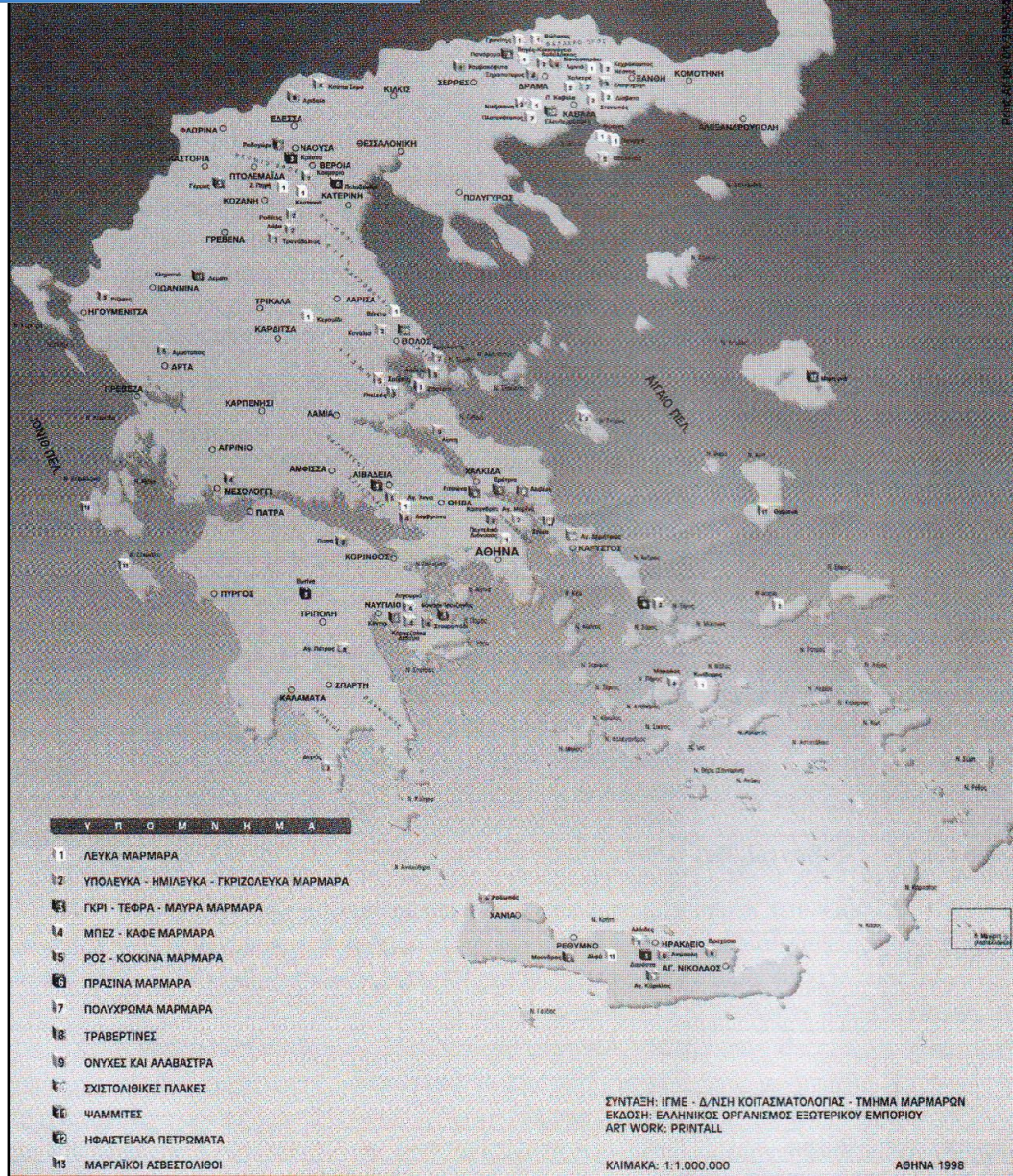
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[9].

/	
- ( >98% )	
	& & , - , - & , - - & , , , , & , , , & , & , , . , , . , , . ,
	(>92%) <sup>1</sup> (100%)- (100%)- (92%) (87%) (87%) (90%) (62%) (80%)
	- - , , ,
	, , , , ,
	.
	,
( )	.
	, .

**ΜΑΡΜΑΡΟΦΟΡΩΝ ΠΕΡΙΟΧΩΝ ΤΗΣ ΕΛΛΑΔΟΣ**



1.22:

[9].

2003

35 . p.

140

(14 . )

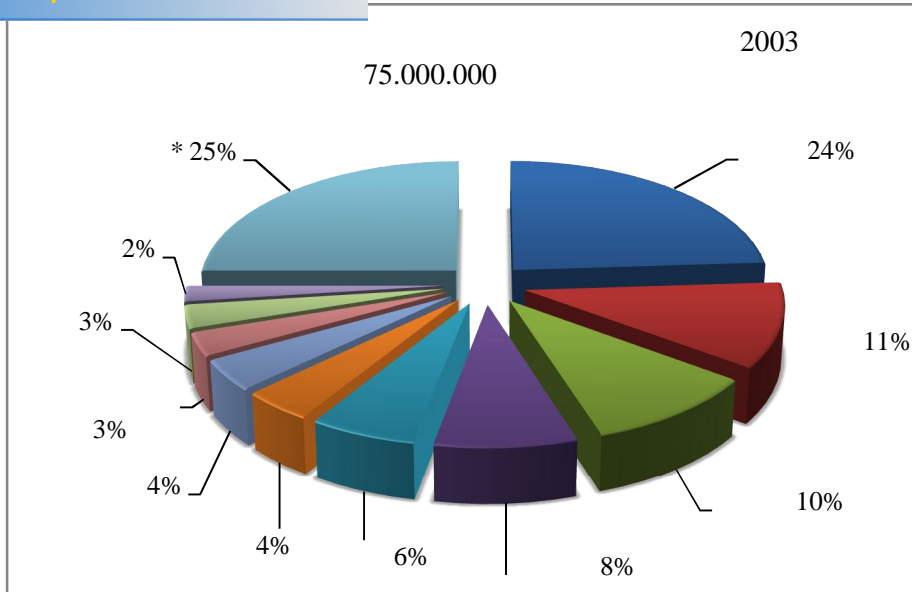
53%.

10

48%.

2003

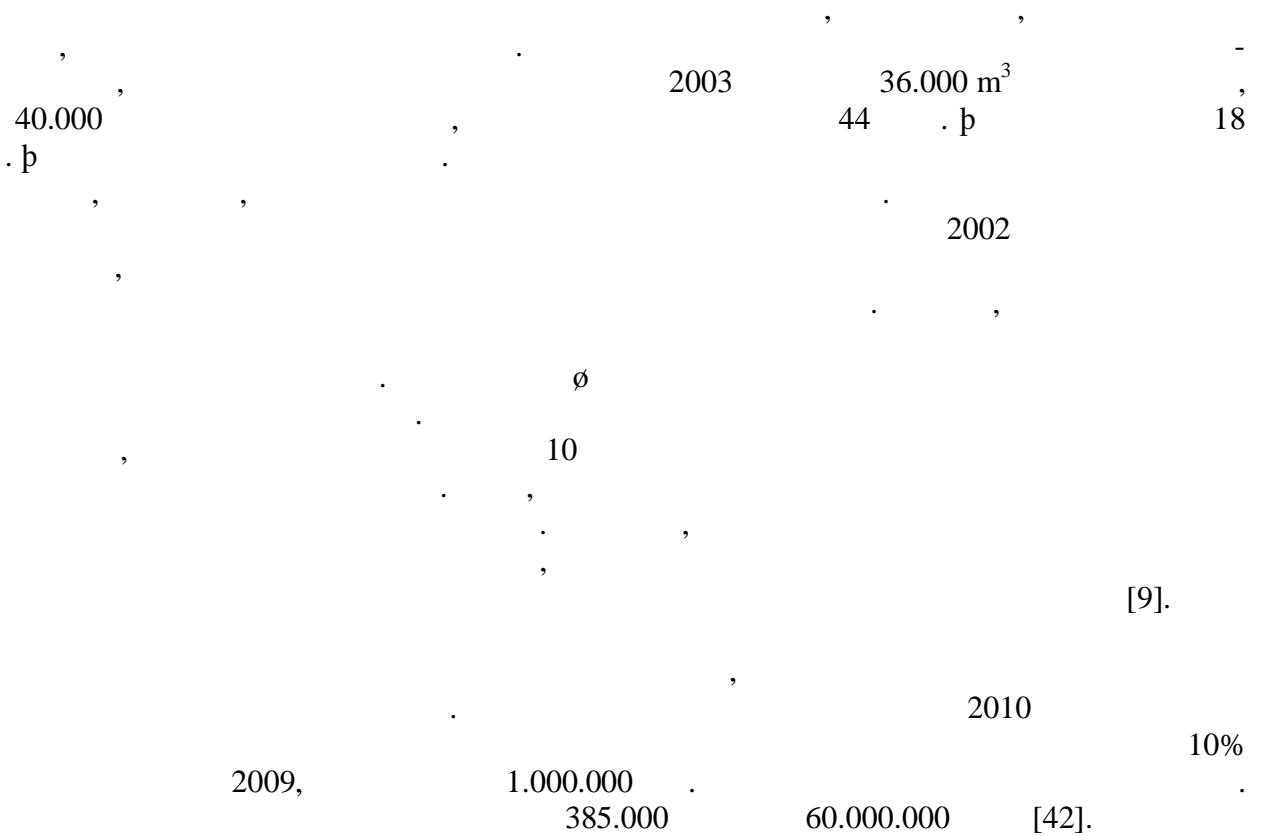
40% 41%,  
3-4%



\*

1.23:

(2003) [9].



1.2.6.1

[15].  
( , , , 30 )  
∅  
7 [9,36]. , 11 . 11  
( CaO) ( 1.6) [9,11,12,35].

1.7: [12].


SiO<sub>2</sub>  
SO<sub>2</sub> ( Ca(HSO<sub>3</sub>)<sub>2</sub> ) CaO  
MgO ( . , )  
( . . ) Ca

Ca

[11].

[12].

a. (>80%)

b. (>90%), « »

[11].

[12].

[11].

[9]. 1.7 1.8

**1.8:**

[11,12,28].

>1 m	
>30 cm	
1-30	
1-20 cm	
3-8 cm	
0,2-5 cm	
<4 cm	
<3 cm	
<0,2 mm	
<0,1 mm	



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- - 
  - MgO:
  - 
  -
- ), CaSO<sub>4</sub>, CaCl<sub>2</sub>, LiCl, LiBr.
- ( ).
- ) (
- ) [9].

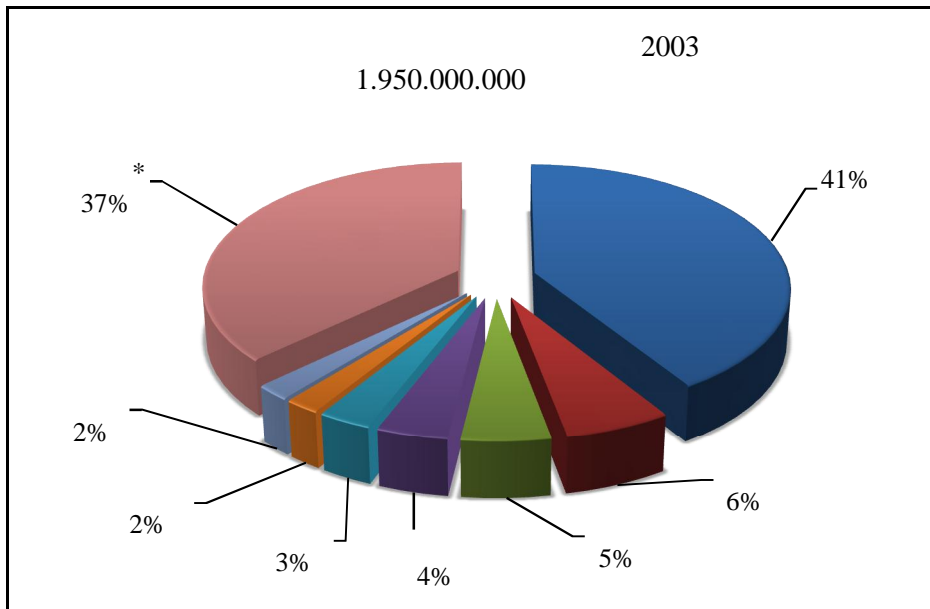
[9]. ( . . %)

			>30 cm							
			= 1-200 mm							
			<40 mm, >95% CaCO <sub>3</sub> , <1% SiO <sub>2</sub>							
			<5 mm, <5% MgO							
			>65% CaCO <sub>3</sub> , <4% MgO, <1,5% A. ., <0,1% F, <0,5% (P+Zn+Pb), <3% L.O.I.							
			= 0,2-2 mm, >60% CaCO <sub>3</sub> , 5-20% MgO							
			<30 mm, >97% CaCO <sub>3</sub> , <3% (SiO <sub>2</sub> +Al <sub>2</sub> O <sub>3</sub> +FeO+MnO), <0,02% P, <0,1% S							
			>98,5% CaCO <sub>3</sub> , <0,5% SiO <sub>2</sub>							
			<0,1 mm, >95% CaCO <sub>3</sub> , 2% SiO <sub>2</sub> , 1-2% MgO, 1% Al <sub>2</sub> O <sub>3</sub> , 1% Fe <sub>2</sub> O <sub>3</sub> , 0,02% MnO, 1000 ppm Cl							
			>98% CaCO <sub>3</sub> , é0 SiO <sub>2</sub> , é0 Al <sub>2</sub> O <sub>3</sub> , é0 (As+F+Hg+[b+H.M.])							
				CaO	MgO	Al <sub>2</sub> O <sub>3</sub>	Fe <sub>2</sub> O <sub>3</sub>	SO <sub>3</sub>	A.Y.	C
				%	%	%	%	%	%	%
➤	:	1-5 mm	>55	<0,8	<0,35	<0,08	<0,05	<0,6	<0,1	<0,05
➤	:	1-5 mm	>30	>21,5	<0,40	<0,25	<0,20	<0,6	<0,4	<0,05
			= 1-5 mm, <0,1% Fe <sub>2</sub> O <sub>3</sub> , <0,001 Cr <sub>2</sub> O <sub>3</sub> , <0,1%							
			<10 m, >95-97% CaCO <sub>3</sub> , >90%			<35 mg			<30 ml/100 g	
			<10 m, 98% CaCO <sub>3</sub> , <0,03% (Na <sub>2</sub> O+K <sub>2</sub> O), <0,02% MnO, <0,005% CuO, <0,2% L.O.I.							
			>97% CaCO <sub>3</sub> , <1,2% SiO <sub>2</sub> , <0,5% (Al <sub>2</sub> O <sub>3</sub> +Fe <sub>2</sub> O <sub>3</sub> ), <0,5% MgO, <0,004% P, S=							
			>98,8% CaCO <sub>3</sub> , <1% (Mg + ), <0,05% Fe, <0,002% H.M., 0,0005% F, <3 ppm As, <3 ppm Pb, <0,5 ppm Hg, <0,2% A.Y.							
. . =	, . . =		, L.O.I.=							

[12,28].

1.

CaO), (14,3% SiO<sub>2</sub>), (3% Al<sub>2</sub>O<sub>3</sub>) (1,6  
 (44,4%  
 (1,5% Fe<sub>2</sub>O<sub>3</sub>),  
 Ca-Al-Si clinker (3-5% Portland). clinker  
 Portland.  
 1.480 C  
 145 10  
 70% (41%) [9].



\*

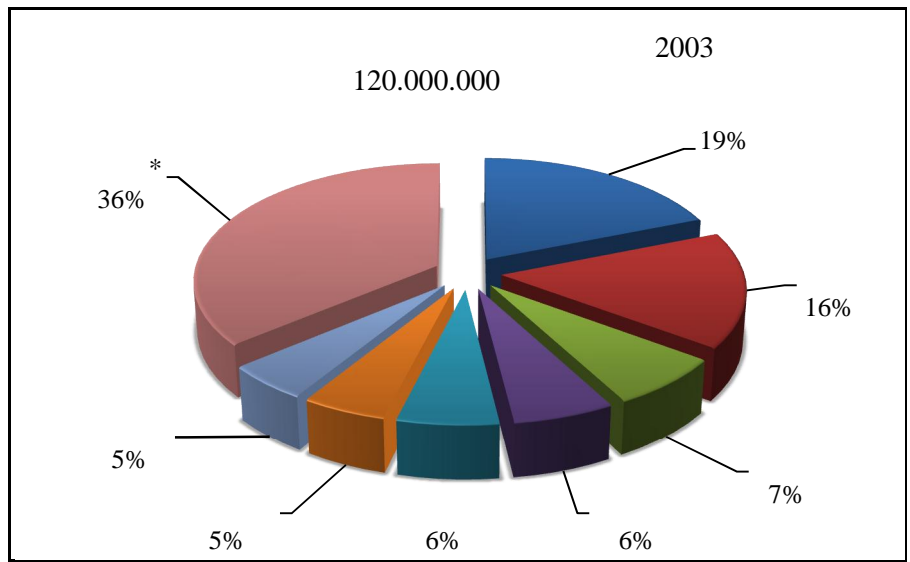
140

1.24:

(2003) [9].



2. , , [13]. , ( 20-70 mm) , ( 10 mm) , (CaO). . 65,5 kJ/mol. CaO ( , , ), [9,11]. 35% [9]. 70



\* , , , , , , , .  
**1.25:** (2003) [9].

3. , , , ( , , . . . ) , , . ( ) , . 4. , , , (

0% [13].

1

2

( 0,5%). 2.000-2.200 C

, 950 kg , 550 kg

(80% CaC<sub>2</sub> 13%

CaO). CaC<sub>2</sub> 1 [11,13].

6. ( , CaO ),

1,05 ,

1 (CaCl<sub>2</sub>), ,

97%, , 0,2% (CaO

, 800-1000 C) .

. [13].

[12].

, SO<sub>2</sub>, SO<sub>3</sub>, HCl HF ,

SO<sub>2</sub> ( , ), [35].

(pH)

100-8 mehs [13].

### 1.2.6.3

(PCC) (GCC)

(GCC),

GCC ,

(PCC) , CO<sub>2</sub> [Ca(OH)<sub>2</sub>]

[9,28].

CaCO<sub>3</sub> : 2,71,

2.711 kg/m<sup>3</sup>, 3, 0,2 cal/g<sup>o</sup>C, 23-

30x10<sup>-6</sup> cm<sup>o</sup>C, 1,59 [9].

✓

✓

✓

(<10 m -45 m ).

✓  
✓

[12].  
 (GCC)  
 , GCC ( >80%),  
 (>90%),  
 ( , ).  
 (PCC), ,



1.26: [42].

PCC. GCC  
 pH.  
 [9,35,47].  
 CaCO<sub>3</sub> 30%  
 [12].  
 [47]. [9].

(3-5 m) PVC, [12].  
GCC PCC (0,5-3 m),  
PCC [9,35].

**1.2.6.4**

[22,27].  
[7,33].  
[9,13].  
( g ), 2-4  
[12,33,35].  
MgO,  
1.000-1.100 °C ( 10 mm)  
(CaO.MgO).  
65,5 kJ/mol. [9]. CaO.MgO  
[9,11,12,28].  
[12]. MgO 21%  
[13].

U. , [11].

**1.2.6.5**

, , .

[13,29].

[9].

- 1) ( : )
- 2)
- 3) ( )
- 4)
- 5)
- 6) [13,31]

, , , .

[9].

, . ( , , . . . )

, « »

, ( . . . )

( . . . ) , . . . )

( ) , , . . . ).  
 , , , . . . [11].  
 ,  
 [47]. ,  
 ,  
 [11].



1.27: [37].

6.000.000 tn

[24-26].  $\text{CaCO}_3$

**1.2.6.6.1**

( ... , , ), ,  
 [11].

... ( )

, quartz

- - 
  - 
  - 
  - 
  - 
  - 
  - 
  - 
  -
- [12].

>98,6% CaCO<sub>3</sub> <1% SiO<sub>2</sub> [11].

**1.2.6.6.2**

( ... )  
 [13].

[11].

[12].



1.28:

[42].

### 1.2.6.6.3

), ( ) : (« » ), (« »  
 , <1,5% , <0,1% Fe <3% : >65% CaCO<sub>3</sub>, <5% MgO,  
 MgCO<sub>3</sub>. <5% MgO <3%  
 , Al<sub>2</sub>O<sub>3</sub>, SiO<sub>2</sub>, Fe<sub>2</sub>O<sub>3</sub>,  
 MgO. , CaO, SiO<sub>2</sub>,  
 Al<sub>2</sub>O<sub>3</sub> FeO.



Mn, H Al-Fe, Ca, Mg, P Mo Al,  
 pH  
 Ca Mg [11].  
 [12]. pH,  
 pH  
 0,2 mm 2-4 2 mm  
 6-12 MgO  
 CaCO<sub>3</sub>  
 75%  
 a. ( pH 3,5-6 6-7)  
 pH 3,5.  
 b. Ca Mg, Ca<sup>2+</sup> Mg<sup>2+</sup> Na<sup>+</sup>  
 35%

d.

e.

f.

g.

[11].

### 1.2.6.6.5

Ca Mg [9,11]. 13%  
0,2-1 mm. (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>, NH<sub>4</sub>NO<sub>3</sub>,  
[11].

### 1.2.6.6.6

( . . ) >98,5% CaCO<sub>3</sub> <0,5% SiO<sub>2</sub> [9,11].

### 1.2.6.6.7

Ca P,

98-98,5% CaCO<sub>3</sub>, (SiO<sub>2</sub>, Al<sub>2</sub>O<sub>3</sub> . . .)  
(As, F . .), Mg  
MgSO<sub>4</sub>·7H<sub>2</sub>O.

<3mm

- , Ø

1.2.6.6.8

∅

1.9

CaCO<sub>3</sub>

CaCO<sub>3</sub>

CaCO<sub>3</sub>

30%,

1.10: ( ) (%) [11].

	1975	1984	1968	1985
	1600	1200	88	62
	100	450	2	33
	365	450	10	5

CaCO<sub>3</sub>

30 98% <2 m,

CaCO<sub>3</sub>

(36%

1984),

( 1972 1985) 7%

95-97% CaCO<sub>3</sub> <2 m

Jensen SO<sub>2</sub>

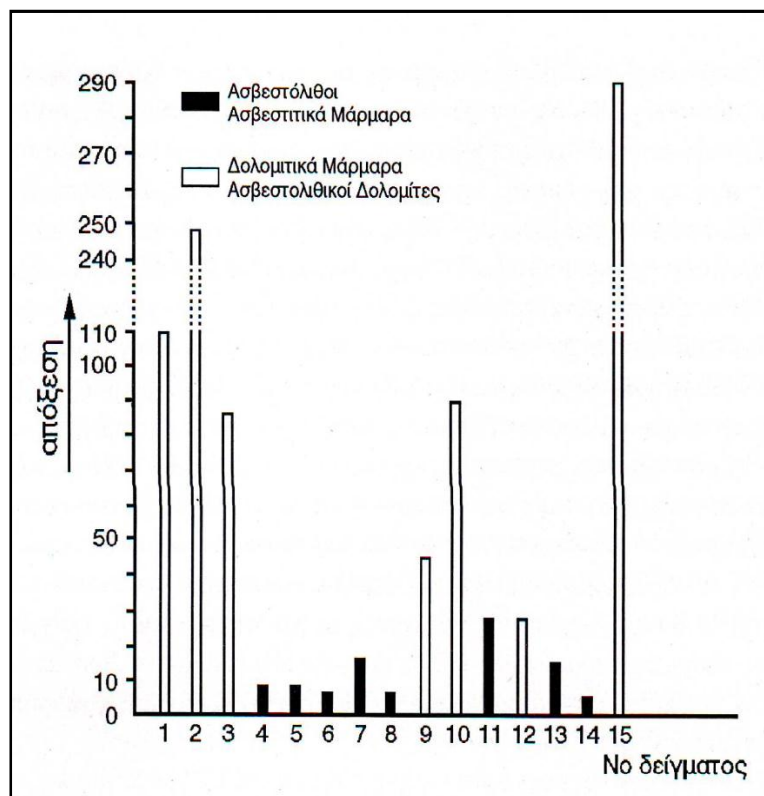
(0,1-0,2%)  
86% Elrepho R-457

<2 m. ,  
35 mg.

(1989),

∅

( 1.26).



:1. ,2. ,3. ,4. ,5. ,6. ,7.  
 ,8. ,9. ,10. ,11. ,12. ,13. ,14.  
 ,15. -

1.29:

[11].

### 1.2.6.6.9

∅

CaCO<sub>3</sub>.

Mg,

0,001 mm

0,1 mm

1.10.

, caulks

latex

[12].

1.11:

[12].

	( m)	( m)	
	22-40	420	
	12-22	100	
	3-10	44	
	0,7-2	10	

>80% ( 18-21, ) 1,5-4 m<sup>2</sup>/g, >96% ( 0,6-0,8 g/cm<sup>3</sup> ), pH=9-9,5.

[11].

180 C 7-9 bar,

1:13:0,7,

[9].

CaO

### 1.2.6.6.11

[11].

(CaCO<sub>3</sub>)

(Na<sub>2</sub>CO<sub>3</sub>) [9].

(

[12].

1.11.

<0,1% Fe<sub>2</sub>O<sub>3</sub>, <0,001 Cr<sub>2</sub>O<sub>3</sub> <0,1%

[11].

1.12:

[11].

	(%)	(%)
CaO	>54,85	>29,50
MgO	<0,8	>21,40
Fe <sub>2</sub> O <sub>3</sub>	<0,075	<0,25
Al <sub>2</sub> O <sub>3</sub>	<0,35	<0,40
	<0,05	<0,20
	<0,1	<0,40
	<0,6	<0,6
	<0,05	<0,10

30 kg CaO/t      60-65 kg CaO/t [9].

CaO

SiO<sub>2</sub> [11].

SO<sub>2</sub>, SO<sub>3</sub>, HCl      HF,

( . . ) [9].

: >95% CaCO<sub>3</sub>, <2%  
50%

SiO<sub>2</sub>, <1% Fe<sub>2</sub>O<sub>3</sub>, <1% Al<sub>2</sub>O<sub>3</sub>, <1% MgO, <0,02% MnO, <0,001% Cl,  
30      90%      90

[11].

[12].

### 1.2.6.6.13

[9].

( . . . . . )

3-8 cm,

85-95% CaCO<sub>3</sub>, <5% MgO

pH,

[11].

[15,24,27,63].

2.1

:

« »  
4000 kgm<sup>-3</sup> 63,5 200,6  
5,0 ( )

[64,69,71].

20

Pb, Cu, Cd, Cr, Fe . [26,15,27].

( )

[55,65].

[80].

Minamata,

ōItai-Itaiō

[27,80].

[15,25,71,80].

[83].

... Cd<sup>2+</sup>, Pb<sup>2+</sup>, Hg<sup>2+</sup>, Ag<sup>2+</sup> As<sup>2+</sup>,

[69,71].

( . . . ),

[65,83].

[69].



[83].

[71,83].

20

Zn

USEPA (US Environmental Protection Agency, 100

EPA).  
mg/day. A

[52,26,71,81,91].

Zn ,  
100 150 mg/day,  
Cu,

200-800 mg/day Zn

[91].

[64].

[64].

[71].

Cu,

Cu

$Cu(OH)^+$ ,  $Cu_2(OH)^{2+}_2$   
(

$CuCO_3$ .

$Cu^{2+}$

pH)

Cu [64].

Wilson,

[71,81, 94]

[84].

[64].

Agency, EPA)

USEPA (US Environmental Protection

[52,71].

Itai-Itai.

[71,76,81].

[27,56].

[85].  
(priority pollutants) USEPA (US

Environmental Protection Agency, EPA)  
[52,63,70].

[56,85].  
[63].

[63,25,71,83].

[71,81].

[52].

[71,83].

Cr(VI).

Cr(VI)

Cr(III).

Cr(VI)

: Cr (III)

[71,81].

: As(0)

s(V) arsenate, As(III) arsenite

As(III) arsine.

(As(III)

As(V)

(arsenate)

$3 \text{SO}_4^{2-}$ ,  $2 \text{SO}_4^-$ ,  $\text{SO}_4^{2-}$   $\text{SO}_4^{3-}$  [5783,97,].

Cr, Cu, As, Ni, Cd, Pb, V, Bi, Co, [25,26, 64,66,73]. Mn, Fe, Zn, Hg,

[61,65,66,70,71,80,99].

[88].

Zn,

[53, 91].

[64,94].

[25,27,70,94,88].

[27,58,76].

[88].

[89,90,92].

[83].

[92,99].

(Printed Circuit Board, PCB).

[85].

Ni, Zn, Cu, Cd, Pb (Cr

$\text{mg}\cdot\text{dm}^{-3}$ )

[89,90,92].

[57,60].

- 
- 
- 
- 
- 

[83].

[88].

[83].

[80].

[61].

[99].

## 2.3

### 2.3.1

[27,04, / 1].

[83].

[48]. 2.1 (Maximum Concentration Limits, MCL) Zn, As, Ni, Cd, Cu, Pb, Cr Hg  
 USEPA (US Environmental Protection Agency) [83].

2.1: (Maximum Concentration Limits, MCL)  
 USEPA (US Environmental Protection Agency) [83].

	MCL (mg/L)
	0,050
	0,01
	0,05
	0,25
	0,20
	0,80
	0,006
	0,00003

[19].

[15,71,83].

### 2.3.2

#### 2.3.2.1

[71,82].

DDTC, DMTC),  
triazine ( ),

(thiocarbonate salts, STC)

dithiocarbamate (DTC,  
2,4,6-trimercapto-1,3,5-

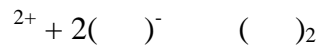
[82].

### 2.3.2.2

pH [71,82,83,69].

CaO, Ca(OH)<sub>2</sub>, Mg(OH)<sub>2</sub>, NaOH

NH<sub>4</sub>OH:



[82,83].

[82].

[71,83].

1000 mg/L.

[83].

pH 8-11 [71].

pH

( 2.1) [72,83].

pH

[54].

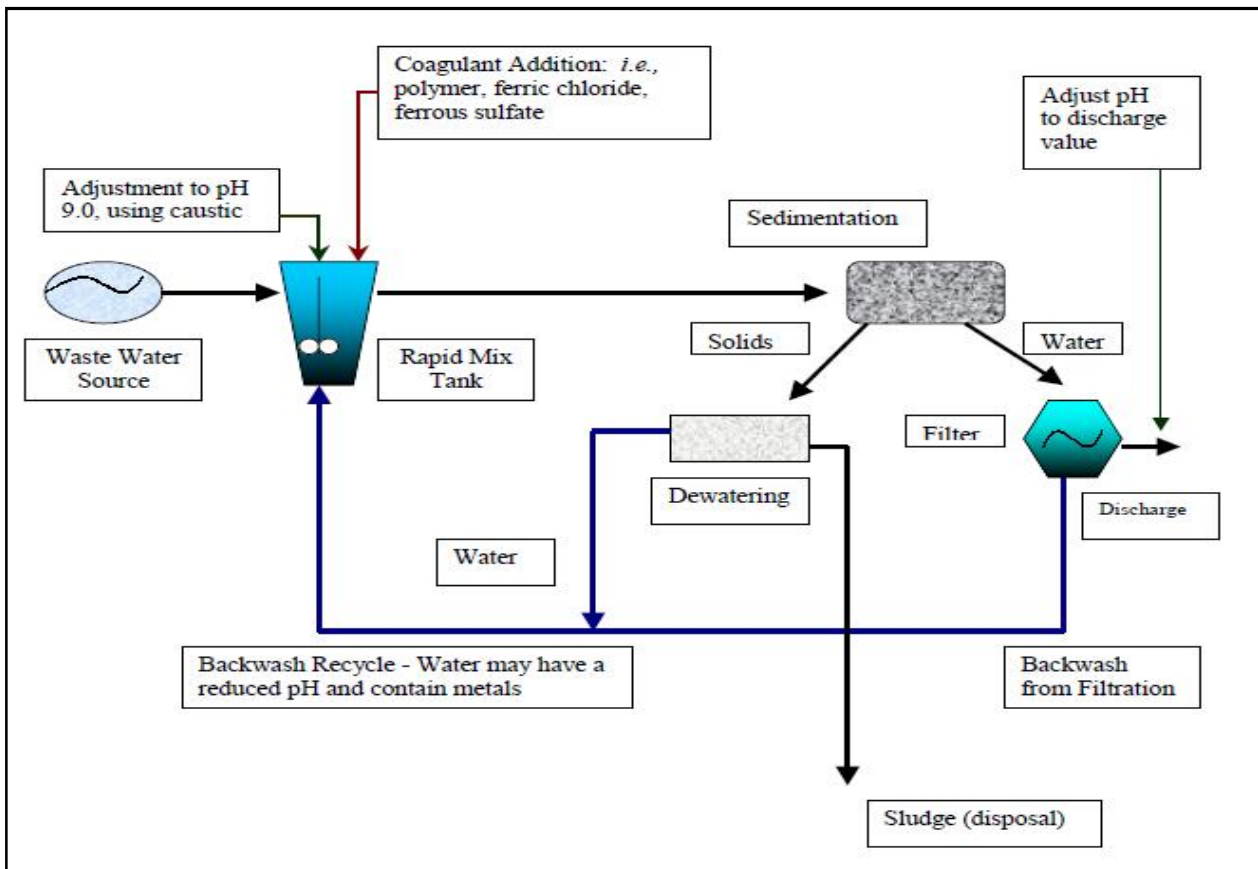
[71,72].

pH

[19,71,72,83].

[19,57].

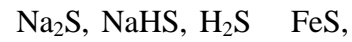
[19].



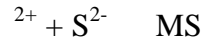
2.1:

[83].

[71].

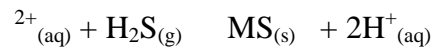


[82]:



pH

(Sulphate-Reducing Bacteria, SRB). SRB  
SRB

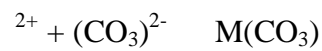


[71].

#### 2.3.2.4



[82]:



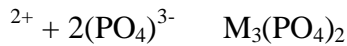
[61].

[83].



ρπ δ [01].

[82]:



### 2.3.2.5

[71].

[82].

### 2.3.2.6

potassium/sodiumthiocarbonate, sodiumdimethyldithiocarbamate), ( . . . trimercaptotriazine,

BDET<sup>2-</sup> (1,3-benzenediaminoethanethiol),

(dithiocarboxy)piperazine) (BDP) 1,3,5-hexahydrotriazinedithiocarbamate (HTDC), , N,N0-bis-  
-dipropyl dithiophosphate.

(potassium ethyl xanthate)

### 2.3.3

[71].

[83].

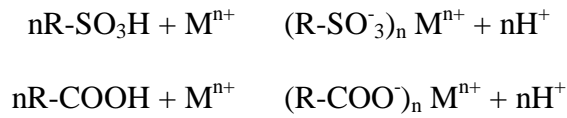
[99].

[83].

[71].

(-SO<sub>3</sub>H)

(-COOH).



pH,

[71].

EDTA

R-EDTA-Na.

(iminodiacetic)

Cu, Ni, Cd Zn. [99].

[71].  
 Ni, Cu, Zn, Cd, Pb)

(chabazite) [99].

(Cr,  
 Cs)

[71].

pH.

pH

pH,



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pH,

[99].

pH

[83].

[99].

### 2.3.4

#### 2.3.4.1

/  
[102,99].

/  
[83,88,100].

( )

[100-102, 99].

( )

( )

[101].

i.

ii.

iii.

[99].

Vaals,

( ).

Van der

2.2  
[88,100].

2.2:

[100].

	<p>der Vaals, Van</p> <p>4-40 kJ/mole</p>	<p>&gt;200 kJ/mole</p>

[25,71,83].

- 
- 
- 
- 

[25].

[75].

[71,96].

[83].

[71].

#### 2.3.4.2

(Activated Carbon, AC)

[71].

Ni, Cr (II), Cd(II), Zn(II), Cu(II), Cr(VI),

[81].

pH,

[67,75].

[71].

[87].

(Carbon Nanotubes, CNTs),

1991

Iijima,

- 1) (Single-Walled CNTs, SWCNTs)
- 2) (Multiple-Walled CNTs, MWCNTs).

HNO<sub>3</sub>, NaClO KMnO<sub>4</sub>.

[71].

#### 2.3.4.4

[63].

, silica gel,

silica,

Fe,

70

[15,61,71,74].

[74].

#### 2.3.4.4.1

[25].

[65].

[75,81,86].

Pb (II), Cd (II), Zn (II) Cu (II).

( )

[83].

[81].

( NaA zeolite, 4A zeolite

MMZ (Magnetiaccally Modified Zeolite))

pH

[83].

[68].

pH

[15,27,68,75,86].

[74,81].

[71].

900 C,

( )

[83].

[53,54].

[83].

2.3:

[83].

	(mg/g)					
	Pb <sup>2+</sup>	Cd <sup>2+</sup>	Zn <sup>2+</sup>	Cu <sup>2+</sup>	Cr <sup>6+</sup>	Ni <sup>2+</sup>
	1,6	2,4	0,5	1,64		0,4
	123					8
HCl-			63,2	83,3		
/ ( )	81,02		20,6	29,8		80,9
	85,6					
	155,0					
	4					
	398					

[53,61].

[78].

[23,50, 68,76,77].

[23,68].

[54].

CaCO<sub>3</sub>,

[23,57,59,77,98].



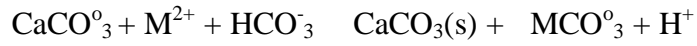
[50,51].

Ca

( )

[50].

[25,26]:



CaCO<sub>3</sub>

Cd, Co, Zn, Pb,

[57].

Cd,

Cu Zn

CdCO<sub>3</sub> ZnCO<sub>3</sub> Ca  
CuCO<sub>3</sub>  
Cd Zn

Cd Zn Cu ( )  
CdCO<sub>3</sub> CaCO<sub>3</sub>, Zn Cu

[57,76,77,79].

Cu

Zn

Cd, Zn, Mn Co

[49,57,59,77].

MeCO<sub>3</sub> ,

[23].

Fe Al

[54].

✓ :

Mg

:

( . . )

Ca

[15].

: Cd>Mn>Zn>Co

[49,50,51,23].

✓

:

[15].

Cu(II), Pb(II), Cd(II), As(V)

[15,20,27,48,97].

[22,19]. , ,

Pb(II), Cd(II),

Co(II) Cr(III) [78].

I.

II.

[15].

Ca

Mg

pH,

[15,20].

[19,98].

CaO

Pb

[15].

[63,64].

(Acid Mine Drainage, AMD) ( . . )

[15,60,62].

Ca<sup>2+</sup>,

Pb, Cd, As, Zn, Cu,  
[62].

(Powdered Marble Waste, PMW)

. H

Cu( ),

Cr( ), n( ) Pb( ) [24-26].

(FeCO<sub>3</sub>) . . ),

( . . ,

(MgCO<sub>3</sub>),

Mg

(CaCO<sub>3</sub>)

[23].

Cd(II), Pb(II), Cr(VI) [75,86].

#### 2.3.4.4.2

, o clinopyrrhotite,

[71].

[75].

[83,81].

ferrosorp plus,

akaganeite [83].

#### 2.3.4.4.3

( )

[71].

[61].

/

[71,92,95].

1. -
- 2.
3. [71].

[83,80].

[95].

( - )

(

),

, ,

[83].

[95].

CSC),

(Coconut Shell Charcoal,

[83,80,92,95].

[80].

[71].

Chatoetomorpha linum, Caulerpa lentillifera, Cladophroa fascicularis.

Oedogonium sp.

Nostoc sp.

Spirogyra, Ecklonia maxima, Ulva lactuca,

Fucus serratus [71,83].

Bacillus cereus, Escherichia coli, Pseudomonas aeruginosa, . . .

Aspergillus niger, Rhizopus arrhizus, Rhizopus oryzae, Saccharomyces cerevisiae, Lentinus edodes, . . . [71].

Ca-alginate

gel,

[93].

pH 2-6 [83]. H

[83].

) [61].

[95].

2.4

[83].

[71,92].

2.4:

[83].

	(mg/g)					
	Pb <sup>2+</sup>	Cd <sup>2+</sup>	Zn <sup>2+</sup>	Cu <sup>2+</sup>	Cr <sup>6+</sup>	Ni <sup>2+</sup>
-	456	493,7	495,9		3,65	158
Spirogyra ( )		2,0	13,9	31,7	0,79	
Ecklonia maxima-	235			133	23,4	
Ulva lactuca				90	112,3	
Oedogonium	145					
Nostoc	93,5					
Bacillus-	467	85,3	418	381	39,9	

2.3.4.4.4

(hydrogels)

ppb,

)

[83].

	(mg/g)					
	Pb <sup>2+</sup>	Cd <sup>2+</sup>	Zn <sup>2+</sup>	Cu <sup>2+</sup>	Cr <sup>6+</sup>	As <sup>5+</sup>
Gel /	433	150		164 135 200		230

- a) :
- b) (gels)

pK<sub>a</sub> 6,2 7. C-3,

sol-gel.

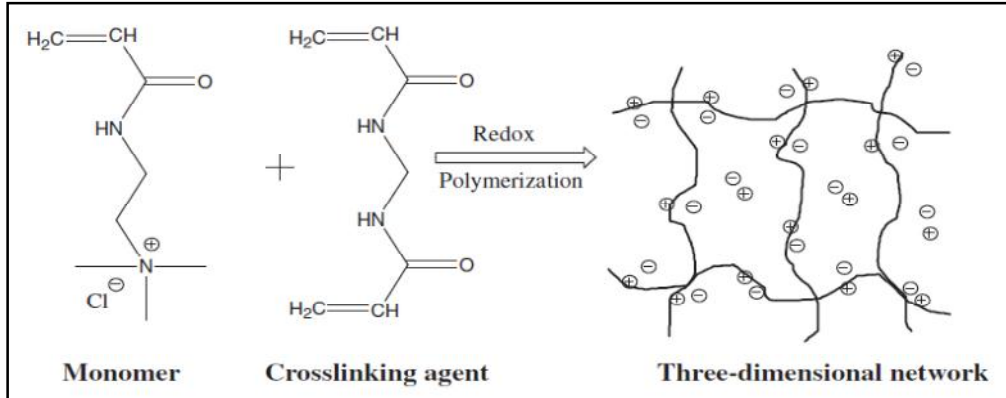
acrylamide), poly(ethyleneglycol dimethacrylate-co-poly(3-vinylpyrrolidone-co-methylacrylate)

2.2

loride.

pH >6.

[83].



2.2:

[83].

### 2.3.5

[83].

[71].

«membrane contactors»,

[83].

#### 2.3.5.1

##### 2.3.5.1.1

(Ultrafiltration, UF)

[71].

5-20 nm

1000-100.000 Da.



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[83].

[71].

(Transmembrane Pressure, TMP)

[83].

Ultrafiltration, MEUF)  
Ultrafiltration, PEUF).

, EYMM (Micellar Enhanced  
, EYM (Polymer Enhanced

**2.3.5.1.2**

Scamehorn

1980

Micelle Concentration, CMC),

(Critical

Dodecyl Sulfate),

SDS (Sodium

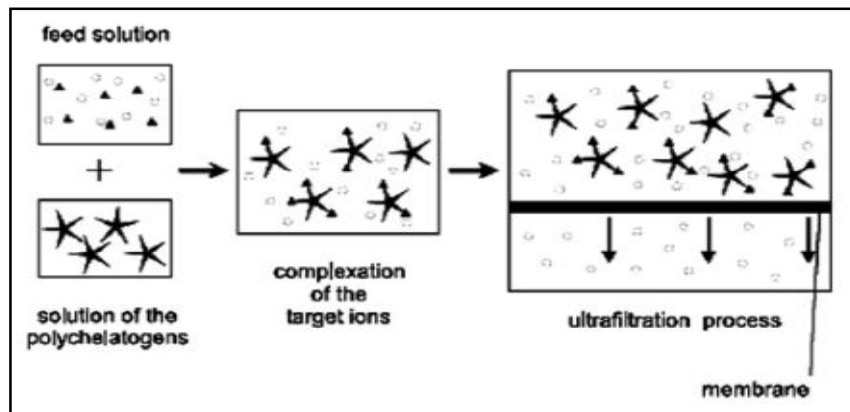
, pH



[71,83].

(PAA),

(PEI),



2.3:

[83].

pH

[71].

### 2.3.5.2

(Reverse Osmosis, RO)

20%

2.3.5.3

(Nanofiltration, NF)

[71].

2.3.6

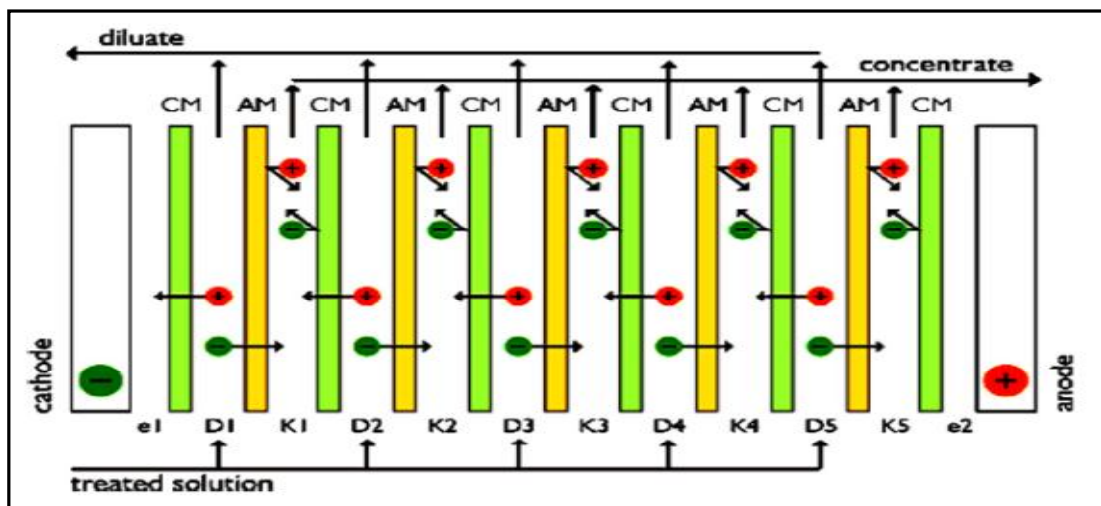
(Electrodialysis, ED)

[71].

[83].

[71].

2.4



2.4: Schematic diagram of an Electrodialysis (ED) stack. CM-Cation-exchange membrane/  
 D-Diluate chamber/ e1 e2-electrode chambers/  
 -anion-exchange membrane/ -concentrate chamber/  
 [83].

Cr Cu,

[83].

### 2.3.7

(PAM), PAC, PFS (Polyferric sulfate)

mercaptoacetyl chitosan, Konjac-graft-poly(acrylamide)-co-sodium xanthate poly-ampholyte chitosan ( -carboxyethylated chitosans),



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### 2.3.8

(Dissolved Air Flotation, DAF),

1990.

( , , . . . )

### 2.3.9

#### 2.3.9.1

(Electrocoagulation, EC)

in situ

**2.3.9.2**

(Electroflotation, EF) /

[71].

**2.3.9.3**

[71,83].

**2.3.10**

(e<sup>-</sup>/h<sup>+</sup>)

TiO<sub>2</sub>, ZnO, CeO<sub>2</sub>, CdS, ZnS . . .

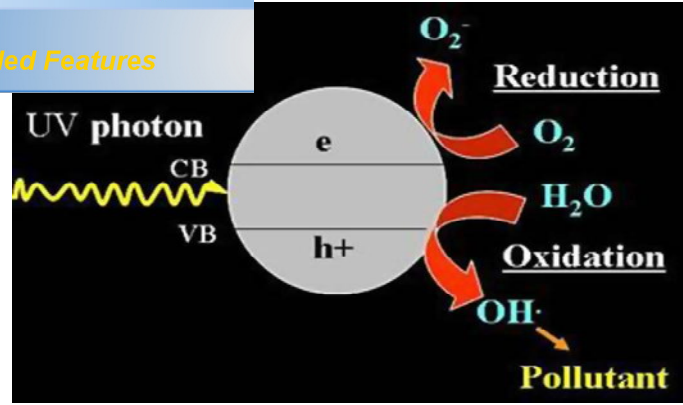
2.5

Cr(VI)

Cu(II)

(arsenite)

TiO<sub>2</sub>



2.5:

TiO<sub>2</sub> [83].

( ppm) [83].

2.4

[83].

[71].

0,1 3 mg/l

[57].

[71].

[83].

pH,

[71,83].

2.6:

[83].

	,	,
,	, pH,	,
,	,	,
,	,	,

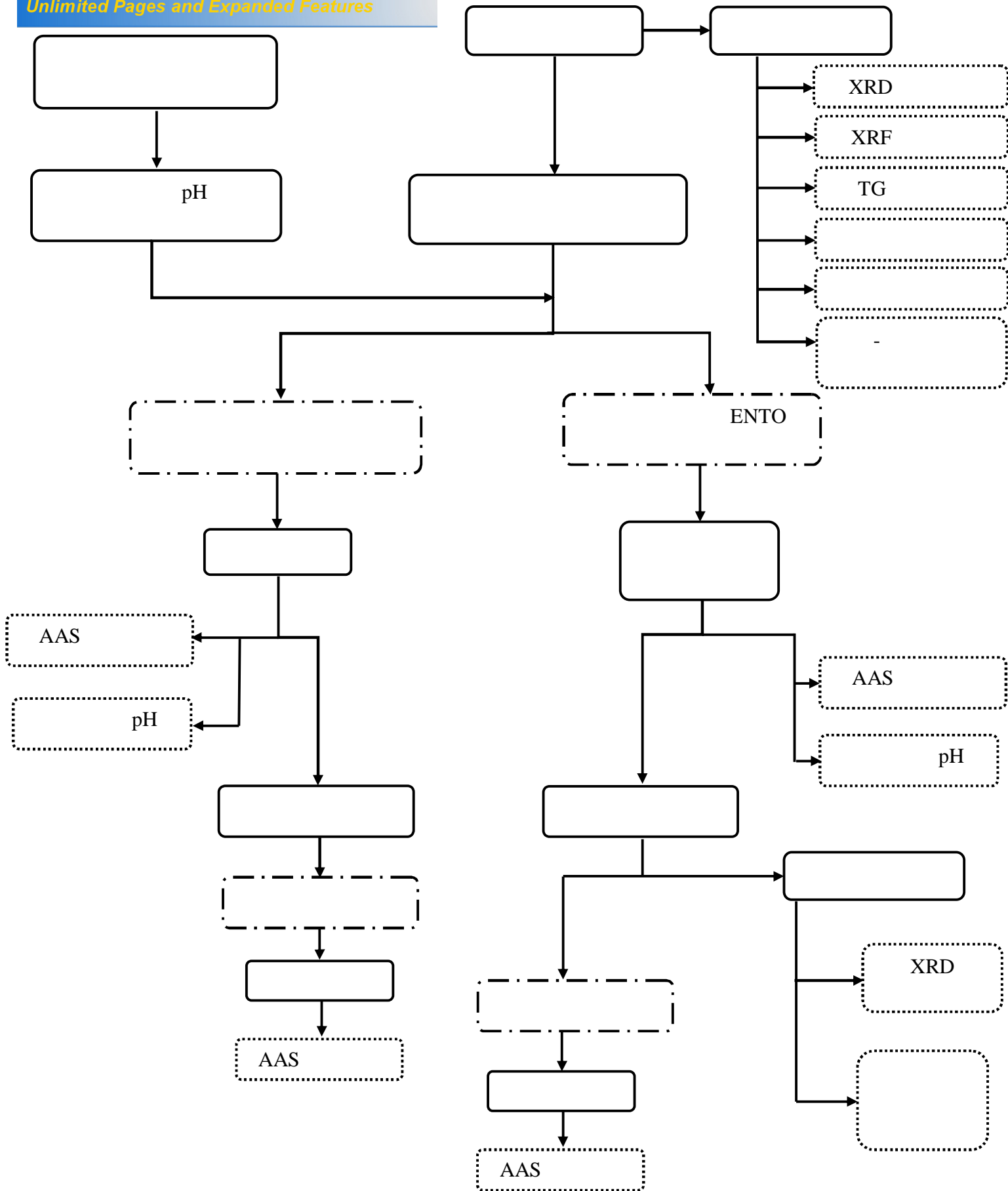




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### 3.1

#### 3.1.1

Zn, Cu, Pb Cd,

mg/L 100 mg/L 5

Na HCl 2% c(HNO<sub>3</sub>), pH 5. Zn, Cu, Pb Cd T 1000 mg/L,

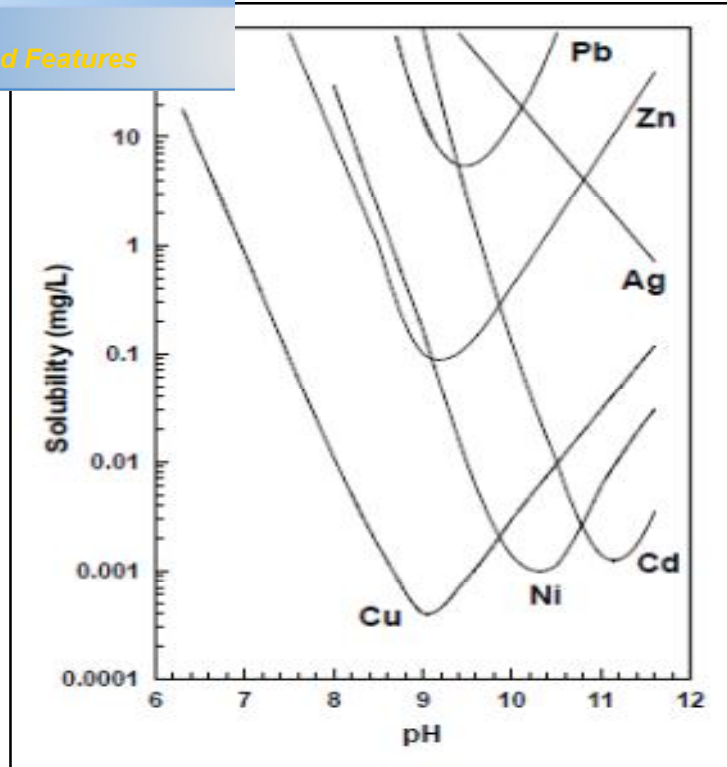
pH<6. pH,

H<sup>+</sup> H<sup>+</sup> pH >5

pH pH>6. pH>7

pH=5

( 3.1) [15,27].



3.1:

pH [110].

pH : 30 ml Pb, Cu, Zn Cd,  
pH 5  
pH 8.  
pH 7  
pH 8 pH ( 3.2).



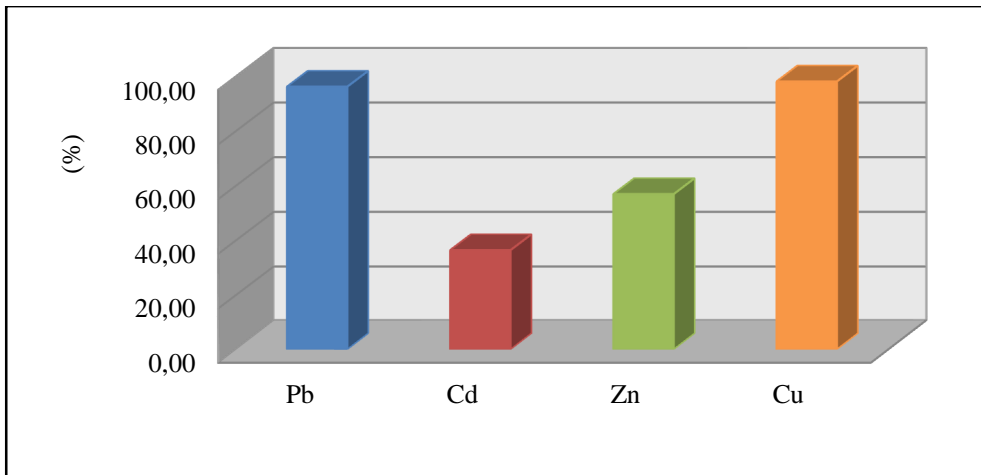
3.2:

Pb, Cd, Zn Cu pH=8.

pH

3.1: (%) Pb, Cd, Zn Cu pH.

	(%)
Pb	96,30
Cd	36,50
Zn	57,00
Cu	98,21



3.3: (%) Pb, Cd, Zn Cu pH.

Cu , 96,30% 98,21%, Pb ,  
 36,50% 57% ,

(MgO>20%)

( . . . )

3.2 .

3.2:

<90 m 90-315 m 315-1000 m 1000-4000 m	315-1000 m	<1 mm (<90 m:28,7% 90-315 m: 28,5% 315-1000 m: 42,8%)	<1 mm (<90 m:19,4% 90-315 m: 16,2% 315-1000 m: 64,4%)	<1 mm (<90 m: 29,7% 90-315 m: 40,1% 315-1000 m: 30,2%)

pH

50 ml

pH

400 rpm

pH  
5 g

20 min

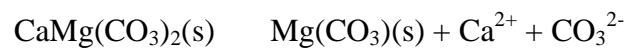
3.3.

	pH
	10,32
	9,15
	10,06

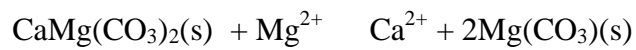
pH.  
/

pH

[109]:



:



[19]:



3.2.1 (Atomic Absorption Spectrometry, S)

3.2.1.1

(AAS) (single-element method),  
 1955 WALSH  
 S [112]. [115].

- [MeX] MeX (1)
- MeX Me + X (2)
- Me Me\* (3)
- Me Me<sup>+</sup> + e<sup>-</sup> (4)

Me = [ eX].  
 (1) (2).  
 (3),  
 (4).  
 AAS (1) (2),

AAS

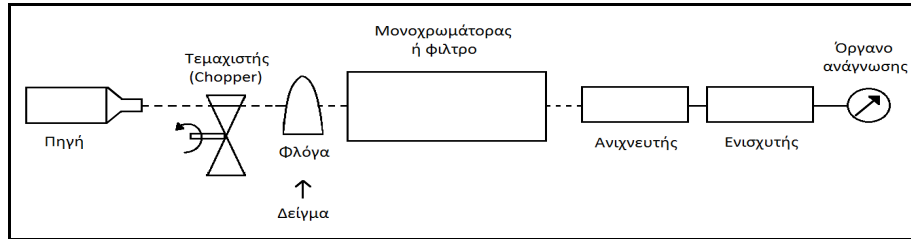
3.4  
 (AAS) [112].

[113]. AAS Hg, Pb Cs.

( [112]. )



[113].



3.4:

[112].

(Chopper),  
 $\emptyset$   
 (ac)

AAS

LAMBERT-BEER ( )

[112].

$$\log \frac{I_0}{I} = A = \epsilon c b$$

$\log \frac{I_0}{I} = A$   
 :  
 :  
 :  
 b:  
 c:

%

$$A = \log(I_0/I)$$

BEER,  
 % [112].

S 70 , 0,01-10 ppm.

5% ø . .

0,5-2%. S , , : Zr, Hf, Ta, Re, U  
H, N, O.

AAS :

- 1) , ,
- 2) , ,
- 3) , ,
- 4) , ,
- 5) , ,
- 6) , ,
- 7) , ,
- 8) [112] , ,
- 9) [115] , ,

AAS

[115]

[112,113]. , ,

AAS

[113].

**3.2.1.2**

( , . .) [11].

Mg, Zn Fe.

Perkin Elmer 3300.

2.300 C.

1. 0,5 g : 200
2. 20 ml HCl 2 1:1
3. 50 C, . .

250 ml

S.

- a)
- b)
- c)
- d)
- e)
- f)
- g)

3.2.1.3

%  
:

$$\% = \frac{\cdot \square}{\square} \cdot 100$$

: = mg/ml  
= mg  
V =

3.4:

	%		
	Fe	Mg	Zn
	0,020	11,6	0,016
	0,021	12,2	0,005

Fe Zn  
Mg  
( Mg)

(X-Ray Diffraction, XRD)

3.2.2.1

ó

ø

XRD

(3-D)

XRD,

XRD

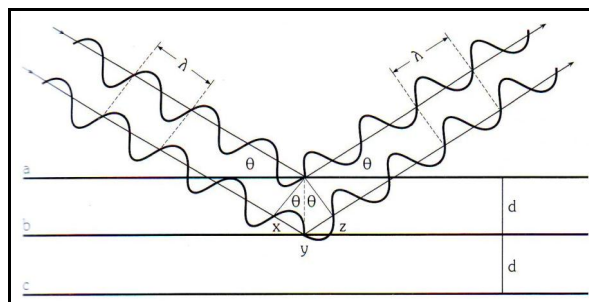
( 1912 von Laue)

d

Bragg (1912):

$$n \cdot \lambda = 2d \cdot \sin \theta$$

d



3.5:

[112].

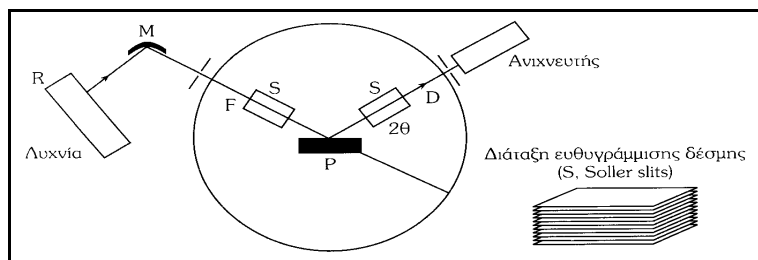
- 1) (Powder Diffraction or Polycrystalline Diffraction).
- 2) H (Single Crystal Diffraction).

$0,5 \cdot 10^{-2} - 0,5 \cdot 10^{-3}$  mm

Rietveld.

0,1-0,6 mm

(Powder X-Ray Diffraction)



3.6 :

Bragg-Brentano [112].

3.6  
(Powder Diffractometer)  
(R)

Bragg-Brentano.

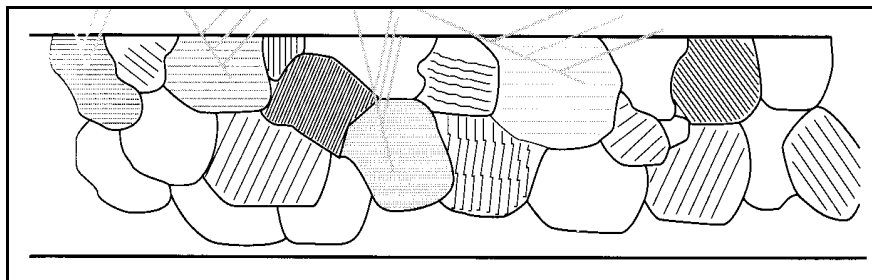
(Cu)

(CuK $\alpha$ ),  
(S=Soller slits)

( $\lambda$ ).  
25 m)

(F)

( $\theta$ ).  
Bragg,  
3.7).



3.7:

(Powder diffraction) [112].

(SSD), . . . (D),  
Peltier [112]. (CCD)

*XRD ó* *XRD*  
( 2 ) 2  
Bragg d  
[116].  
ø (Cps) 2  
PDF (Powder  
Diffraction Files) , JCPDS (Joint Committee on  
Powder Diffraction Standards).

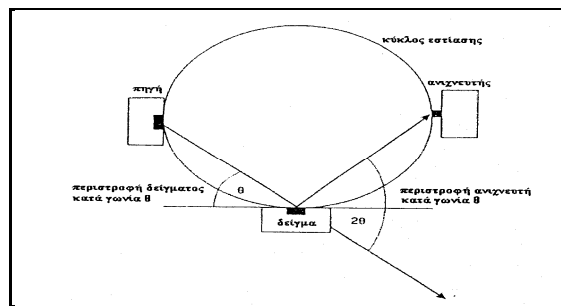
2 ,  
[112].  
1%,  
0,1%,  
XRD [113].

*XRD*  
(XRD)  
:  
• ( ) ø  
•  
• ó  
•  
•

✓  
 ✓  
 ✓  
 ✓  
 ✓  
 ✓  
 ✓  
 ✓  
 ✓  
 ✓  
 ✓  
 ✓  
 ✓  
 ✓  
 ✓  
 [112]

3.2.2.2

Siemens, D5000 :  
 • ό ( 2 W)  
 • ( 30 mA 40KV  
 • 2 , (1) (0,1  
 mm) , ( 3.8)  
 • ,



3.8: [111].

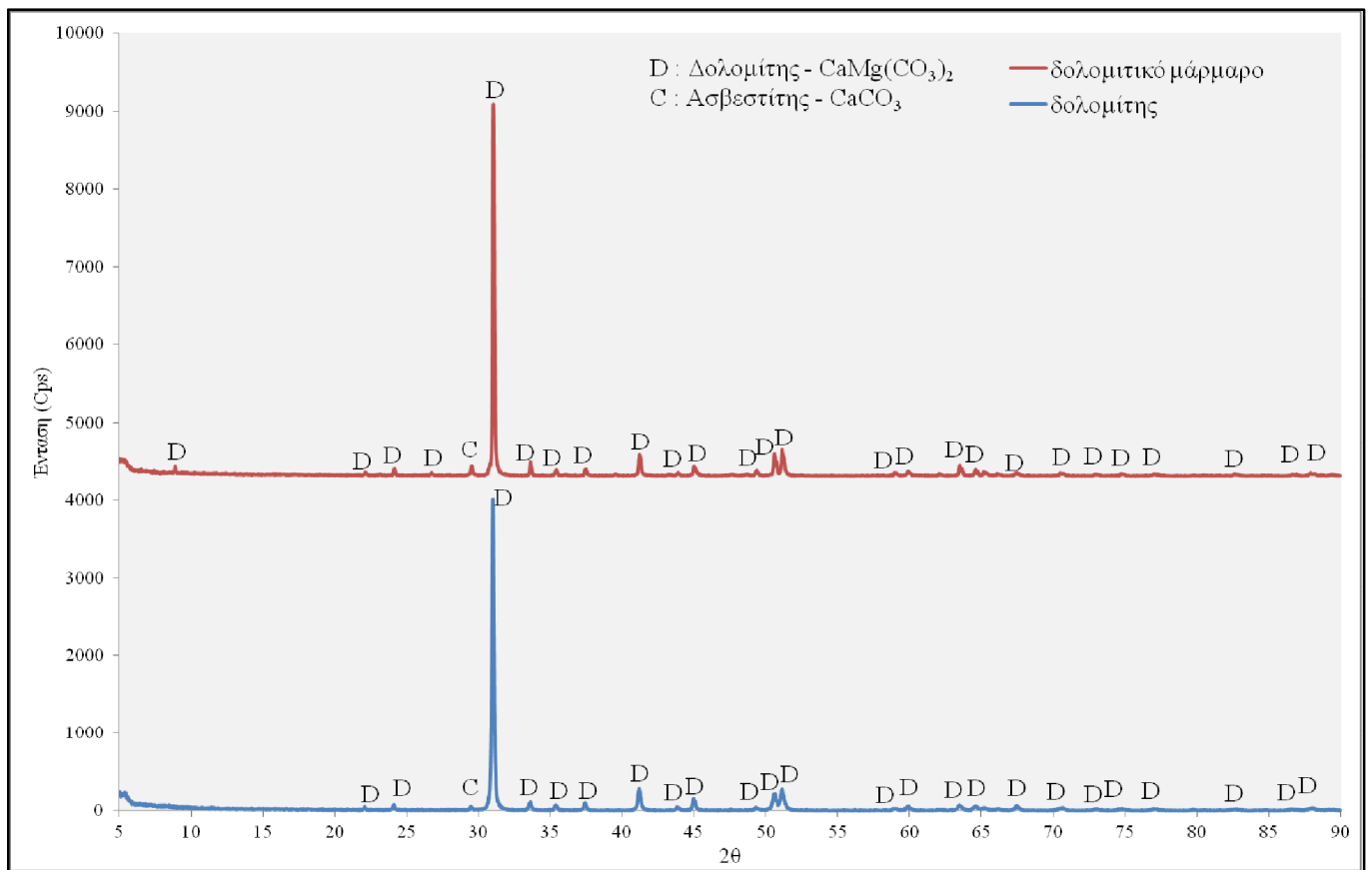
DIFFRAC. AT. Search Program

SIEMENS.

- i.
- ii.
- iii.
- iv. ( 2 : 5-90 , step: 0.020°, step time:2 s, Temp: 25°C) .
- v.
- vi.

Software [113].

3.2.2.3



3.9:



PDF  
2 =31 .

2 =29

2

3.5 .

3.5:

2

2 ( )		
[103-108]		
9	8,9	8,88
22	22,08	22,12
24	23,64	23,64
31	31,02	31,04
33	33,62	33,62
35	35,38	35,44
37	37,44	37,5
41	41,2	41,22
44	43,84	43,9
45	44,98	45
49	49,32	49,38
50,5	50,6	50,6
51	51,14	51,18
59	58,96	58,96
60	59,92	59,92
63	63,48	63,52
65	65,2	65,18
67	67,48	67,5
71	70,58	70,52
73	72,94	72,9
75	74,68	74,74
77	77	77,02
83	82,68	82,68
87	86,66	86,69
88	88	87,94
29	29,46	29,54

\*

( -Ray Fluorescence, RF)

3.2.3.1

ó

,

,

( )

,

( )

.

,

,

,

( ),

,

,

,

.

«

»

,

(WD-XRF, Wavelength-Dispersive XRF)

Bragg

( )

(ED-XRF, Energy-Dispersive

XRF),

.

.

( )

3.10 .

:

1.

2.

3.

(collimators),

4.

-

5.

.

,

,

.

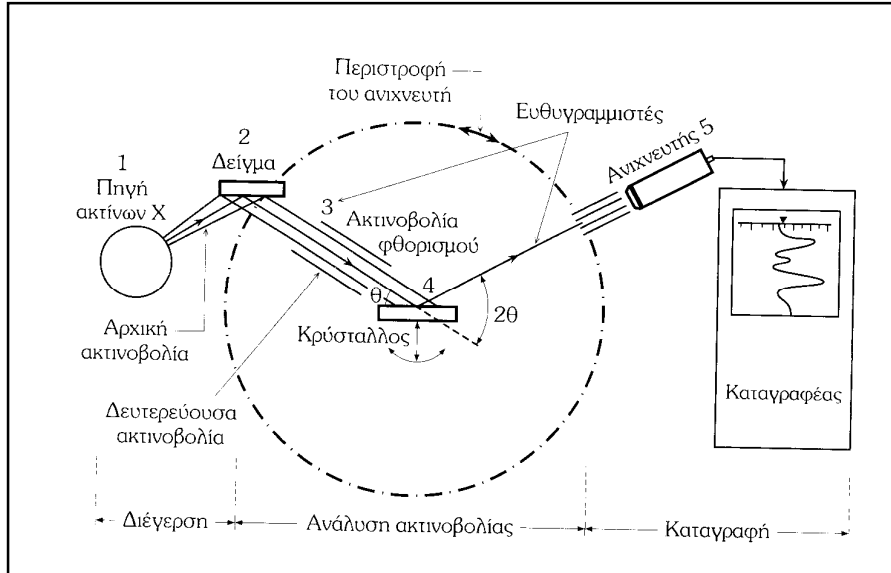
,

,

.

,

(gearless) [112].



3.10:

WD-XRF

( ) [112].

XRF

XRF,  
Be, C, N, O F

Be U.  
11(Na)<Z<20

×20 (Ca)

XRF

2 . 10 140  
Bragg

D

L

2

( . . : =2:1).

, = /t,

t

∅

2

∅

$$C_x = \frac{N_x}{N_s} \cdot 100\%$$

$$C_x = \frac{A_x}{A_s} \cdot 100\%$$

$10^{-1}$   $10^{-4}$  %

( 2-5%),

10

ppm.

**XRF**

H

[7]. XRF ( , , , , )

100% 100 ppm. , XRF

U,

0,01 ppm,

(100-1000 ppm).

WD-XRF

∅

ED-XRF

0,1-10 ppm

, . . Fe,

1-5% (B, Be),

ED-XRF

5-10

**XRF**

=20-55 70-90.

XRF

—

—

—

—

—

✓

( . . ).

[112].

(ICP-AES), ( S) XRF

83

X.

(He) [117].

### 3.2.3.2

[11].

40 m [112].

(Philips 1606).

### 3.2.3.3

3.6:

XRF

/	(%)
CaO	31,9
MgO	20,3
SiO <sub>2</sub>	0,269
Fe	590 ppm
Al	290 ppm
Sr	170 ppm
K	77 ppm
S	65 ppm
P	50 ppm
Cu	10 ppm
Cl	53 ppm

XRF

	(%)
CaO	33,05
MgO	19,3
SiO <sub>2</sub>	0,524
Sr	310 ppm
Al	200 ppm
Fe	200 ppm
Cu	89 ppm
P	33 ppm

3.6 3.7

MgCO<sub>3</sub>, MgO, Ca, 19,3%

33,05%, 31,9%

SiO<sub>2</sub>, Sr, Fe, Cu, P Al, ppm.

Cl, K, S,

### 3.2.4.1

( ) (TG) (GA) [114].  
 [115].  
 TG  
 (Thermobalance).  
 (TG) (TGA)  
 1915 . Honda.  
 ( , , . . . ), [114].  
 H TG , [113]. TG  
 [114]. TG  
 ( , , . . . ) TG  
 ( , ), ( , )  
 TG : , [113].  
 TG (instrumental control)/ (data handling).  
 TG TG  
 TG DTG.  
 DTG TG

DTG ( ),

- ASTM :
- I. DTG
  - II.
  - III.
  - IV.

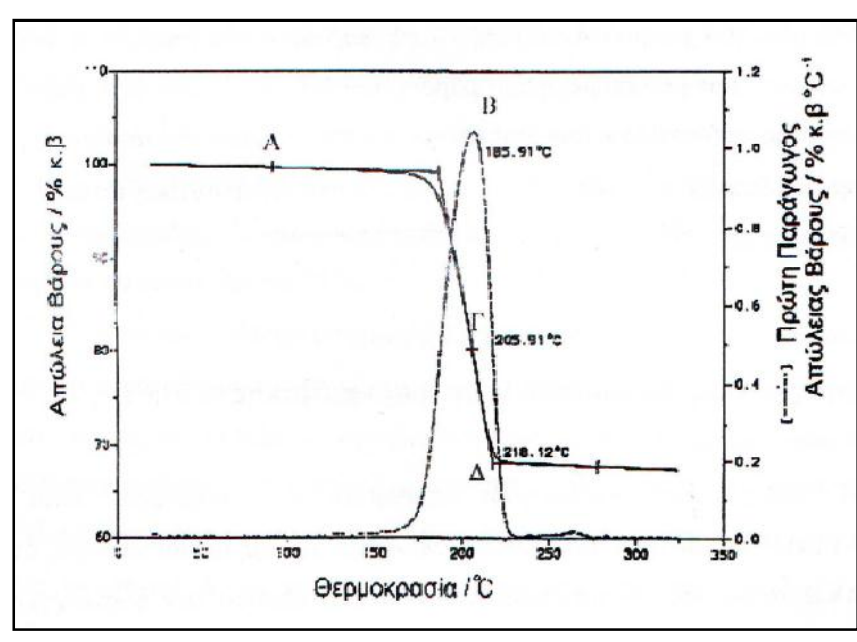
II IV III

✓ IV

ASTM ISO

TG

TG [114].

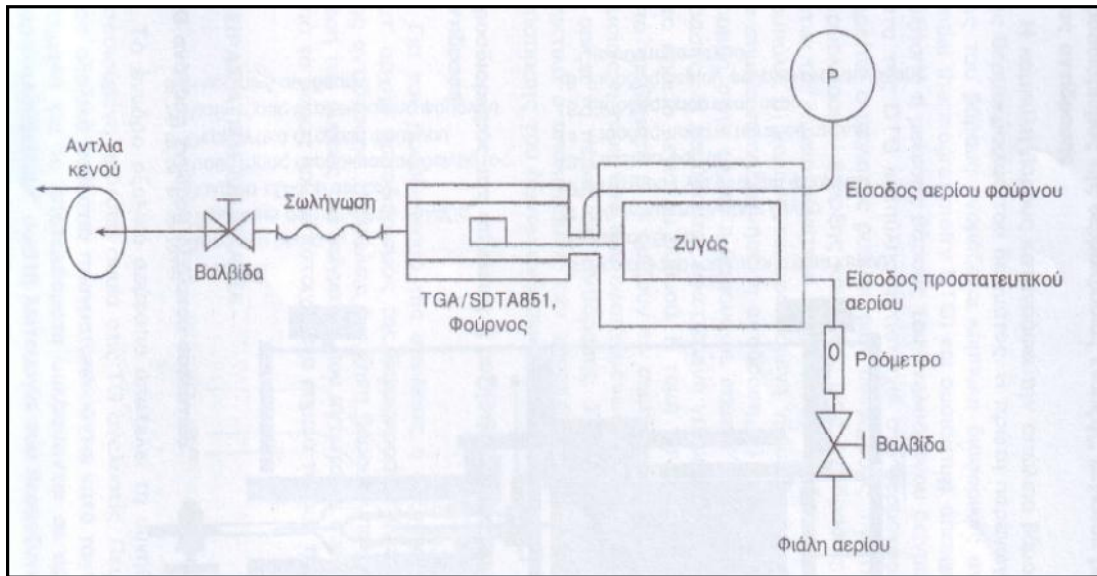


3.11: TG DTG K<sub>2</sub>CO<sub>3</sub>- A: TG, : [114].



TGA/SDTA 851

Mettler Toledo,



3.12:

[113].

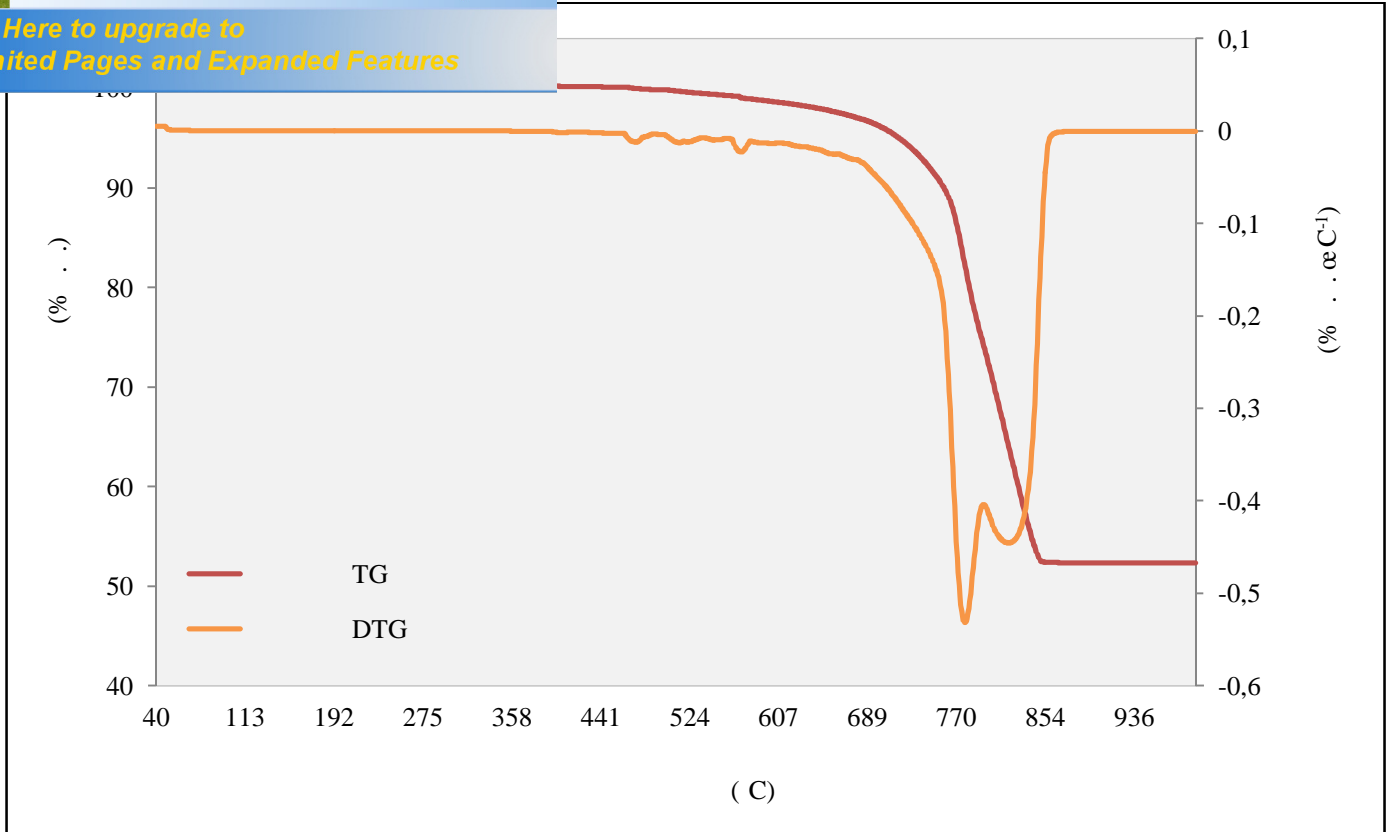
N<sub>2</sub>

TG DTG,

3.2.4.3

DTG

TG



3.13: TG DTG

TG =850 C. 47,7%.

DTG =724 C.

[19,103-105,108]

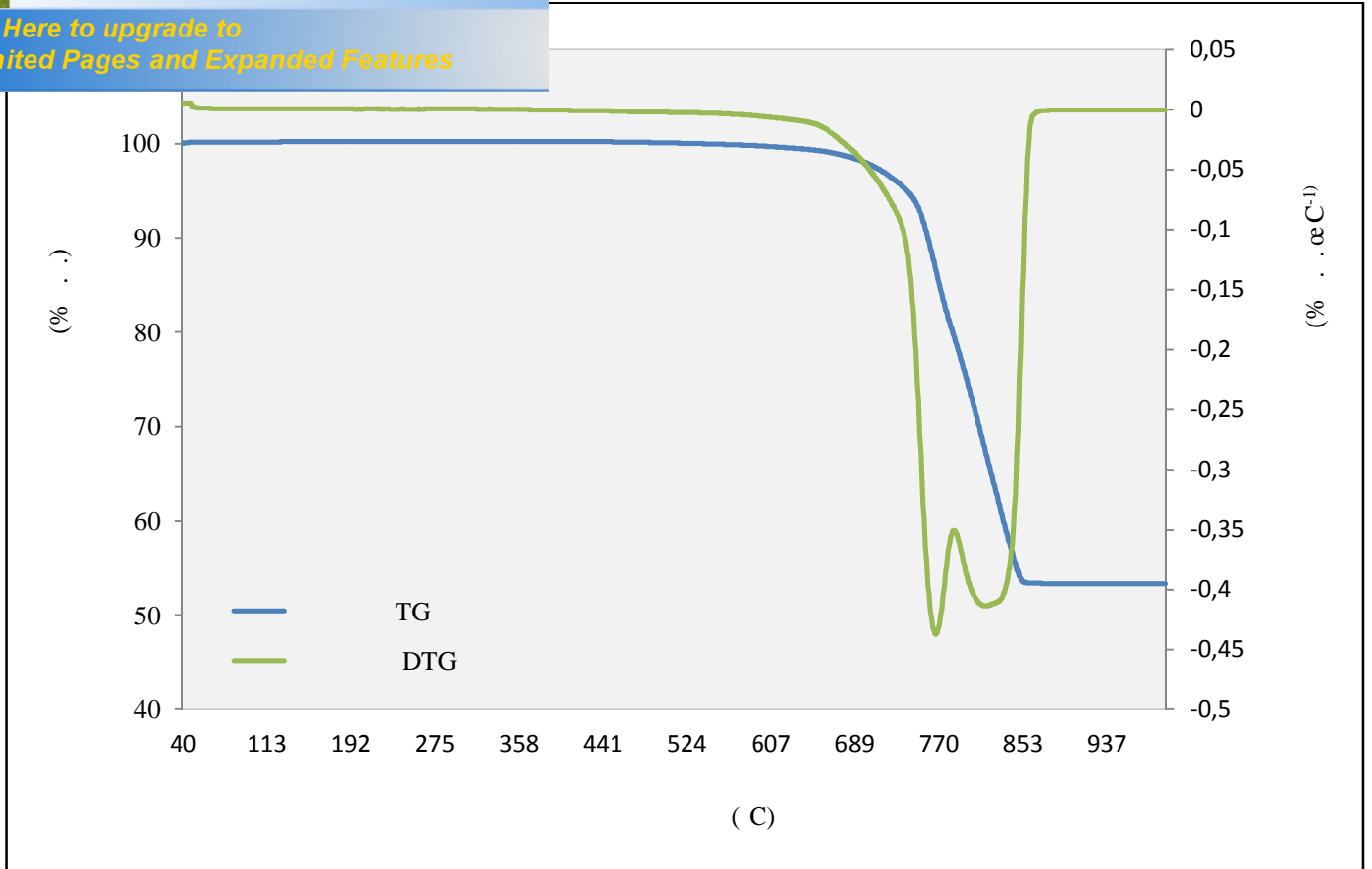
, MgO, CO<sub>2</sub>, CaCO<sub>3</sub>,

CaMg(CO<sub>3</sub>)<sub>2</sub> CaCO<sub>3</sub> + MgO + CO<sub>2</sub>

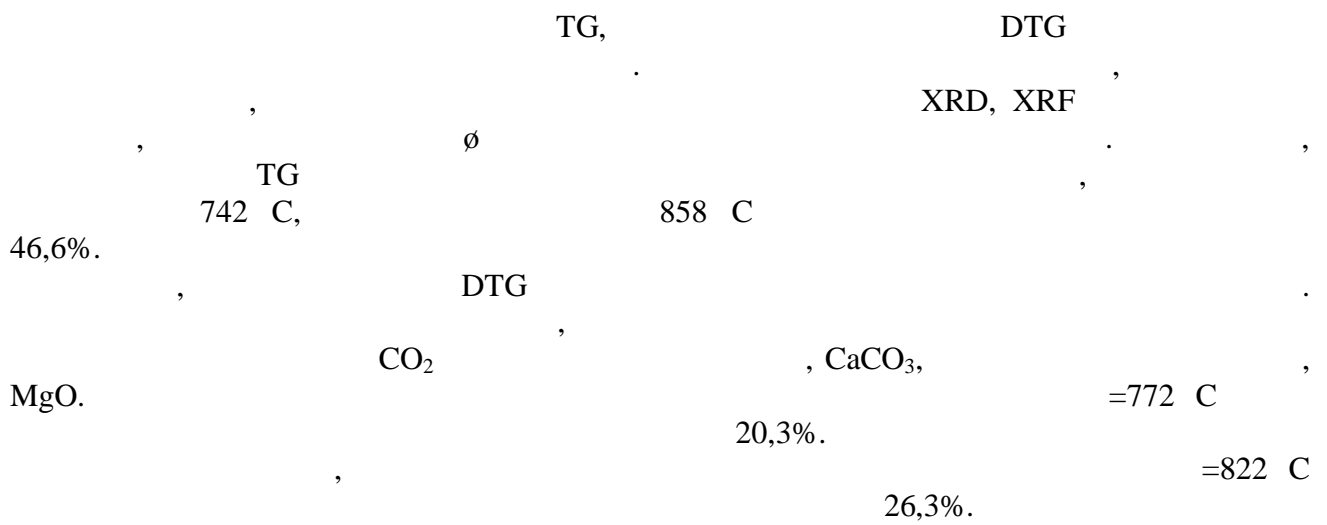
DTG : 26,1% =784 C.

CaCO<sub>3</sub> CaO + CO<sub>2</sub>

21,6%, =824 C.



3.14: TG DTG





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TG .

( C)	724	-	-	742	-	-
( C)	850	-	-	858	-	-
( C)	-	784	824	-	772	822
(%)	47,7	26,1	21,6	46,6	20,3	26,3

( Loss of gnition, LOI)

3.2.5.1

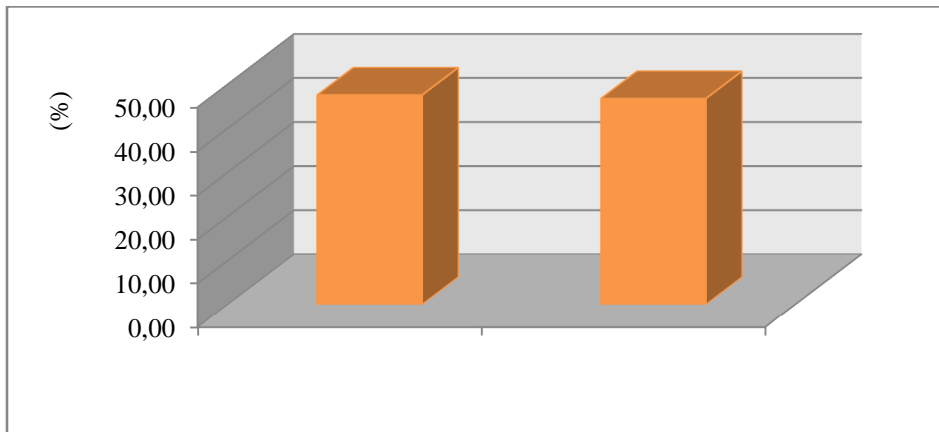
1 g , 1000 C , 1 h. ( )

3.2.5.2

$$LOI (\%) = \frac{\text{Mass of residue after LOI} - \text{Mass of residue before LOI}}{\text{Mass of sample}} \times 100$$

3.9:

	(%)
	46,70
	47,50



3.15: :

CO<sub>2</sub>.

**3.2.6.1**

[122].

z-

25 mV ( )

z-

( ... )

( )

[123].

( z- )

o,

/

z-

Laser Doppler Anemometer.

[124].



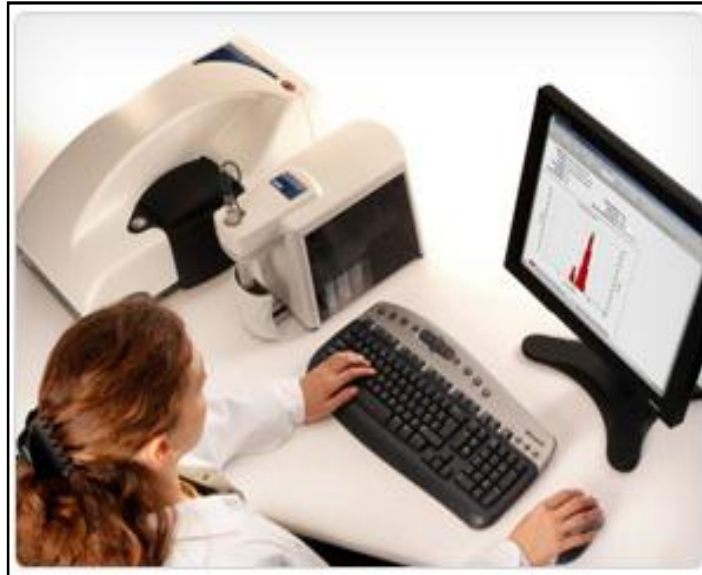
**3.16:**

[125].

< 10 m.

1 ml  
3.16 3.17

MALVERN Z-SIZER NANO-ZS.



3.17: [125].

3.2.6.3

3.10 3.11

3.10:

	(mV)
	-19,1
	-20,0





Fe, Al, Mn, P, Sr . . . ,

SiO<sub>2</sub>,

, CaCO<sub>3</sub>,

, MgO,

CO<sub>2</sub>,  
 CaO

CO<sub>2</sub>.

47%

CO<sub>2</sub>.

-20 mV,  
 0,01-0,03 mS/cm.

[13,31] .

[31].

[5].

3.3.1

3.3.1.1

H

5 g, 30 ml, 5, (20-25 °C), 1h, pH 400 rpm, 1 h

pH

1 h

100 °C



3.20:

3.3.1.2

Pb, Cd, Zn Cu

:

%

$$= \frac{[?] [?] [?] [?]}{[?]} \times 100$$

C<sub>0</sub>:

mg/L

C :

mg/L.

$$q = \frac{C_o - C}{M} \alpha V$$

$C_o$ : mg/L  
 $C$ :  
 mg/L  
 $M$ : g  
 $V$ : o (L).

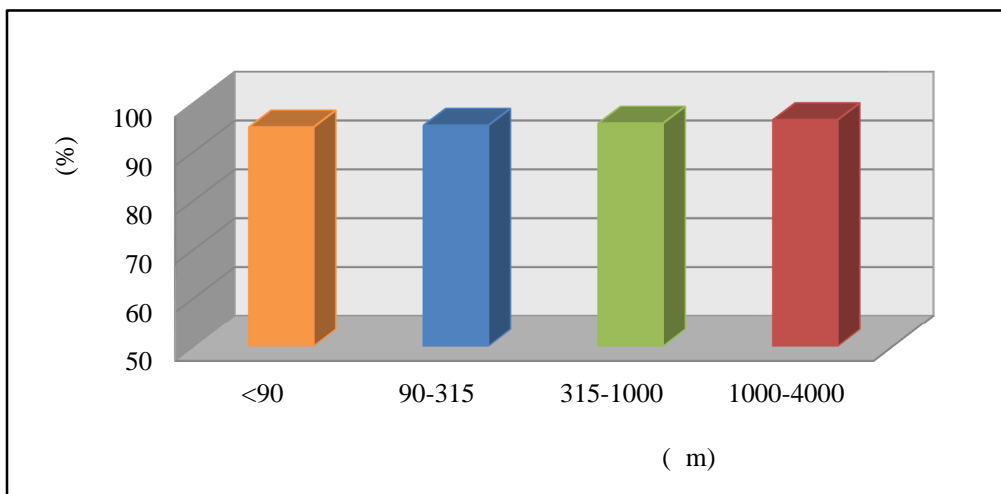
Pb

3.12:

Pb

Pb

( m )	pH	( g Pb/g )	(%)
<90	8,87	28,49	94,98
90-315	9,44	28,60	95,34
315-1000	9,37	28,73	95,76
1000-4000	8,66	28,94	96,48



3.21:

Pb

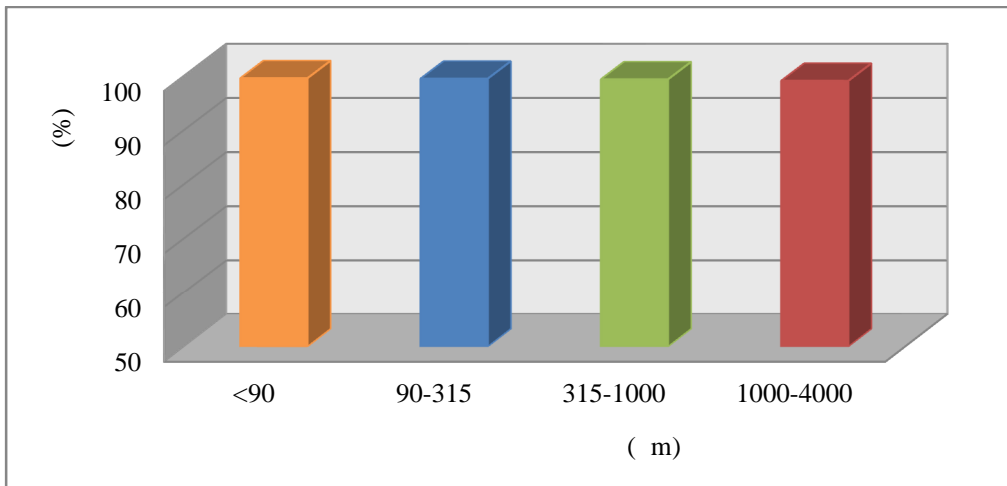
Pb

, pH Ca<sup>2+</sup> Mg<sup>2+</sup> 5 95-96% 9.

\_\_\_\_\_ Cd

3.13: Cd Cd

( m )	pH	( g Cd/g )	(%)
<90	8,73	29,87	99,58
90-315	9,32	29,85	99,50
315-1000	9,16	29,81	99,38
1000-4000	8,64	29,73	99,10



3.22: Cd Cd

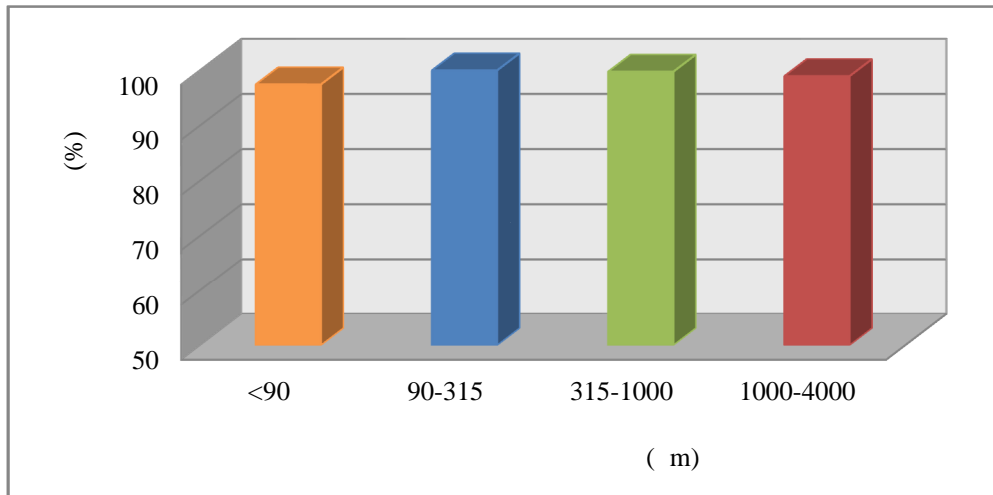
99% pH  
 pH=9,

3.14:

Zn

Zn

( m )	pH	( g Zn/g )	(%)
<90	7,15	584,40	97,40
90-315	8,41	599,70	99,95
315-1000	8,25	598,29	99,72
1000-4000	7,91	593,45	98,91



3.23:

Zn

Zn

3.14

pH

Ca Mg

7-8

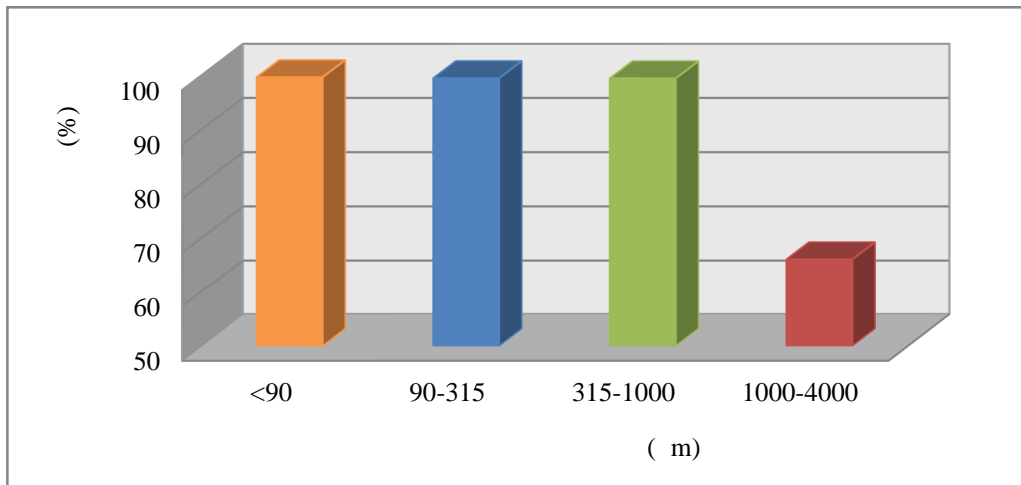
98-99%.

3.15:

Cu

Cu

( m )	pH	( g Zn/g )	(%)
<90	7,47	597,85	99,64
90-315	7,57	596,79	99,47
315-1000	6,12	596,67	99,45
1000-4000	6,97	396,30	66,05



3.24:

Cu

Cu

<90 m, 90-315 m, 315-1000 m, 1000-4000 m, 99%,  
 66%. pH  
 6-7 pH  
 Ca Mg  
 3.12 3.15

3.16:

Pb, Cd, Zn Cu

( m)	pH	( g /g )	(%)
Pb			
<90	6,77	28,12	93,74
90-315	6,64	27,81	92,70
315-1000	6,68	28,18	93,92
1000-4000	6,43	26,33	87,78
Cd			
<90	6,77	22,31	74,36
90-315	6,64	11,16	37,20
315-1000	6,68	14,52	48,40
1000-4000	6,43	3,90	13,00
Zn			
<90	6,77	352,8	58,80
90-315	6,64	275,4	45,90
315-1000	6,68	328,2	54,70
1000-4000	6,43	246,6	41,10
Cu			
<90	6,77	587,35	97,89
90-315	6,64	562,50	93,75
315-1000	6,68	580,50	96,75
1000-4000	6,43	447,48	74,58

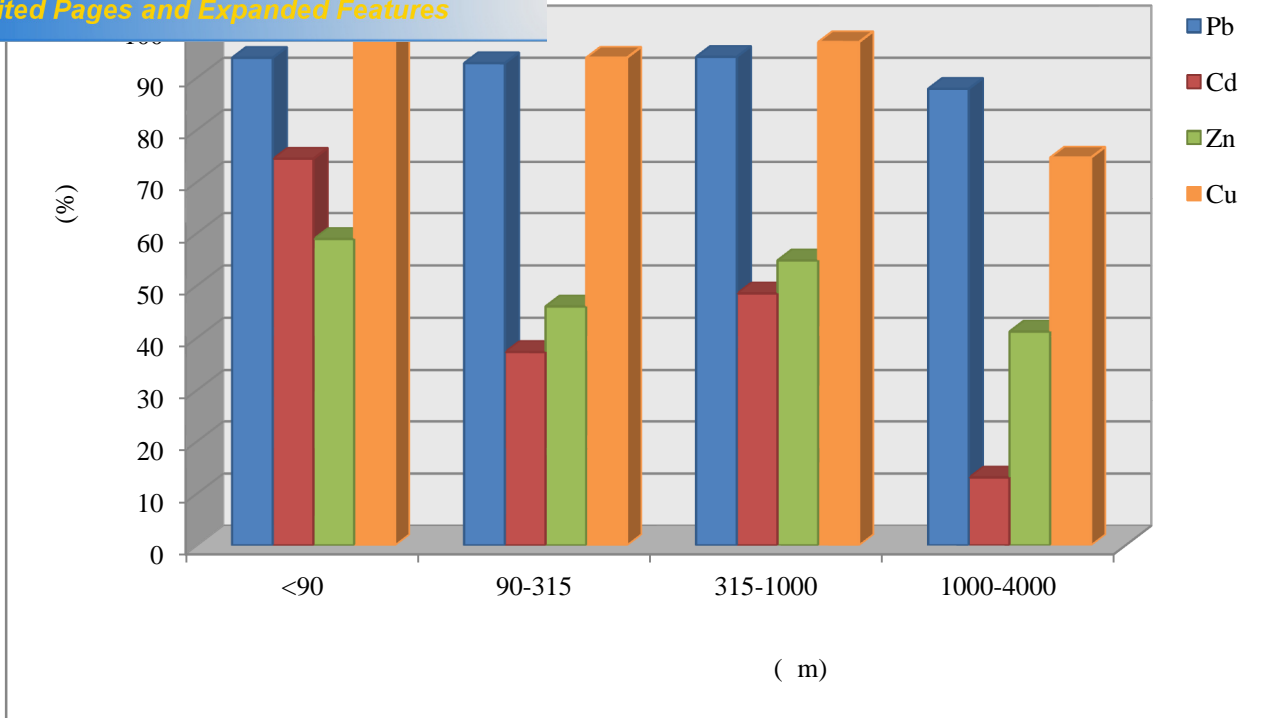
m

pH

5

1000-4000

6-7,



3.25:

Pb, Cd, Zn Cu

88% 94-98% 93-94%  
75%  
<90 m, 90-315 m 315-1000 m 74%, 37% 48%,  
1000-4000 m, 13%  
59%, 46%, 55% 41% <90 m, 90-315  
m, 315-1000 m 1000-4000 m,

Cu>Pb>Zn>Cd.

4000 m,

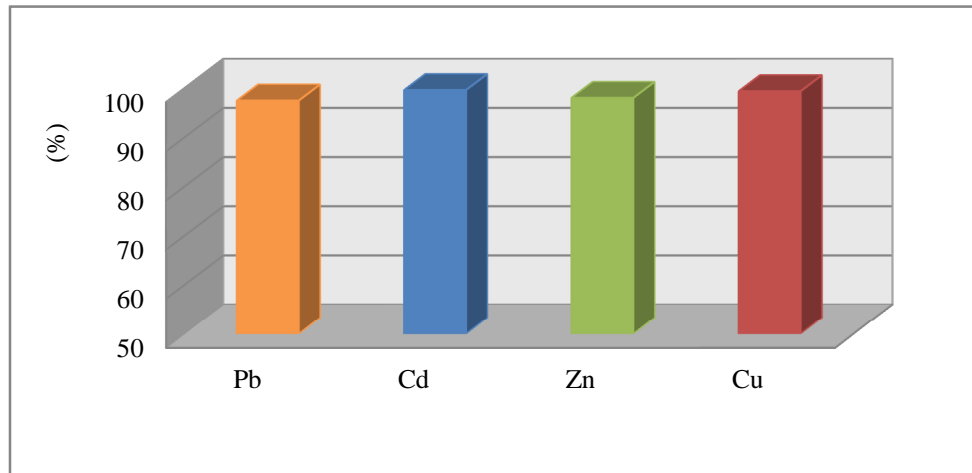
1000-



Pb, Cd, Zn Cu

**3.17:** Pb, Cd, Zn Cu, ,

	( m )	pH	( g /g )	(%)
Pb	315-1000	9,90	29,25	97,50
Cd	315-1000	9,74	29,89	99,64
Zn	315-1000	8,91	588,30	98,05
Cu	315-1000	8,48	596,44	99,41

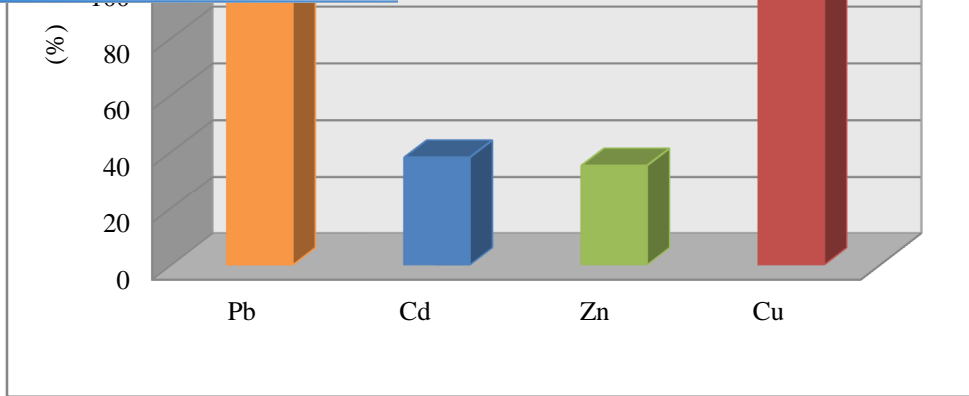


**3.26:** Pb, Cd, Zn Cu, ,

Pb, Cd, Zn Cu

**3.18:** Pb, Cd, Zn Cu

	( m )	pH	( g /g )	(%)
Pb	315-1000	7,73	29,15	97,16
Cd	315-1000	7,73	11,40	38,00
Zn	315-1000	7,73	210,60	35,10
Cu	315-1000	7,73	570,82	95,14



3.27 :

Pb, Cd, Zn Cu

97-99%

Cd>Cu>Zn>Pb.

99%,

98%,

97%.

Pb, Cd, Zn

Cu

97%

95%,

38%

35%,

8-10

pH

7-8

Ca

Mg.

### 3.3.1.3

pH=5.



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n Cu  
95%,  
315-1000 m

Cu>Cu>Zn>Pb,

Cd>Cu>Zn>Pb.

1000-4000 m

Cu

Cu>Pb>Zn>Cd

Pb>Cu>Cd>Zn.

1000-4000 m,

- I.
- II.

[15].

[15,20].

Ca Mg

[19,98].

[15].

3.3.2.1

3.19:

	( m)	( m)
Zn	< 90	315-1000
Cu	1000-4000	
Pb	< 90	
Cd	1000-4000	
Zn, Cu, Pb,Cd	315-1000	

400 rpm (20-25 °C). (5 g), 30 ml, 1 h  
100 °C, 1 h

3.3.2.2

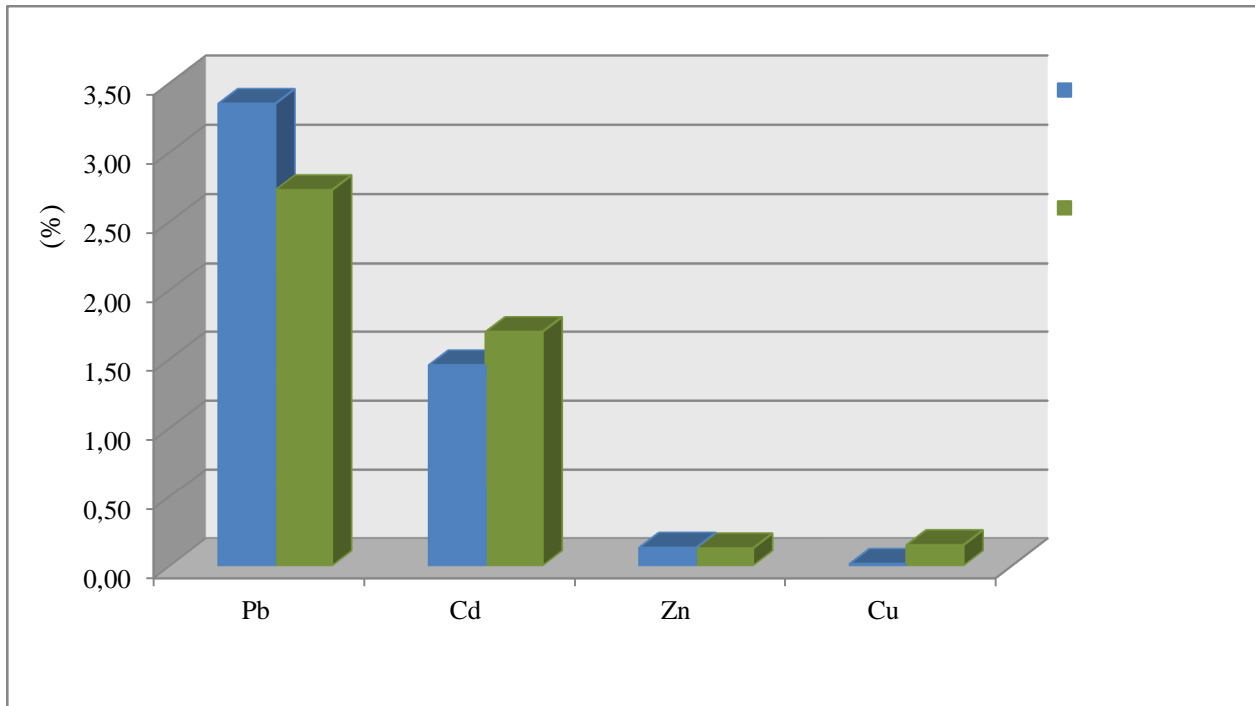
:

$$\% = \frac{m}{m} \times 100$$

m : (5 g), mg  
m: mg.

3.20: Pb, Cd, Zn Cu

	(% )	
Pb	3,35	2,73
Cd	1,45	1,69
Zn	0,13	0,12
Cu	0,01	0,15



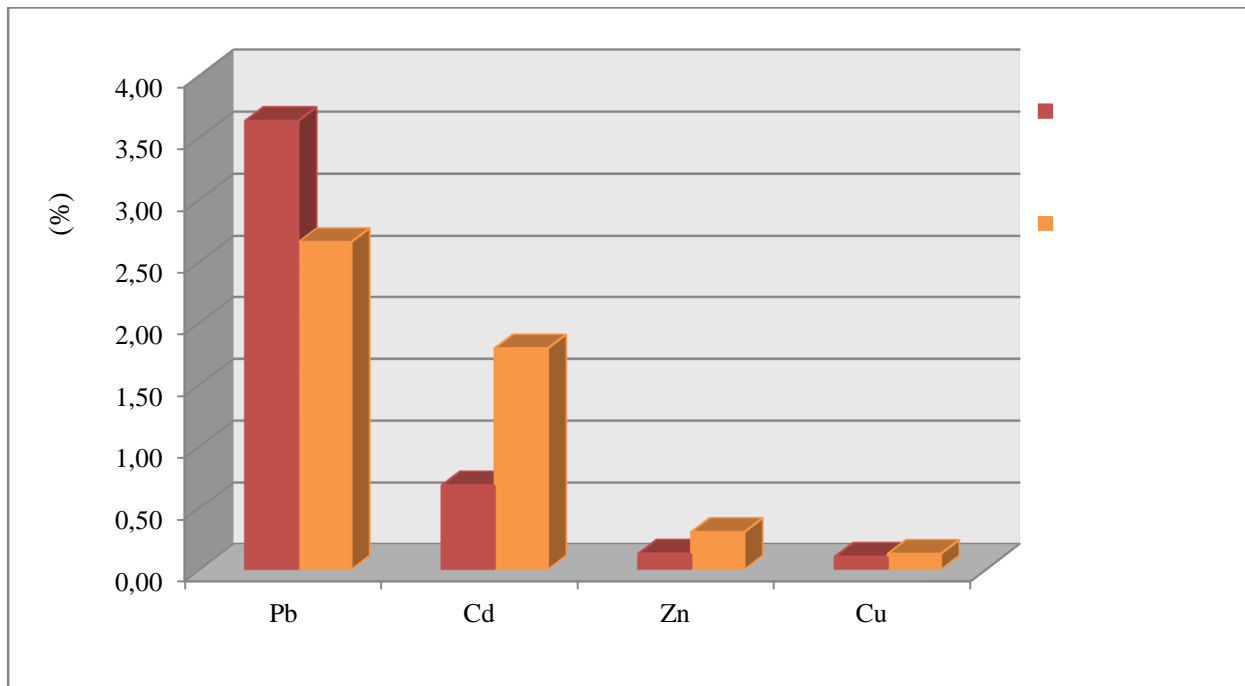
3.28: Pb, Cd, Zn Cu

2%,

Ca

**3.21:** Pb, Cd, Zn Cu

	( % )	
Pb	3,63	2,66
Cd	0,68	1,79
Zn	0,13	0,30
Cu	0,10	0,13



**3.29:** Pb, Cd, Zn Cu

Ca Mg,

### 3.3.2.3

Pb, Cd, Zn Cu

Zn Cu. Pb Cd  
4%,  
: Pb>Cd>Zn=Cu.  
Zn Cu  
Ca Mg

3.4.1

3.4.1.1

Pb, Cd, Zn Cu

30 cm

5 cm.

<1mm.

1,25 ml/min  
100 ml,

pH

60 C

24 h



3.30:





3.31:

:  
:  
✓  
✓

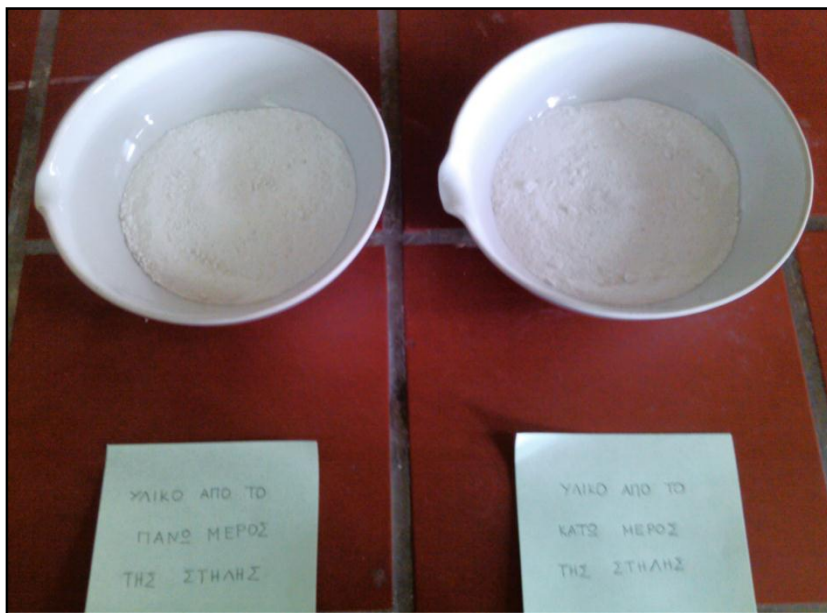
( 3.32).

( 3.33 3.34).

( 3.35).



3.32:



3.33:

( ) ( )



3.34:

( ) ( ) .



3.35:

3.4.1.2

Pb, Cd, Zn Cu

:

$$\% = \frac{2222222222}{222} \times 100$$

C<sub>0</sub>:

mg/L

C:

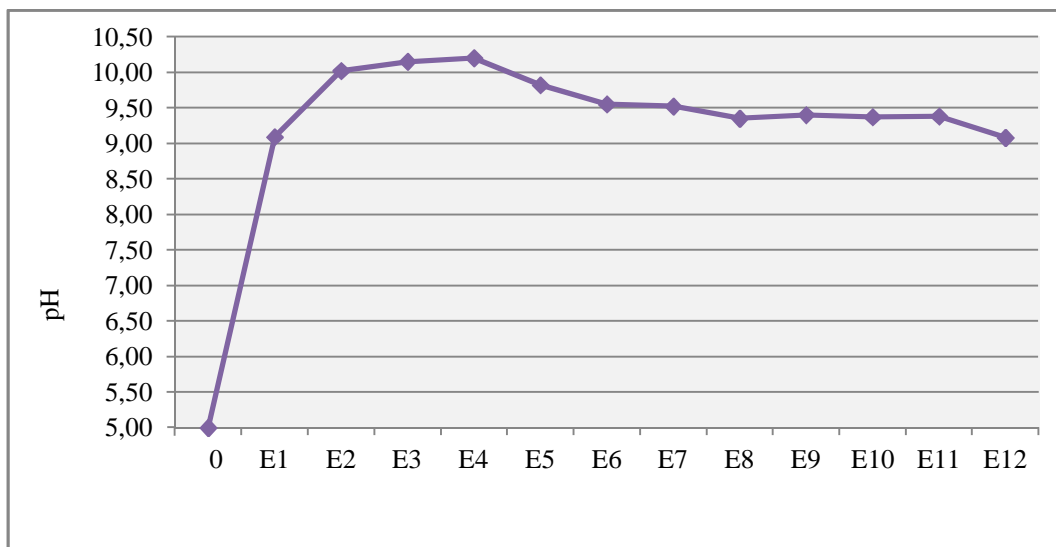
mg/L.

3.22:

Pb, Cd, Zn Cu

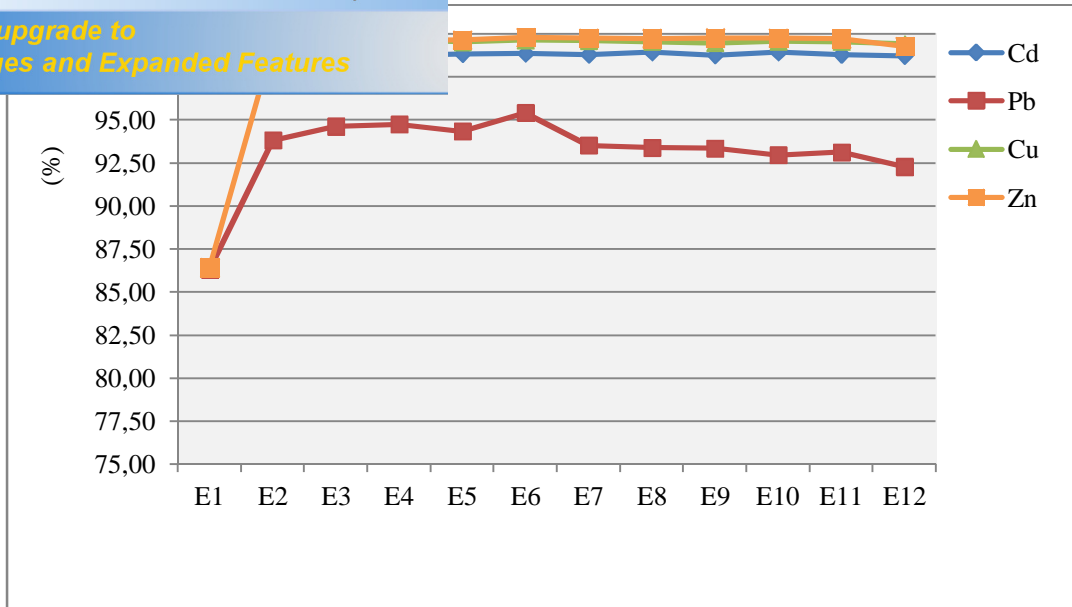
	pH	%			
		Pb	Cd	Zn	Cu
1	9,09	86,30	98,64	86,43	97,81
2	10,02	93,82	99,06	99,33	99,52
3	10,15	94,62	98,82	99,44	99,63
4	10,20	94,74	98,74	99,64	99,69
5	9,82	94,34	98,84	99,65	99,51
6	9,55	95,42	98,86	99,80	99,63
7	9,52	93,52	98,80	99,74	99,60
8	9,35	93,38	98,92	99,73	99,55
9	9,40	93,35	98,76	99,76	99,47
10	9,37	92,96	98,92	99,74	99,55
11	9,38	93,12	98,80	99,70	99,54
12	9,08	92,28	98,72	99,27	99,41

pH 9-10. 5  
Mg<sup>2+</sup> 2 4 10  
pH 9,5. 5  
Ca<sup>2+</sup>



3.36: pH

Pb, Cd, Zn Cu



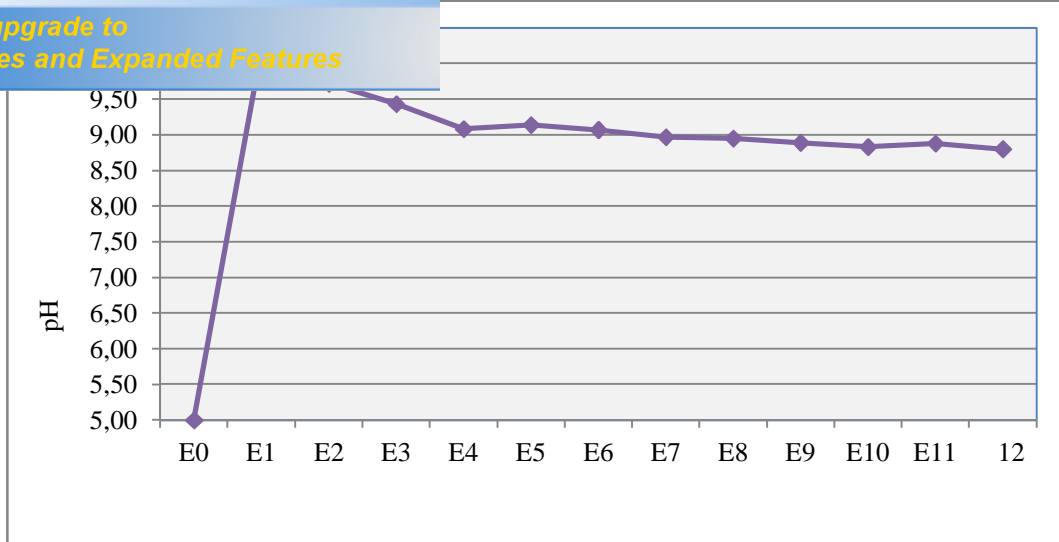
3.37: (%) Pb, Cd, Zn Cu

Cd, Zn Cu 3.22, Pb, 98-99% 93-95%, 99-100% 3.37, Cu 1,25 ml/min. To

3.23 : Pb, Cd, Zn Cu

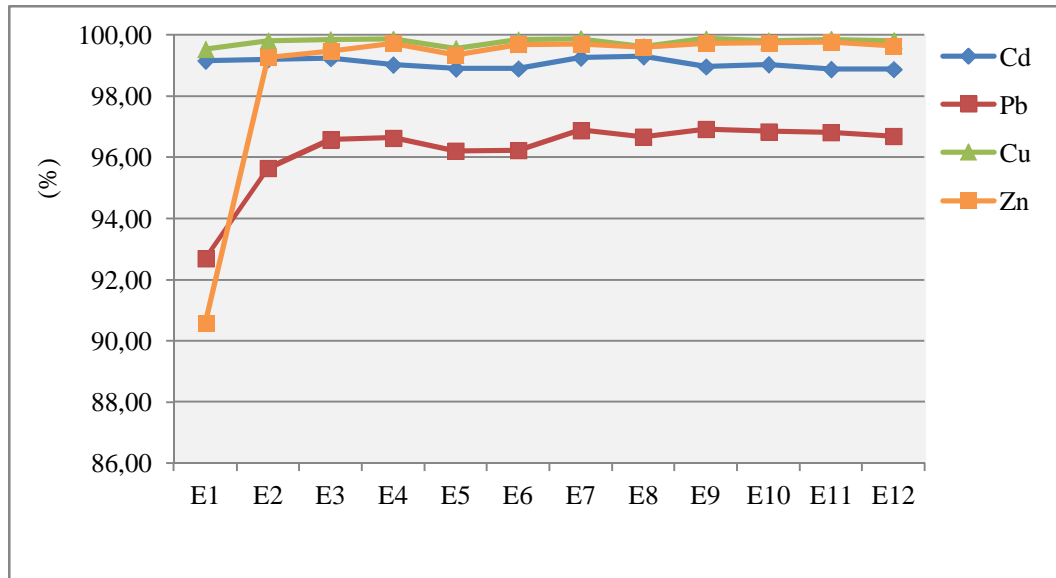
	pH	(%)			
		Pb	Cd	Zn	Cu
1	10,21	92,69	99,16	90,58	99,53
2	9,72	95,64	99,21	99,27	99,81
3	9,43	96,58	99,24	99,48	99,85
4	9,08	96,64	99,03	99,73	99,86
5	9,14	96,21	98,90	99,34	99,56
6	9,07	96,23	98,90	99,69	99,84
7	8,97	96,89	99,26	99,71	99,88
8	8,95	96,67	99,30	99,60	99,62
9	8,89	96,92	98,97	99,73	99,88
10	8,83	96,85	99,04	99,74	99,81
11	8,88	96,82	98,88	99,76	99,86
12	8,80	96,70	98,88	99,65	99,82

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3.38: pH Pb, Cd, Zn Cu

9-10. pH  
 10, 2 pH  
 9. , pH Ca Mg



3.39 : (%) Pb, Cd, Zn Cu

3.23

99-100%,

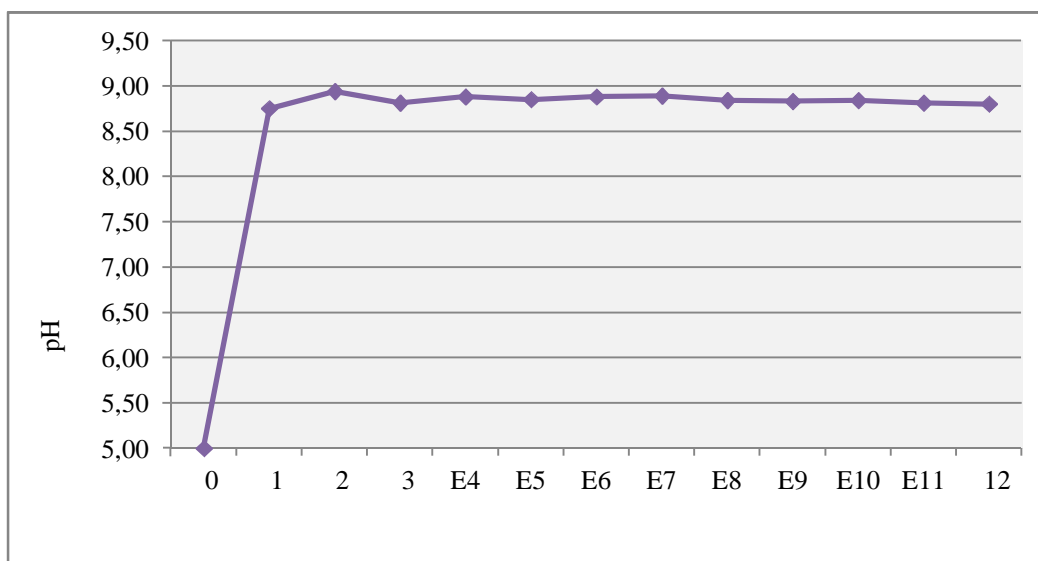
98-99%,

96-97%.

3.24:

Pb, Cd, Zn Cu

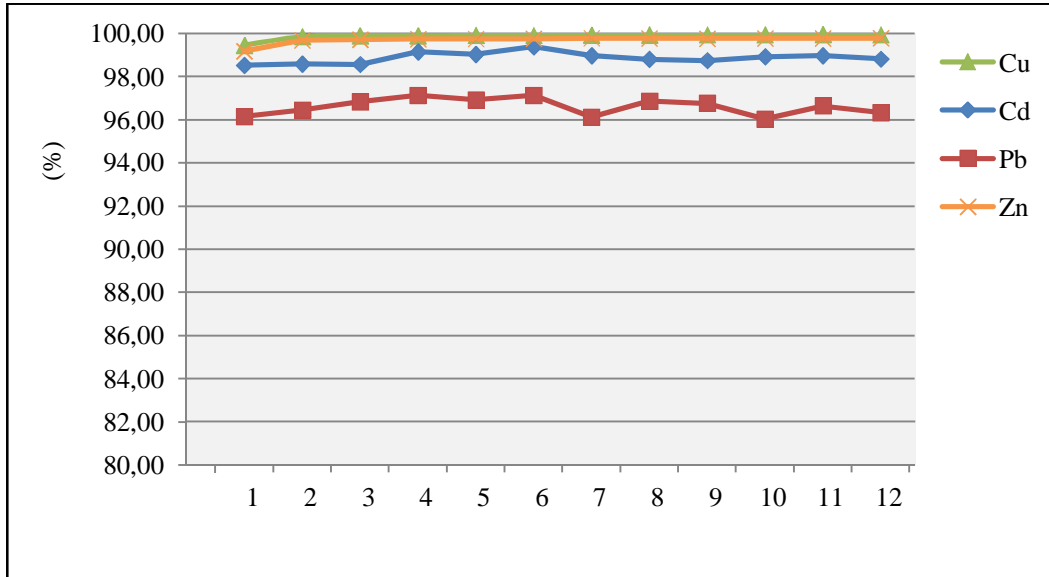
	pH	(%)			
		Pb	Cd	Zn	Cu
1	8,75	96,16	98,52	99,17	99,45
2	8,94	96,45	98,58	99,68	99,84
3	8,81	96,83	98,56	99,71	99,89
4	8,88	97,14	99,14	99,74	99,89
5	8,85	96,91	99,02	99,75	99,91
6	8,88	97,13	99,37	99,76	99,91
7	8,89	96,11	98,96	99,78	99,93
8	8,84	96,86	98,80	99,77	99,93
9	8,83	96,76	98,74	99,76	99,94
10	8,84	96,02	98,92	99,77	99,93
11	8,81	96,64	98,96	99,77	99,95
12	8,80	96,33	98,81	99,77	99,94



3.40: pH

Pb, Cd, Zn Cu

3.40 pH  
Ca



3.41: (%) Pb, Cd, Zn Cu

Pb, Cd, Zn Cu

100%, , 98-99% 96-97%, « »

3.4.1.3

pH=5





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Cd,

,  
.

93-97%,

98-99%

99-100%.

Zn=Cu>Cd>Pb.

, 93-95%

96-97%,

( )

5

pH  
8-10.

pH

9-10,

8-9.

Ca,

3.4.2.1

5 g , 50 ml , 1 h , 400 rpm , (20-25 °C). 1 h , 100 °C,

3.4.2.2

:

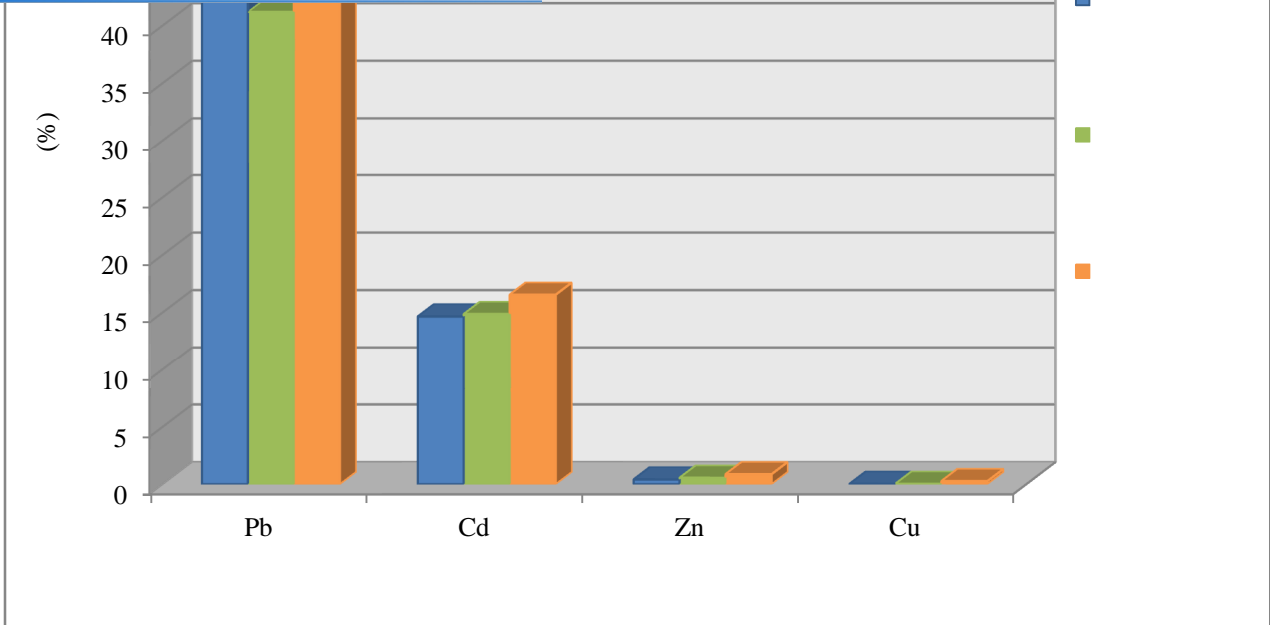
$$\% = \frac{\square}{\square} \times 100$$

m : (5 g), mg  
m: mg.

3.25: (%) Pb, Cd, Zn Cu ,

	(%)		
Pb	42,15	41,15	43,92
Cd	14,62	14,83	16,51
Zn	0,41	0,55	0,86
Cu	0,02	0,03	0,26

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3.42: (%) Pb, Cd, Zn Cu ,

3.25, Pb, Cd, Zn Cu, 40%, 15%, <1%.

Ca Mg



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∅

, Pb>Cd>Zn=Cu.

40%,  
15%.

1%.

Zn

Cu

Pb

Cd

3.4.3.1

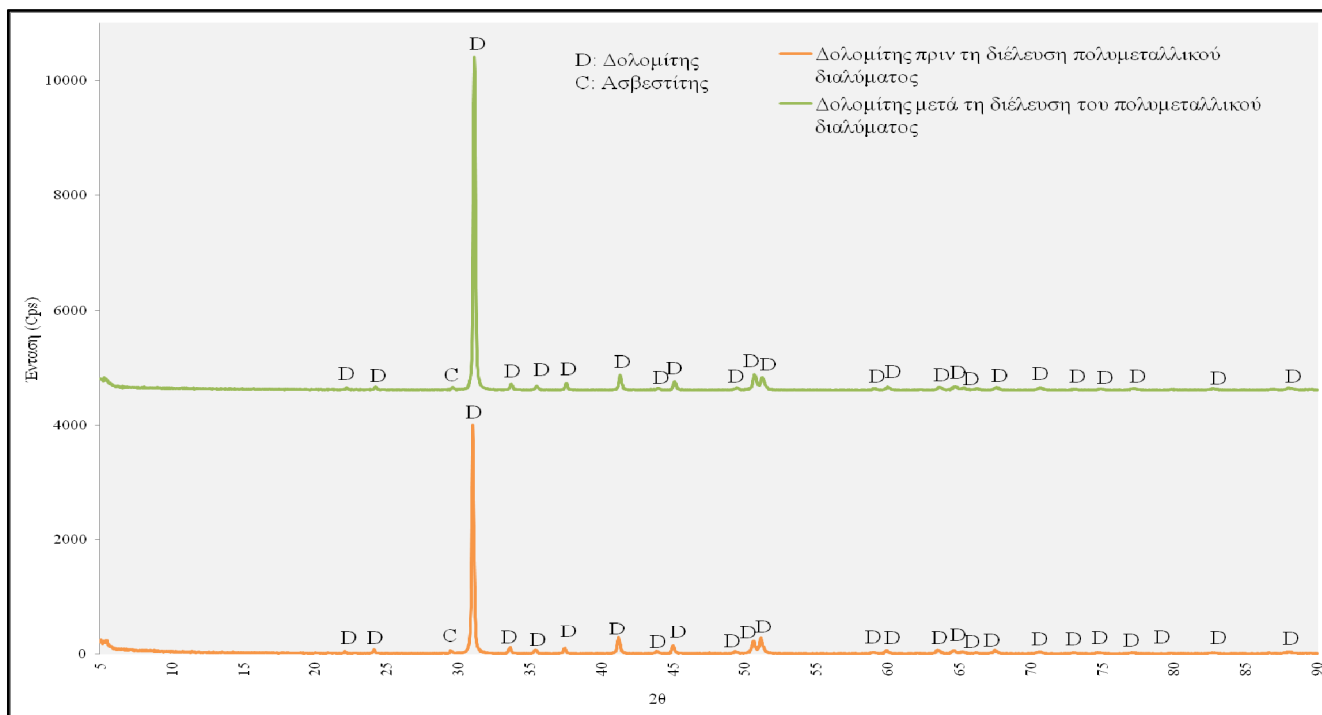
(X-Ray Diffraction, XRD)

3.4.3.1.1

3.2.2.2.1

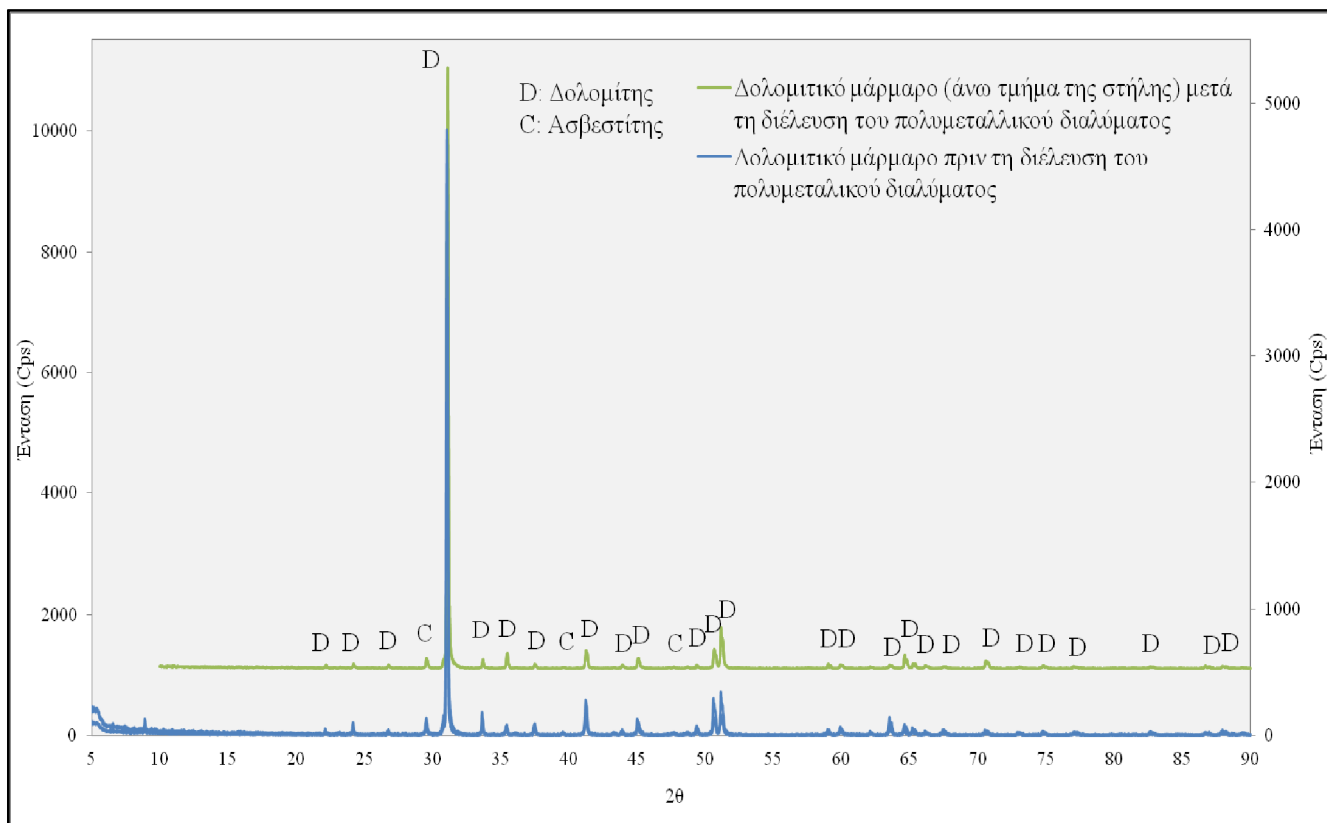
- i. 2 : 10-90
- ii. Step: 0.050°
- iii. Step time: 2 s
- iv. Temp: 25°C

3.4.3.1.2



3.43:

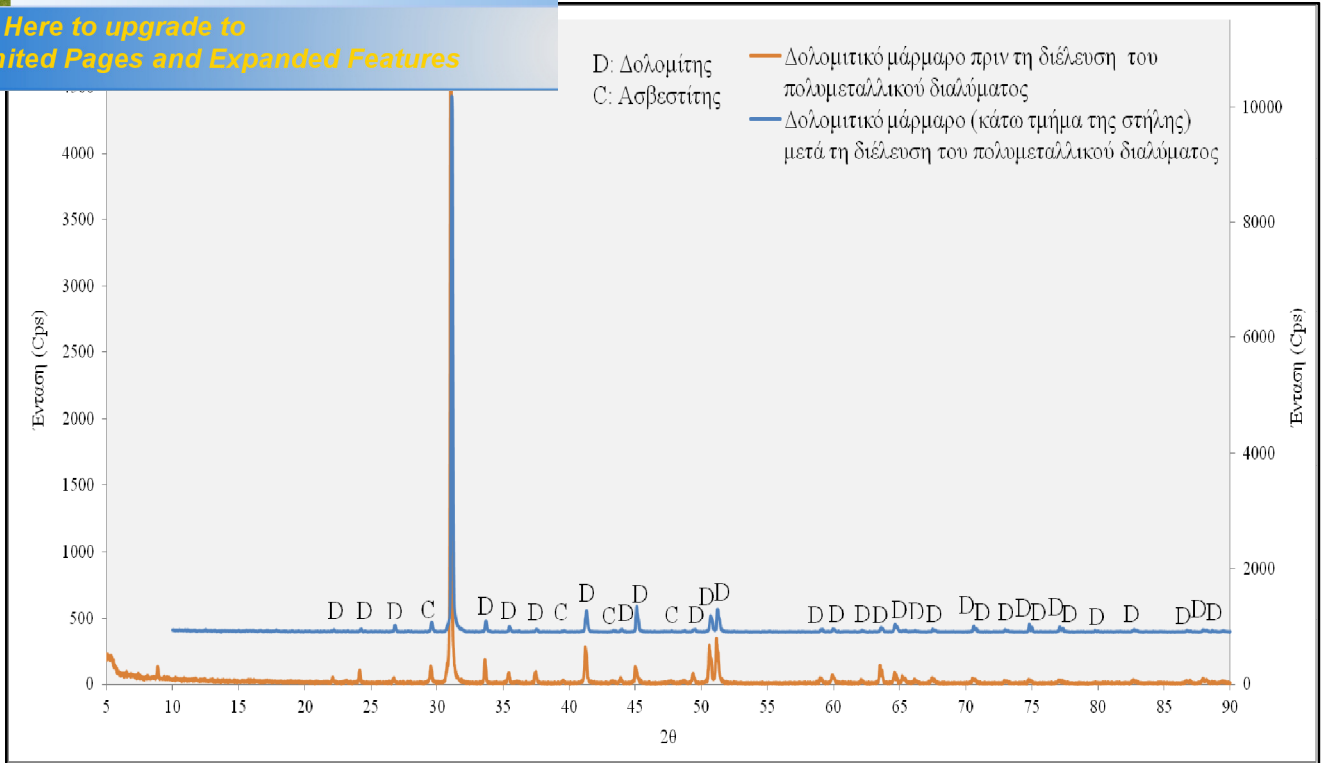
3.44 3.46



3.44:

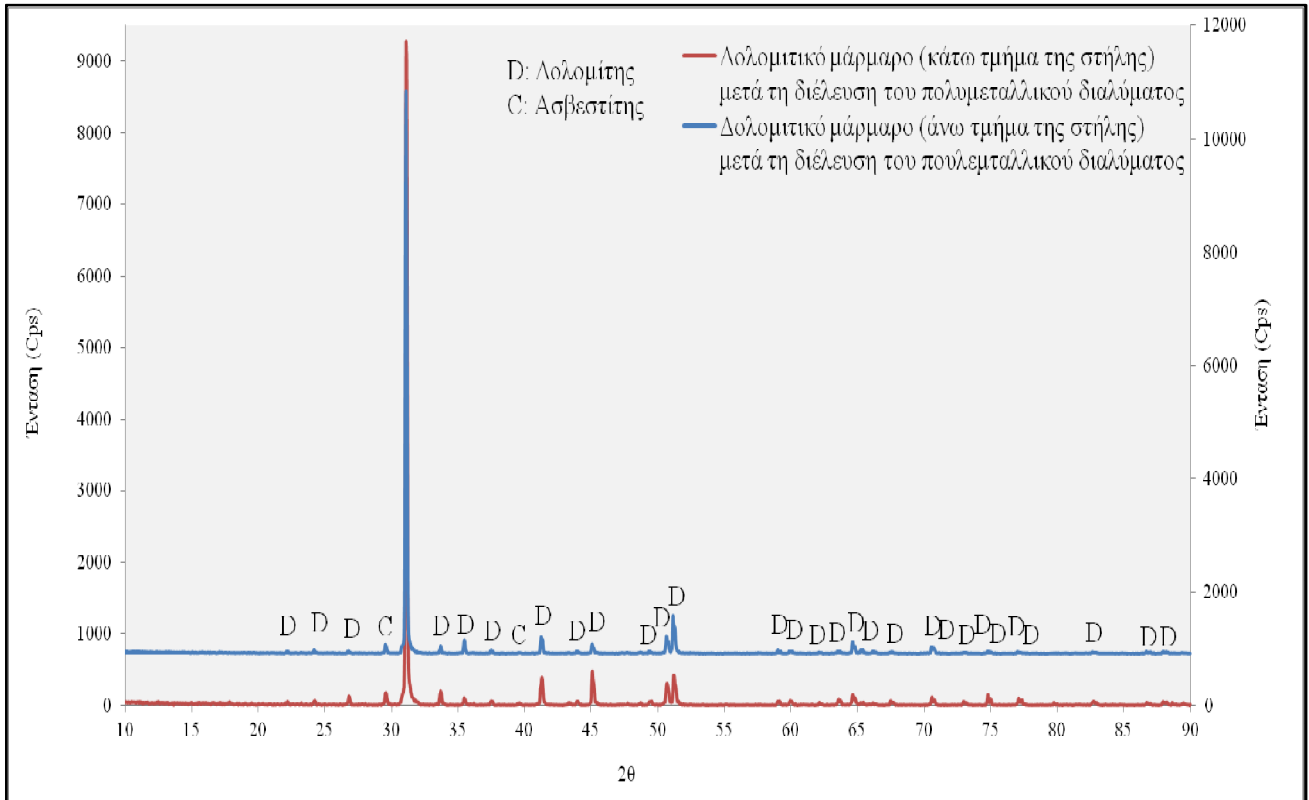
( )

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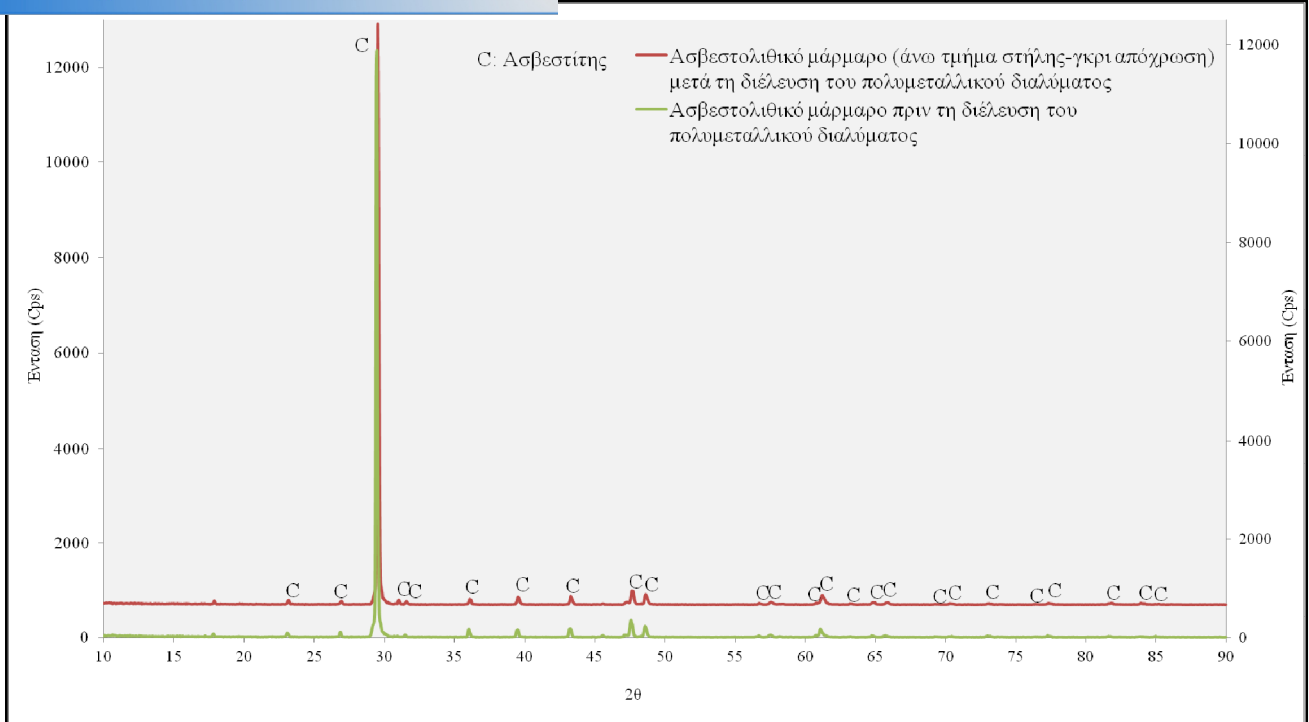


3.45:

( )



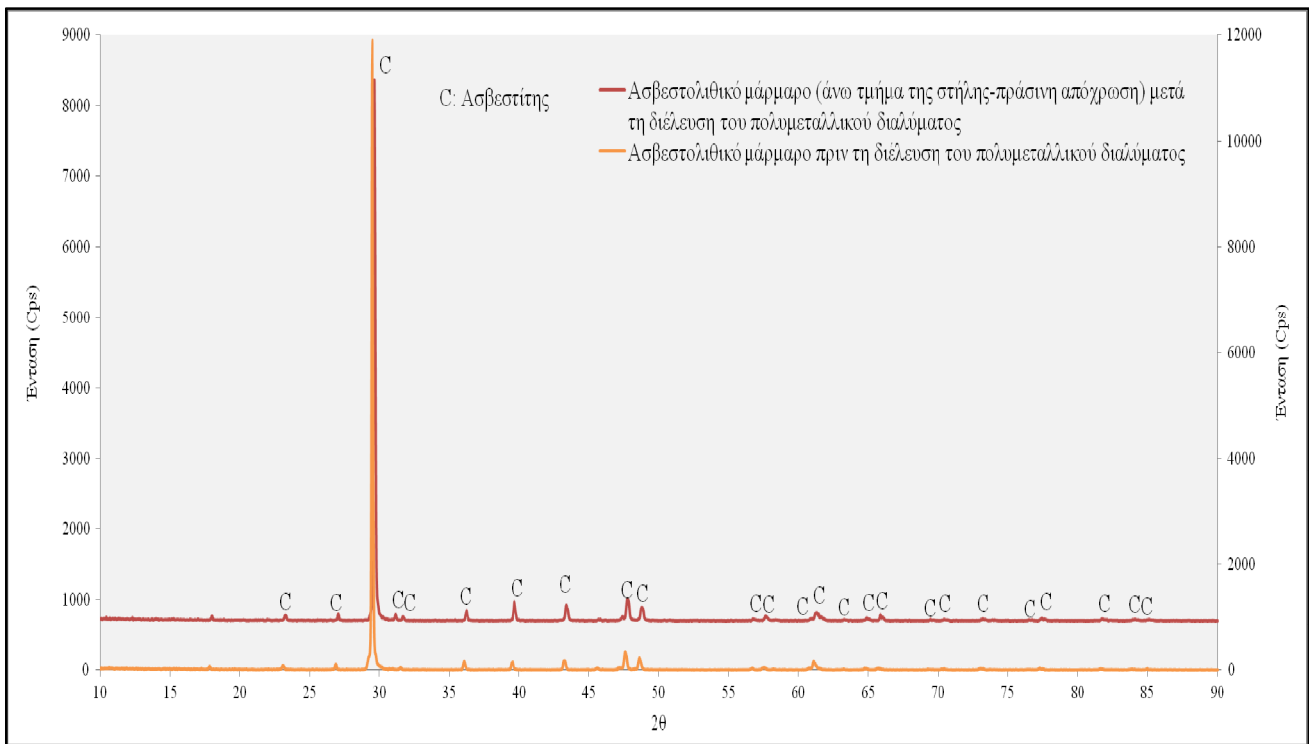
3.46:



3.47:

)

(



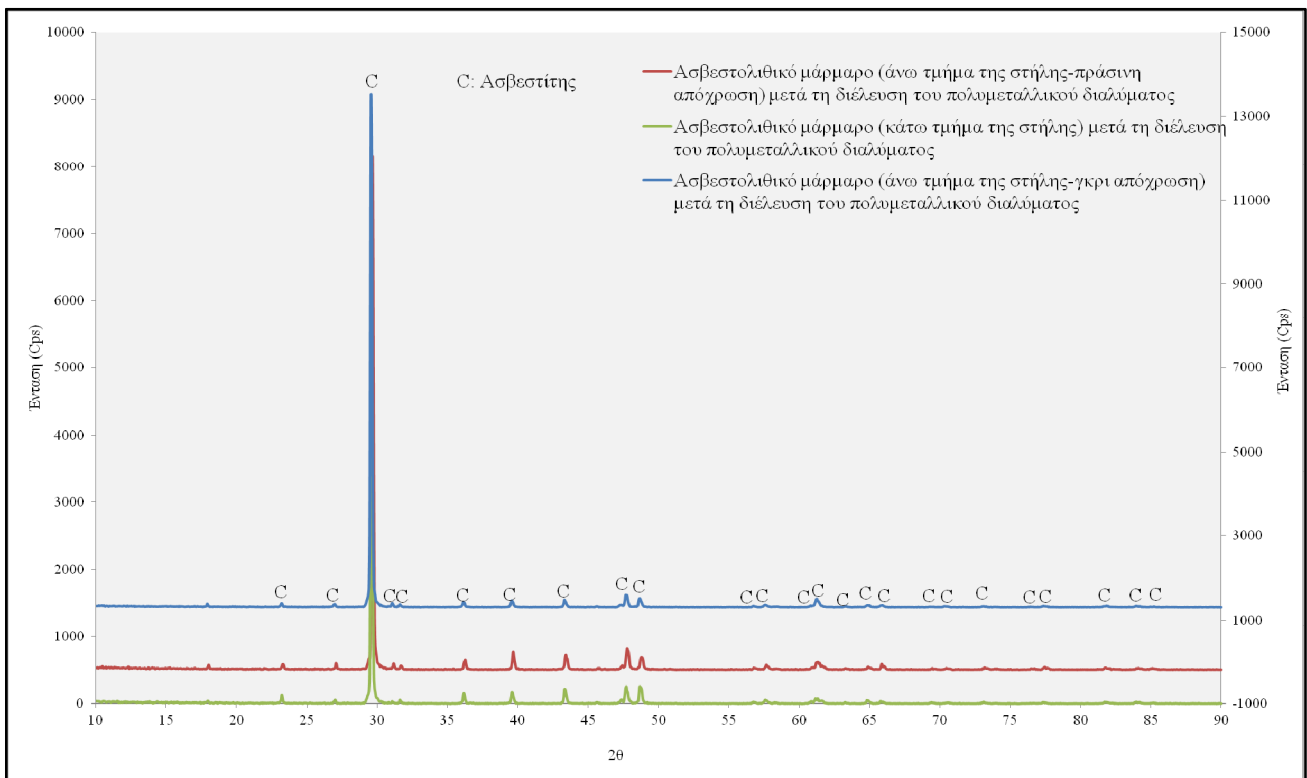
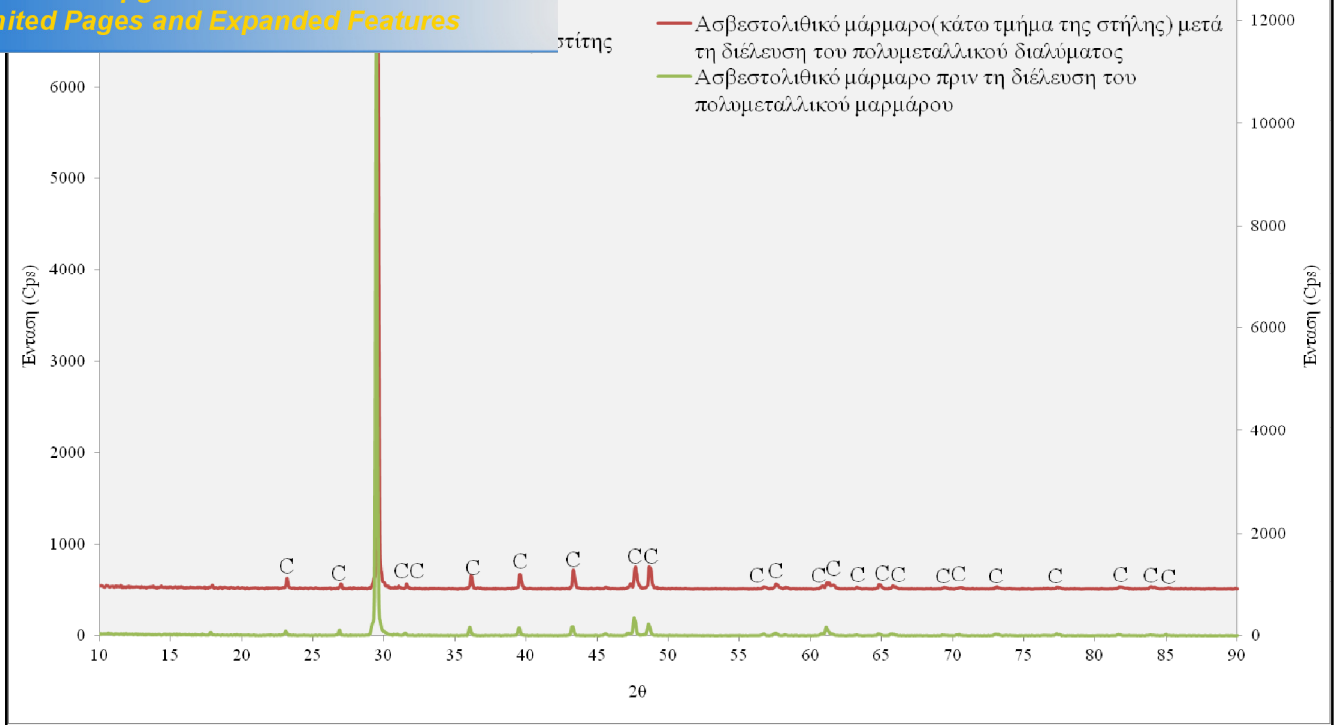
3.48:

)

(



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3.50,

3.4.3.2.1

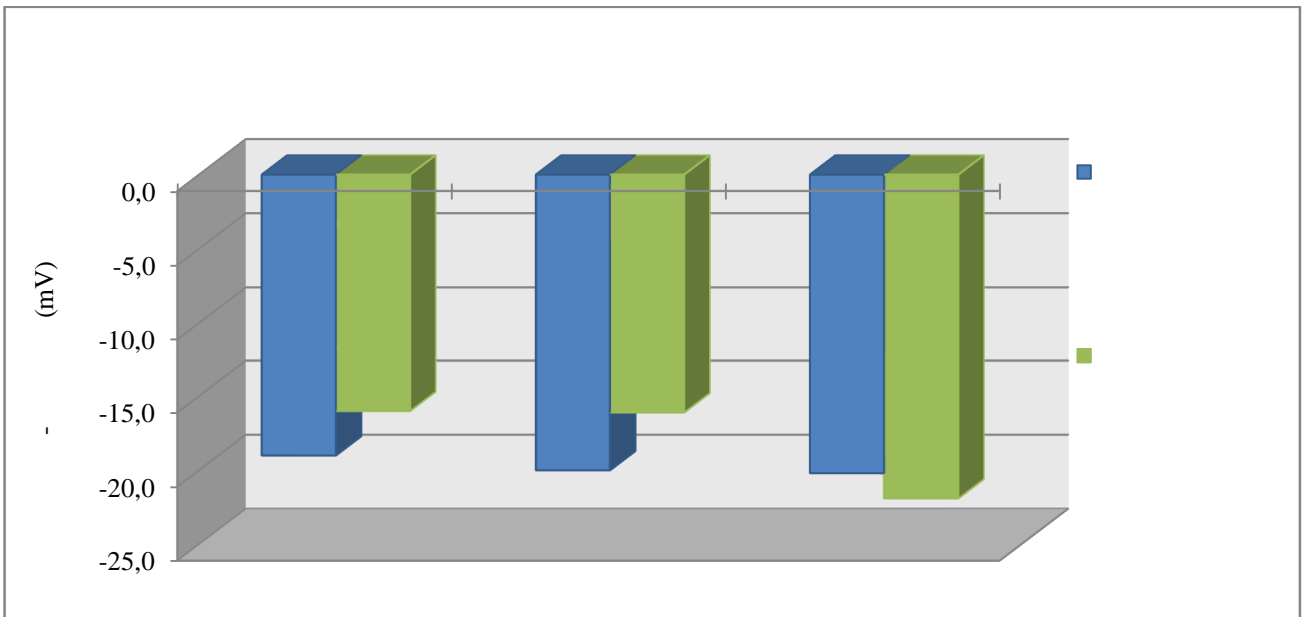
3.2.6.2.

3.4.3.2.2

3.26 3.27

3.26:

	(mV)	
	-19,0	-16
	-20,0	-16,1
	-20,2	-21,9



3.51:

-19,0 - -20,0 mV

-16,0 mV.

-21,9 mV.

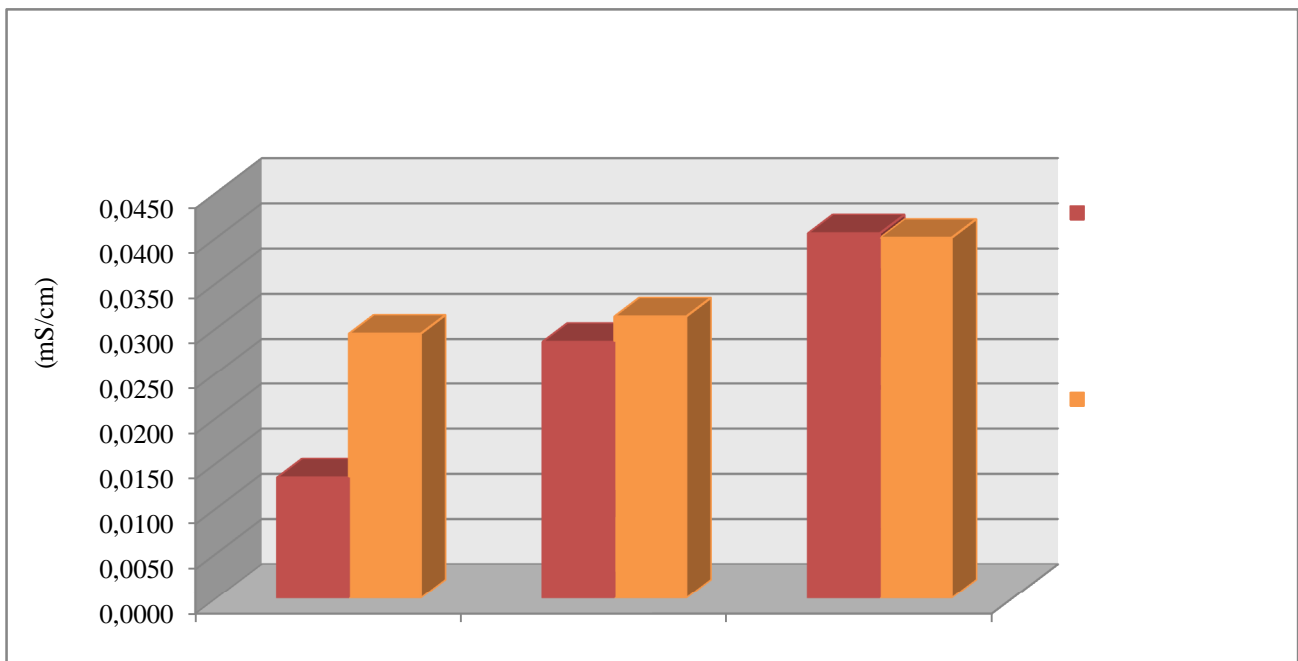
Mg

-16 mV

Ca,

3.27:

	(mS/cm)	
	0,0134	0,0294
	0,0285	0,0313
	0,0405	0,0400



3.52:



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### 3.4.3.3

Mg



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✓

, SiO<sub>2</sub> Fe, Al, Mn, P, Sr . . .

✓

95%,  
Cd>Cu>Zn>Pb.  
1000-4000 m

pH=5

pH=5

: Cu>Pb>Zn>Cd.

1000-4000 m

pH=5,  
Cd>Cu>Zn>Pb.  
Pb Cu  
Pb>Cu>Zn>Cd.

97%  
Cd Zn

pH



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99%,

99-100%,  
93-97%.

98-

96-97%  
, 93-95%,  
Zn=Cu>Cd>Pb.

«

»

pH

Ca Mg

✓

: (i)

(ii)

Ca Mg

4%.

Pb>Cd>Zn=Cu.





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Pb>Cd>Zn=Cu.

Pb      Zn      Cu  
          Cd

✓

Mg



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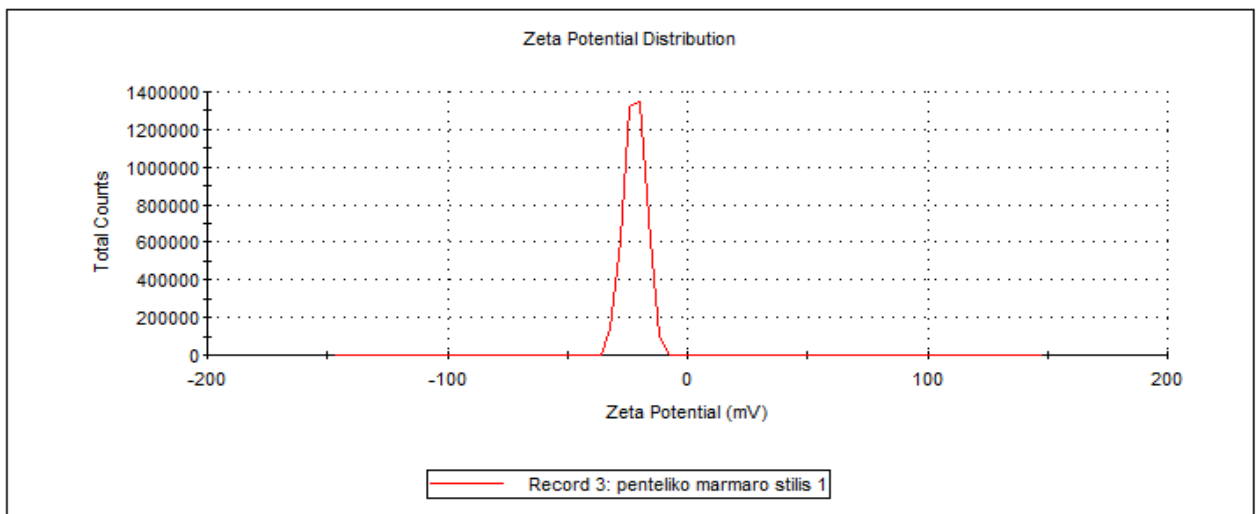
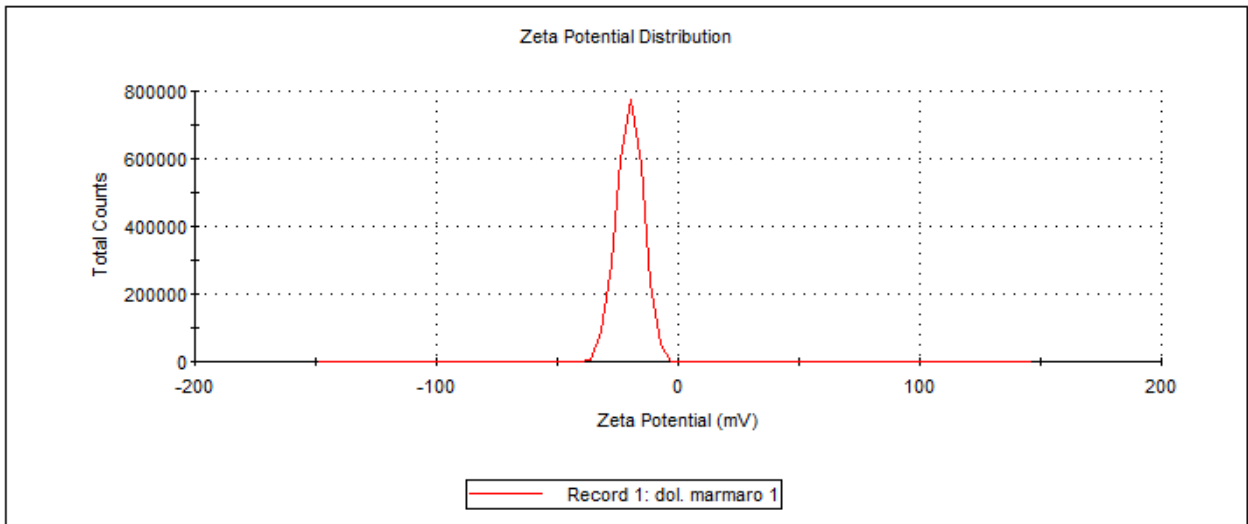
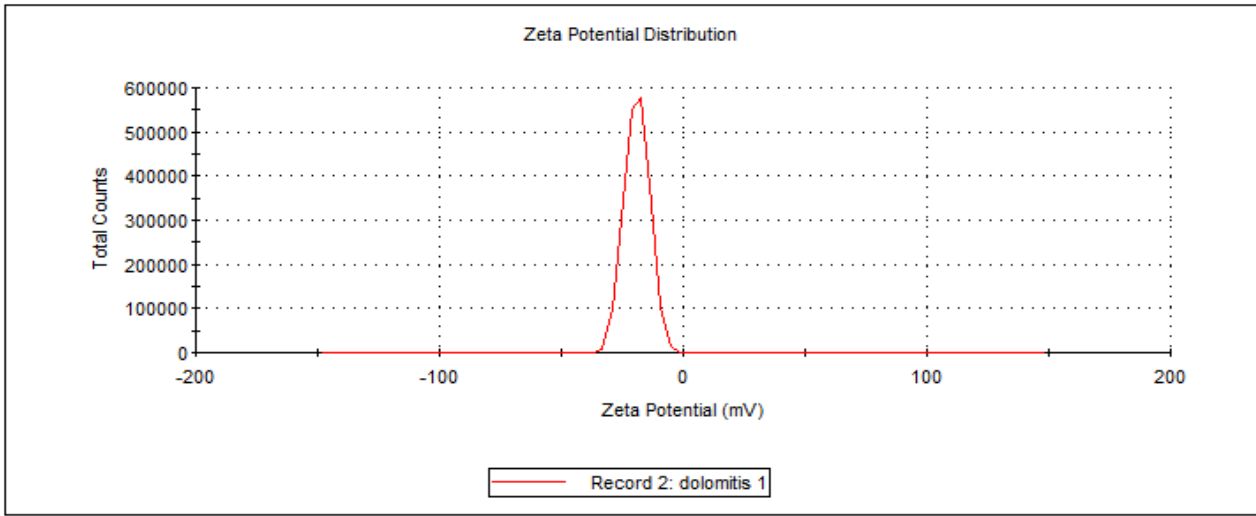
:

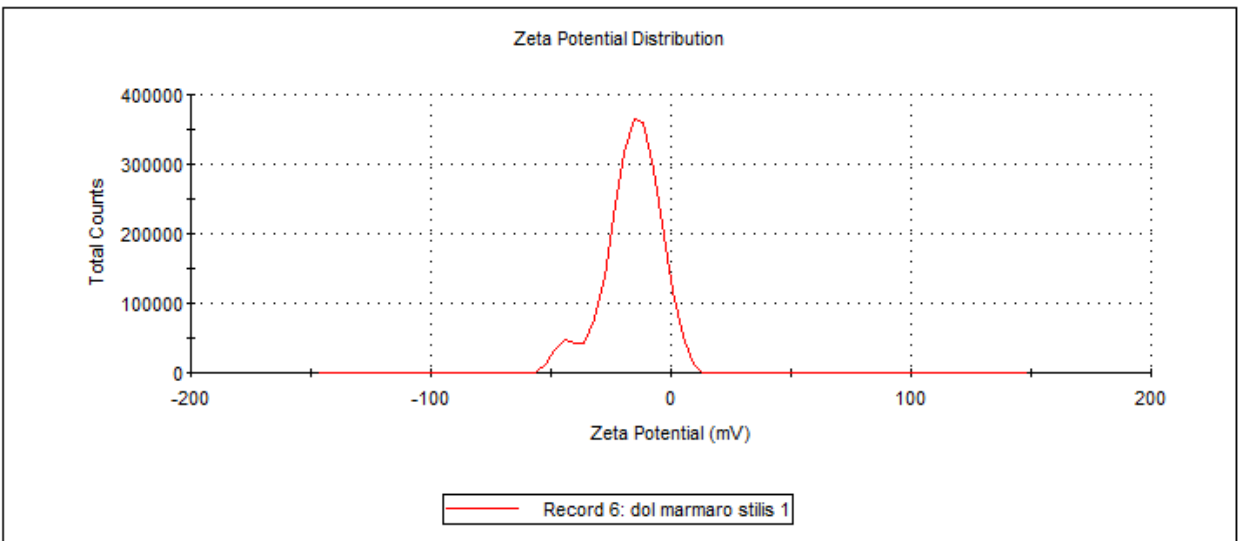
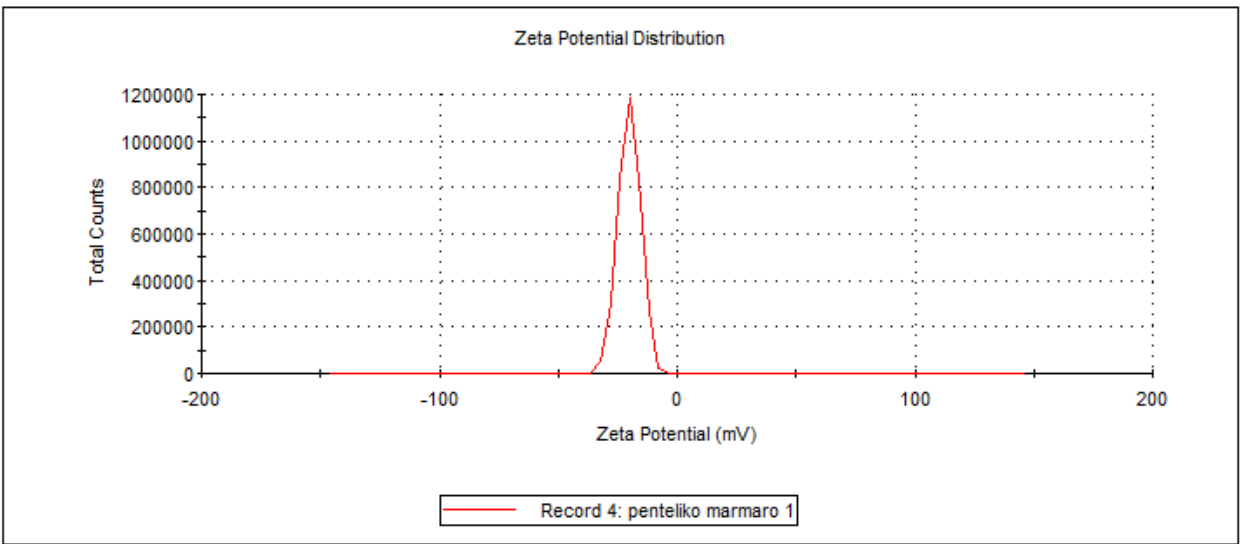
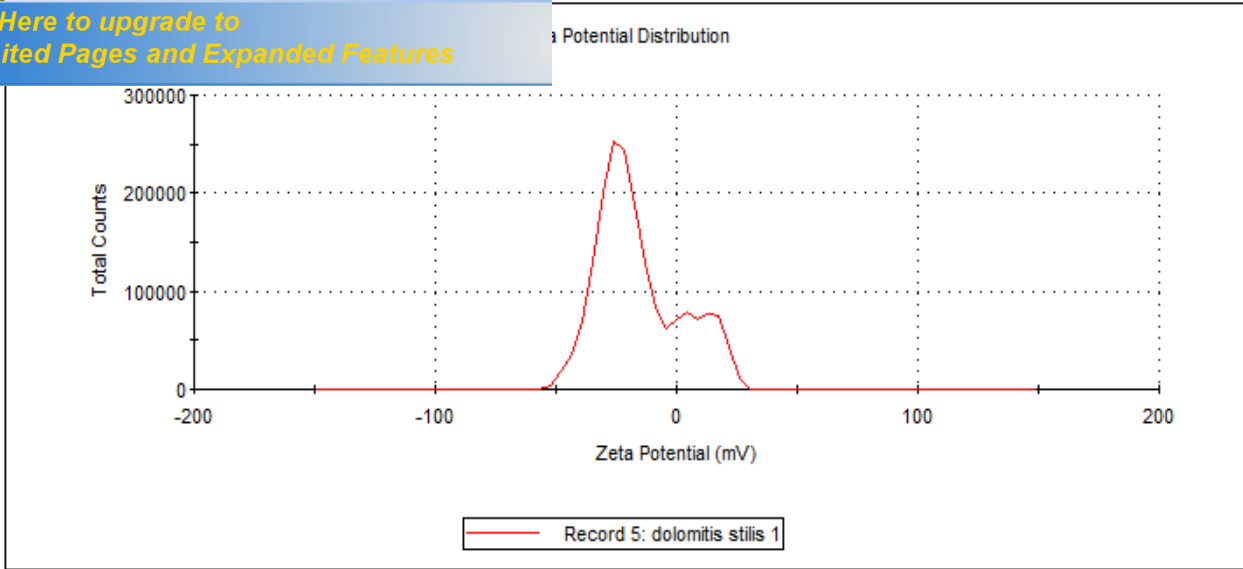


CEC



, Cr As,





1. W.A. Deer, R. A. Howie, J. Zussman, An Introduction to The Rock-Forming Minerals, 2<sup>nd</sup> Edition, Longman, Essex, 1994, pp. 620, 641, 644-645.
2. L. L. Y. Chang, R. A. Howie, J. Zussman, Rock-Forming Minerals: Non-silicates: Sulphates, Carbonates, Phosphates, Halides, Volume 5B, Second Edition, London, 1998, pp. 97-98, 189, 208-209.
3. Cornelis Klein, Cornelius S. Hurlbut, Jr., Manual of Mineralogy, 21<sup>st</sup> Edition, John Wiley & Sons, 1999, pp. 403-405, 411, 415-417, 580-581, 588.
4. A. V. Milovsky, O. V. Kononov, Mineralogy, Mir Publishers, Moscow, 1985, pp.215-217, 222.
5. Dexter Perkins, Mineralogy, 2<sup>nd</sup> Edition, Prentice Hall, New Jersey, 2002, pp. 142, 376, 379-380
6. L. G. Berry, Brian Mason, Mineralogy: Concepts-Descriptions-Determinations, W.H. Freeman and Company, 1959, pp. 399-401, 410-413.
7. William D. Ness, Introduction to Mineralogy, Oxford University Press, New York, 2000, pp. 334-335.
8. H. Battey, A Pring, Mineralogy for students, 3<sup>rd</sup> Edition, Longman, Essex, 1994, pp. 227-230.
9. . , « » , , 2005, . 70-86
10. , 2004, . 58.
11. , & , University Studio Press, 1996, . 13-14, 22-23, 41-42, 45-47, 59-61, 7, 207-224.
12. P. W. Harben, M. Kuffvart, Industrial Minerals ó A Global Geology, Industrial Minerals Information Ltd., London, 1996, pp. 81-86.
13. O , 1979, . 481-483, 506-510520-524, 527-528541.
14. William H. Blackburn, William H. Dennen, Principles of Mineralogy, 2<sup>nd</sup> Edition, Wm.C. Brown Publishers, 1994, pp. 274, 280-281.
15. Erol Pehlivan, Ali Mujdat Ozkan, Salih Dinc, Serife Parlayici, Adsorption of Cu<sup>2+</sup> and Pb<sup>2+</sup> ion on dolomite powder, Journal of Hazardous Materials, 167, 2009, 1044-1049.
16. John Waren, Dolomite: occurrence, evolution and economically important associations, Earth-Science Reviews, 52, 2000, 1-81.
17. Hans G. Machel, Jeff Lonnee, Hydrothermal dolomite-a product of poor definition and imagination, Sedimentary Geology, 152, 2002, 163-171.
18. Oleg S. Pokrovsky, Jacques Schott, Fabien Thomas, Dolomite surface speciation and reactivity in aquatic systems, Geochimica et Cosmochimica Acta, Vol. 63,1999 , pp. 3133ó 3143.
19. G. M. Walker, G. Connor, S. J. Allen, Copper (II) removal onto dolomitic sorbents, Chemical Engineering Research and Design, 82(A8), 2004, 961-966.
20. Yousef Salameh, Mohammad Ahmad, Steven Allen, Gavin Walker, Ahmad Albadarin, Nassir Al-Lagtah, Kinetic and Thermodynamic Investigations on Arsenic Adsorption onto Dolomitic Sorbents, Chemical Engineering Transactions, Volume 21, 2010, 793-798.
21. G.M. Walker, L. Hansen, J.-A. Hanna, S.J. Allen, Kinetics of a reactive dye adsorption onto dolomitic sorbents, Water Research, 37,, 2003, 2081ó2089.

- , Kheira Marouf-Khelifa, Jacques Schott, Amine Khelifa, from aqueous solutions by dolomitic sorbents, *Journal of Colloid and Interface Science*, 297, 2006, 45653.
23. Patrick V. Brady, Hans W. Papenguth, John W. Kelly, Metal sorption to dolomite surfaces, *Applied Geochemistry* 14, 1999, 569-579.
  24. Khary Cave, Federico I. Talens-Alession, Comparative effect of Mn(II) and Fe(III) as activators and inhibitors of the adsorption of other heavy metals on calcite, *Colloids and Surfaces A: Physicochem. Eng. Aspects*, 268, 2005, 19623.
  25. Shaban E. Ghazy, Abdullah H.M. Gad, Lead separation by sorption onto powdered marble waste, *Arabian Journal of Chemistry*, 2010, 1-10.
  26. S. E. Ghazy, A. H. Gad, Separation of Zn(II) by sorption onto powdered marble wastes, *Indian Journal of Chemical Technology*, Vol. 15, 2008, 433-442.
  27. Sevgi Kocaoba, Comparison of Amberlite IR 120 and dolomite's performances for removal of heavy metals, *Journal of Hazardous Materials*, 147, 2007, 488-496.
  28. Anthony M. Evans, *Ore Geology and Industrial Minerals ó An Introduction*, Blackwell Science, 3<sup>rd</sup> Edition, 1996, pp. 288-289.
  29. Charles C. Plummer, David McGearry, *Physical Geology*, Wm.C. Publishers, 7<sup>th</sup> Edition, 1996, pp. 124, 145.
  30. Robert S. Boynton, *Chemistry and Technology of Lime and Limestone*, John Wiley & Sons, 2<sup>nd</sup> Edition, 1980, pp. 9-10.
  31. Terry R. West, *Geology Applied to Engineering*, Prentice Hall, New Jersey, 1995, pp. 28, 55-56, 67-68.
  32. J. A. H. Oates, *Lime and Limestone-Chemistry and Technology, Productions and Uses*, Wiley-VCH, Weinheim, 1998, pp. 12,14.
  33. M. Kuffvart, *Developments in Economic Geology 18-Industrial Minerals and Rocks*, Elsevier, Prague, 1984, pp. 299.
  34. F. G. H. Blyth, M. H. de Freitas, *A Geology for Engineers*, Edward Arnold, 7<sup>th</sup> Edition, London, 1998, pp. 126.
  35. Peter W. Harben, *The Industrial Minerals HandyBook-A Guide to Markets, Specifications & Prices*, Industrial Minerals Information, 4<sup>th</sup> Edition, 2002, pp.78-92.
  36. Donald D. Carr, Lawrence F. Rooney, *Industrial Minerals and Rocks*, American Institute of Mining, Volume 2, Metallurgical and Petroleum Engineers, 5<sup>th</sup> Edition, 1983, pp. 833.
  37. <http://www.geo.auth.gr>
  38. <http://www.metal.ntua.gr>
  39. <http://www.mindat.org>
  40. <http://www.earthquestminerals.com>
  41. <http://geology.com>
  42. [www.sme.gr](http://www.sme.gr)
  43. [www.igme.gr](http://www.igme.gr)
  44. [www.lazaridismarmor.eu](http://www.lazaridismarmor.eu)
  45. <http://web.eps.utk.edu>
  46. [www.geology.about.com](http://www.geology.about.com)
  47. Nikos Arvanitidis, Northern Greece's industrial minerals: production and environmental technology developments, *Journal of Geochemical Exploration*, 62, 1998, pp. 217-227.
  48. Shinwoo Lee, James A. Dyer, Donald L. Sparks, Noel C. Scrivner, Evert J. Elzinga, A multi-scale assessment of Pb(II) sorption on dolomite, *Journal of Colloid and Interface Science*, 298, 2006, 20630.
  49. James A. Davis, Christopher C. Fuller, Alison D. Cook, A model for trace metal sorption processes at the calcite surface: Adsorption of Cd<sup>2+</sup> and subsequent solid solution formation, *Geochimica et Cosmochimica Acta*, 51, 1477-1490.

- C. T. Resch, Sorption of divalent metals on calcite, *Water, Air, and Soil Pollution*, 55, 1991, 1549-1562.
51. J. M. Zachara, J. A. Kutnick, J. B. Harsh, The mechanism of  $Zn^{2+}$  adsorption on calcite, *Geochimica et Cosmochimica Acta*, 52, 1988, 2281-2291.
  52. Yang Du, Fei Lian, Lingyan Zhu, Biosorption of divalent Pb, Cd and Zn on aragonite and calcite mollusk shells, *Environmental Pollution*, 159, 2011, 1763-1768.
  53. Xiaowu Tang, Zhenze Li, Yunmin Chen, Adsorption behavior of Zn(II) on calcinated Chinese loess, *Journal of Hazardous Materials*, 161, 2009, 824-834.
  54. J.A. Gómez del Río, P.J. Morando, D.S. Cicerone, Natural materials for treatment of industrial effluents: comparative study of the retention of Cd, Zn and Co by calcite and hydroxyapatite. Part I: batch experiments, *Journal of Environmental Management*, 71, 2004, 169-177.
  55. A. Martin-Garin, J.P. Gaudet, L. Charlet, X. Vitart, A dynamic study of the sorption and the transport processes of cadmium in calcareous sandy soils, *Waste Management*, 22, 2002, 2016-207.
  56. Neil C. Sturchio, Ronald P. Chiarello, Likwan Cheng, Paul F. Lyman, Michael J. Bedzyk, Yonglin Qian, Hoydoo You, Dennis Yee, Phillip Geissbuhler, Larry B. Sorensen, Yong Liang, Donald R. Baer, Lead adsorption at the calcite-water interface: Synchrotron X-ray standing wave and X-ray reflectivity studies, *Geochimica et Cosmochimica Acta*, 61, 1997, 251-263.
  57. A. García-Sánchez, E. Álvarez-Ayuso, Sorption of Zn, Cd and Cr on calcite. Application to purification of industrial wastewaters, *Minerals Engineering*, 15, 2002, 539-547.
  58. Susan L. Stipp, I. Michael F. Hochella Jr., George A. Park, James O. Leckie,  $Cd^{2+}$  uptake by calcite, solid-state diffusion, and the formation of solid-solution: Interface processes observed with near-surface sensitive techniques (XPS, LEED, and AES), *Geochimica et Cosmochimica Acta*, 56, 1992, 1941-1954.
  59. Rob N. J. Comans, Jack J. Middelburg, Sorption of trace metals on calcite: Applicability of the surface precipitation model, *Geochimica et Cosmochimica Acta*, 51, 1987, 2587-2591.
  60. Hamidi A. Aziz, Mohd. N. Adlan, Kamar S. Ariffin, Heavy metals (Cd, Pb, Zn, Ni, Cu and Cr(III)) removal from water in Malaysia: Post treatment by high quality limestone, *Bioresource Technology*, 99, 2008, 1578-1583
  61. G. Rangel-Porras, J.B. García-Magno, M.P. González-Muñoz, Lead and cadmium immobilization on calcitic limestone materials, *Desalination*, 262, 2010, 1610.
  62. K. Komnitsas, G. Bartzas, I. Paspaliaris, Efficiency of limestone and red mud barriers: laboratory column studies, *Minerals Engineering*, 17, 2004, 183-194.
  63. S. E. Ghazy, A. H. Ragab, Removal of Lead from Water Samples by Sorption onto Powdered Limestone, *Separation Science and Technology*, 42, 2007, 653-667.
  64. S. E. Ghazy, A. H. Ragab, Removal of copper from water samples by sorption onto powdered limestone, *Indian Journal of Chemical Technology*, 14, 2007, 507-514.
  65. Fatima Ouadjenia-Marouf, Reda Marouf, Jacques Schott, Ahmed Yahiaoui, Removal of Cu(II), Cd(II) and Cr(III) ions from aqueous solution by dam silt, *Arabian Journal of Chemistry*, 2010.
  66. Levent Alta, Ahmet Kili, Hasan Koyun, Mustafa Ilik, Adsorption of Cr(VI) on ureolytic mixed culture from biocatalytic calcification reactor, *Colloids and Surfaces B: Biointerfaces*, 86, 2011, 404-408.
  67. Quanyuan Chen, Colin D. Hills, Menghong Yuan, Huanhuan Liu, Mark Tyrer, Characterization of carbonated tricalcium silicate and its sorption capacity for heavy metals: A micron-scale composite adsorbent of active silicate gel and calcite, *Journal of Hazardous Materials*, 153, 2008, 775-783.

- rgouthi, Ayman A. Issa, Majeda A. Khraisheh, Gavin M. II), and Co(II) using natural sorbents: Equilibrium and kinetic studies, *Water Research*, 40, 2006, 2645-2658.
69. M.A. Hashim, Soumyadeep Mukhopadhyay, Jaya Narayan Sahu, Bhaskar Sengupta, Remediation technologies for heavy metal contaminated groundwater, *Journal of Environmental Management*, 92, 2011, 2355-2388.
  70. Mustafa Turan, Ugur Mart, Baris Yüksel, Mehmet S. Çelik, Lead removal in fixed-bed columns by zeolite and sepiolite, *Chemosphere*, 60, 2005, 148761492.
  71. Fenglian Fu, Qi Wang, Removal of heavy metal ions from wastewaters: A review, *Journal of Environmental Management*, 92, 2011, 407-418.
  72. Ping Zhou, Ju-Chang Huang, Alfred W. F. Li, Shirly Wei, Heavy Metal Removal From Wastewater In Fluidized Bed Reactor, *Wat. Res.*, 33, 1999, 1918-1924.
  73. Josep Oliva, Joan De Pablo, José-Luis Cortina, Jordi Cama, Carlos Ayora, Removal of cadmium, copper, nickel, cobalt and mercury from water by Apatite II<sup>TM</sup>: Column experiments, *Journal of Hazardous Materials*, 2011.
  74. Susan E. Bailey, Trudy J. Olin, R. Mark Bricka, D. Dean Adrian, A Review Of Potentially Low-Cost Sorbents For Heavy Metals, *Wat. Res.*, 33(11), 1999, 2469-2479.
  75. Sandhya Babel, Tonni Agustiono Kurniawan, Low-cost adsorbents for heavy metals uptake from contaminated water: a review, *Journal of Hazardous Materials*, B97, 2003, 2196243.
  76. P. Papadopoulos, D. L. Rowell, The reactions of cadmium with calcium carbonate surfaces, *Journal of Soil Science*, 39, 1988, 23-36.
  77. P. Papadopoulos, D. L. Rowell, The reactions of copper and zinc with calcium carbonate surfaces, *Journal of Soil Science*, 40, 1989, 39-48.
  78. A. Godelitsas, M. Kokkoris, P. Misaelides, Investigation of the interaction of Greek dolomitic marble with metal aqueous solutions using Rutherford backscattering and X-ray photoelectron spectroscopy, *Journal of Radioanalytical and Nuclear Chemistry*, 272, 2007, 3396344.
  79. L. Madrid, E. Diaz-Barrientos, Influence of carbonate on the reaction of heavy metals in soils, *Journal of Soil Science*, 43, 1992, 109-121.
  80. Upendra Kumar, Agricultural products and by-products as a low cost adsorbent for heavy metal removal from water and wastewater: A review, *Scientific Research and Essay*, 1(2), 2006, 033-037.
  81. Tonni Agustiono Kurniawan, Gilbert Y.S. Chan, Wai-hung Lo, Sandhya Babel, Comparisons of low-cost adsorbents for treating wastewaters laden with heavy metals, *Science of the Total Environment*, 366, 2006, 4096-426.
  82. Zied Djedidi, Jihen Ben Khaled, Ridha Ben Cheikh, Jean-François Blais, Guy Mercier, Rajeshwar Dayal Tyagi, Comparative study of dewatering characteristics of metal precipitates generated during treatment of monometallic solutions, *Hydrometallurgy*, 95, 2009, 61669.
  83. M.A. Barakat, New trends in removing heavy metals from industrial wastewater, *Arabian Journal of Chemistry*, 4, 2011, 3616377.
  84. Haluk Aydin, Yasemin Bulut, Çi dem Yerlikaya, Removal of copper (II) from aqueous solution by adsorption onto low-cost adsorbents, *Journal of Environmental Management*, 87, 2008, 37645.
  85. Murari Prasad, Huan-Yan Xu, Sona Saxena, Multi-Component Sorption Of Pb(II), Cu(II) And Zn(II) Onto Low-Cost Mineral Adsorbent, *Journal Of Hazardous Materials*, 154, 2008, 2216229.
  86. Guixia Zhao, Xilin Wu, Xiaoli Tan, Xiangke Wang, Sorption of Heavy Metal Ions from Aqueous Solutions: A Review, *The Open Colloid Science Journal*, 4, 2011, 19-31.



- Sollars, R. Perry, Low-cost adsorbent for a waste and The Science of the Total Environment, 116, 1192, 31-52.
88. V. K. Gupta, P. J.M. Carrott, M. M.L. Ribeiro Carrott, Suhas, Low-Cost Adsorbents: Growing Approach to Wastewater Treatmentô a Review, Critical Reviews in Environmental Science and Technology, 39:10, 2009, 783-842.
  89. Manuel Algarra, M. Victoria Jiménez, Enrique Rodríguez-Castellón, Antonio Jiménez-López, José Jiménez-Jiménez, Heavy metals removal from electroplating wastewater by aminopropyl-Si MCM-41, Chemosphere, 59, 2005, 7796786.
  90. E. Álvarez-Ayuso, A. García-Sánchez, X. Querol, Purification of metal electroplating waste waters using zeolites, Water Research, 37, 2003, 485564862.
  91. Flaviane Vilela Pereira, Leandro Vinícius Alves Gurgel, Laurent Frédéric Gil, Removal of Zn<sup>2+</sup> from aqueous single metal solutions and electroplating wastewater with wood sawdust and sugarcane bagasse modified with EDTA dianhydride (EDTAD), Journal of Hazardous Materials, 176, 2010, 8566863.
  92. D. Bo i , V. Stankovi , M. Gorgievski, G. Bogdanovi , R. Kova evi , Adsorption of heavy metal ions by sawdust of deciduous trees, Journal of Hazardous Materials, 171, 2009, 6846692.
  93. Rajender Kumar, Divya Bhatia, Rajesh Singh, Suman Rani, Narsi R. Bishnoi, Sorption of heavy metals from electroplating effluent using immobilized biomass Trichoderma viride in a continuous packed-bed column, International Biodeterioration & Biodegradation, 65, 2011, 1133-1139.
  94. Mohammad Ajmal, Rifaqat A.K. Rao, Rais Ahmad, Jameel Ahmad, Liaqat A.K. Rao, Removal and recovery of heavy metals from electroplating wastewater by using Kyanite as an adsorbent, Journal of Hazardous Materials, B87, 2001, 1276137.
  95. Ayhan Demirbas, Heavy metal adsorption onto agro-based waste materials: A review, Journal of Hazardous Materials, 157, 2008, 2206229.
  96. Anna Stafiej, Krystyna Pyrzynska, Adsorption of heavy metal ions with carbon nanotubes, Separation and Purification Technology, 58, 2007, 49652.
  97. G.M. Ayoub, M. Mehawej, Adsorption of arsenate on untreated dolomite powder , Journal of Hazardous Materials 148, 2007, 2596266.
  98. M.B. Mc Bride, Chemisorption of Cd<sup>2+</sup> on calcite surfaces soil, Science Society of America Journal, 44, 1980, 26-28.
  99. Metcalf & Eddy, : & , 4 , , 2006, . 1491-1492, 1475.
  100. MWIT, John Grittendor et al., Water treatment: Principles and Design, 2<sup>nd</sup> Edition, John Wiley & Sons, New Jersey, 2005, pp. 1246,1248.
  101. Frank Woodard, Industrial Waste Treatment Handbook-Pollution Engineering, Butterworth-Heinemann, Woburn, 2001, pp. 376.
  102. James M. Montgomery, Water Treatment Principles and Design, John Wiley & Sons, 1985, pp.175.
  103. Philip Engler, Mark W. Santana, Martin L. Mittleman, David Balazs, Non-Isothermal. In Situ Xrd Analysis Of Dolomite Decomposition, The Rigaku Journal, 5(2), 1988,3-8.
  104. R. Ozao, M. Ochiai, A. Yamazaki, R. Otsuka, Thermal analysis of ground dolomites , Thermochemica Acta, 183, 1991, 183-198.
  105. S. Gunasekaran, G. Anbalagan, Thermal decomposition of natural dolomite, Bull. Mater. Sci.,30(4), 2007, 3396344.
  106. P. Fritz, D. G. W. Snap, The isotopic composition of secondary dolomites, Geochimica et Cosmochimica Acta, 34, 1970, 1101-1173.
  107. A. Duffy, G.M. Walker, S.J. Allen, Investigations on the adsorption of acidic gases using activated dolomite, Chemical Engineering Journal, 117, 2006, 2396244.

108. Stergioudi, S.A. Tsipas, N. Michailidis, The use of  $\text{TiO}_2$  effect on the microstructure of aluminium metal foams. *Colloids and Surfaces A: Physicochem. Eng. Aspects*, 382, 2011, 1186-123.
109. Ronghua Zhang, Shumin Hu, Xuotong Zhang, Wenbin Yu, Dissolution Kinetics of Dolomite in Water at Elevated Temperatures, *Aquat. GEOchem.* 13, 2007, 309-338.
110. <http://cpe.njit.edu/dlnotes/CHE685/Cls06-2.pdf>
111. , « », , , 2009.
112. - , , 2006, . 123-127, 146-147, 225-226, 239-241, 263-278.
113. , , 2 , , , 2003, . 21, 31-34, 119-120.
114. , : , , 2004, . 119-120, 123-125.
115. , 2001, . 373-378, 386-387.
116. Skoog D. A., Holler F. J., Nieman T. A., , 5 , , 2002, . 350-352.
117. , - , , 2009.
118. [www.aua.gr](http://www.aua.gr)
119. [www.physics.ntua.gr](http://www.physics.ntua.gr)
120. [www.pursue.edu](http://www.pursue.edu)
121. [www.materialscience.uoregon.edu](http://www.materialscience.uoregon.edu)
122. Kirby, B.J., *Micro- and Nanoscale Fluid Mechanics: Transport in Microfluidic Devices*, Cambridge University Press, 2010.
123. Zeta Potential Using Laser Doppler Electrophoresis - Malvern.com
124. Lyklema J., *Fundamentals of Interface and Colloid Science*, 2, 1995, 3.208.
125. [www.atascientific.com.au/nano-particle-sizing.html](http://www.atascientific.com.au/nano-particle-sizing.html)

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